

SGS ARTS COLLEGE TIRUPATI

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1. Remedial Coaching

Slow learners require additional encouragement and support for better academic performance. The college regularly conducts remedial coaching for such students. This led to improved performance of the students in the university examinations.

Special programmes conducted for slow learners

Phone: (0877) 2264599

E-mail: sgsartscollegegettd@gmail.com

Web: sgsac.edu.in



TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

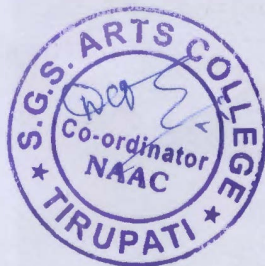
ODD SEMESTERS

CIRCULAR

Date: 11-06-2019

All the HODs of the department are here by instructed to plan remedial classes for slow learners of III & V semesters in order to improve their learning capability and for further improvement in the forthcoming university examinations. The Time –Table for the conduct of remedial classes for slow learners is enclosed. The classes may be conducted out of the college hours i.e. before 10.00A.M / after 4.00P.M


PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI



“OM NAMO VENKATESAYA”
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

NOTICE

Date: 14-06-2019

SGSAC/Chem dept/Remedial classes- 2019/HOD/01

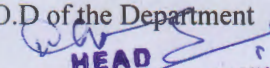
All the slow learners in chemistry of III&V semesters as enclosed in the list are here by intimated that remedial classes are planned to conduct from 17-06-2019 in Room numbers 225& 230 from 4.00Pm to 5.30 P.M. They are further instructed that attendance is mandatory. Stringent action will be taken against the students who will be absent for the remedial classes.

ACADEMIC YEAR 2019-20

II B.Sc, IIISEMESTER

S.No	HALL Ticket Number	Student Name	GROUP& Medium
1	NIL	NIL	MPC EM & TM
1	319008007	T. Goutham	MZC-EM
2	319008026	S.Soma Sekar	
1	318008054	G.Suman	CBZ-TM
2	318008050	P.SAI	
3	318008038	S. Balaji Naik	
4	318008055	E.G. Suri	
5	319008259	P.Narasimha	BBC-EM
6	319008269	D.Sasi Kumar	
7	319008248	M.Chandra	
8	319008242	S.Choudaiah	
9	319008254	S.Manjunadh	
10	319008261	M.Naveen	
11	319008280	T,Vijay	
12	319008247	K.Jagadeswara	
13	319008250	T.Jayaram	

14	319008257	V.Nagendra Prasad	
15	319008246	B.Hari Babu	
16	319008262	A.Pedda Obulesu	
17	319008274	T.SREEKANTH REDDY	
III BSC SEMESTER-V PAPER -5			
1	318008159	M.Kusuma Priya	BBC-EM
2	318008146	S.Gangadhar	
3	318008163	G.,Mahesh	
4	318008284	T.Akash	MZC-EM
5	318008265	K.Vamsi	
1	NIL	NIL	CBZ-TM
1	NIL	NIL	MPC-EM & TM
III BSC SEMESTER-V PAPER -6			
1	318008161	C.Madhu Mohan	BBC-EM
2	318008160	C.Lokesh	
3	318008159	M.Kusuma Priya	
4	318008163	G.,Mahesh	
5	318008284	T.Akash	MZC-EM
6	318008283	M.Sreedhar	
7	318008201	Chittaturu Gnana Sekhar	MPC-EM&TM
1	NIL	NIL	CBZ-TM

H.O.D of the Department

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

To

1. All teaching staff of chemistry
2. Department for circulation.
3. Department notice board

Phone: (0877) 2264599

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TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

S.G.S ARTS COLLEGE, TIRUCHANUR ROAD, TIRUPATI

Institutional Time -Table for Remedial classes (Odd semester)

Programme name : B.A, B.Sc, B.Com

SEMESTER : III & V

TIME: 4.00PM-5.30PM


YEAR : II & III

DAY	SUBJECT	SEMESTER	ROOM NO
MONDAY	CHEMISTRY	III	225
		V	230
	STATISTICS	III	221
		V	222
	COMPUTER SCIENCE	III	115
		V	116
	POLITICS	III	211
		V	212
	COMMERCE	III	307
		V	308
TUESDAY	CHEMISTRY	III	225
		V	230
	STATISTICS	III	221
		V	222
	COMPUTER SCIENCE	III	115
		V	116

	POLITICS	III	211
		V	212
	COMMERCE	III	307
		V	308
WEDNESDAY	MATHEMATICS	III	225
		V	230
	ZOOLOGY	III	Z LAB-2
		V	129
	BIOTECHNOLOGY	III	315
		V	316
	HISTORY	III	211
		V	212
	ACCOUNTS	III	307
		V	308
THURSDAY	MATHEMATICS	III	225
		V	230
	ZOOLOGY	III	Z LAB-2
		V	129
	BIOTECHNOLOGY	III	315
		V	316
	HISTORY	III	211
		V	212
	ACCOUNTS	III	307
		V	308
FRIDAY	PHYSICS	III	225
		V	230
	BOTONY	III	B LAB-2
		V	129

	MICROBIOLOGY	III	314
		V	315
	ECONOMICS	III	211
		V	212
	TELUGU	III	204
		V	205
SATURDAY	PHYSICS	III	225
		V	230
	BOTONY	III	B LAB-2
		V	129
	MICROBIOLOGY	III	314
		V	315
	ECONOMICS	III	211
		V	212
	TELUGU	III	205
		V	206




PRINCIPAL
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DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
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Departmental Time-Table for remedial classes

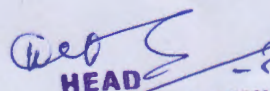
S.G.S. ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2016-17			
SEMESTERS - III & V		DATE : 27/06/2016 to 31/09/2016	
Name Of The Lecturer	MONDAY	TUESDAY	TIME
Dr. V. Venkata Lakshmi	III - SEMESTER	III - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	V - SEMESTER	V - SEMESTER	4.00 TO 5.30

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S.G.S. ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2017-18			
SEMESTERS - III & V		DATE : 19/06/2017 to 03/10/2017	
Name of the Lecturer	MONDAY	TUESDAY	TIME
Dr. V. Venkata Lakshmi	III - SEMESTER	III - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	V - SEMESTER	V - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2018-19			
SEMESTERS - III & V		DATE : 11/6/2018 to 11/9/2018	
Name Of The Lecturer	MONDAY	TUESDAY	TIME
Dr. V. Venkata Lakshmi	III - SEMESTER	III - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	V - SEMESTER	V - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2019-20			
SEMESTERS - III & V		DATE : 17/06/2019 to 29/10/2019	
Name Of The Lecturer	MONDAY	TUESDAY	TIME
Dr. V. Venkata Lakshmi	III - SEMSTER	III - SEMSTER	4.00 TO 5.30
Dr. P. SUGUNA	V - SEMESTER	V - SEMESTER	4.00 TO 5.30


 HEAD
 DEPARTMENT OF CHEMISTRY
 S. G. S. ARTS COLLEGE
 TIRUPATHI



Tiruchanur Road, TIRUPATI-517501

Date: 17-06-2019 To 29-10-2019

[illegible]

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TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501


EVEN SEMESTERS

CIRCULAR

Date: 02-12-2019

All the HODs of the department are here by instructed to plan remedial classes for slow learners of II , IV & VI semesters in order to improve their learning capability and for further improvement in the forthcoming university examinations. The Time –Table for the conduct of remedial classes for slow learners is enclosed. The classes may be conducted out of the college hours i.e. before 10.00A.M / after 4.00 P.M




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DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
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NOTICE

Date: 06-12-2019

SGSAC/Chem dept/Remedial classes- 2019/HOD/02

All the slow learners in chemistry of II, IV&VI semesters as enclosed in the list are here by intimated that remedial classes are planned to conduct from 09-12-2019 in Room numbers 224,225 & 230 from 4.00P.M to 5.30 P.M. They are further instructed that attendance is mandatory for the students mentioned below. Stringent action will be taken against the students who will be absent for the remedial classes.

List of Slow Learners

<u>LIST OF SLOW LEARNERS - 2019-2020</u>			
I B.Sc. , II SEMESTER			
S.No	Hall Ticket No.	Name of the student	Group & Medium
1	320008027	A. Venakatarjun	MPC (EM&TM)
2	320008028	A. Chandu Prakash	
3	320008002	B. Dinesh Naik	CBZ (TM)
4	320008003	C. Lakshmi	
5	320008004	C. Harinath	
6	320008006	C. Chandra	
7	320008008	D. Bala	
8	320008009	D. Munilatha	
9	320008010	E. Sree	

10	320008011	Gowrayya	
11	320008012	G. Rajasekhar	
12	320008014	K. Usha	
13	320008015	K. Swathi	
14	320008017	K. Madhu	
15	320008018	K. Naveen Naik	
16	320008020	M. Hareesh	
17	320008212	A. Giri Ramesh	MZC (EM)
18	320008249	S. Reena	
19	320008261	B. Govardhan	BBC (EM)
20	320008262	V. Jhansi Laskhmi	
21	320008265	G. Ramakanth	
22	320008286	M. Siva Kumar	
23	320008289	P. Sunil	
24	320008293	T. Umamaheswar	
25	320008294	S. Muni Sindhuja	
26	320008297	S.R.P. Jenny	
27	320008302	V. Nethesh Kumar	
II B.Sc. , IV SEMESTER			
S.no	Hall Ticket No.	Name of the student	Group & Medium
1	319008036	T. Asha	CBZ (TM)
2	319008044	K. Himabindu	
3	319008047	A. Manendra	
4	319008239	K. Bhargav	BBC (EM)
5	319008241	B. Chinna	
6	319008242	B. Haribabu	
7	319008246	K. Jagadeeswara	
8	319008247	M. Chandra	
9	319008248	T. Jayaram	

10	319008250	S. Manjunath	
11	319008254	V. Nagendra Prasad	
12	319008257	P. Narasimha	
13	319008259	M Naveen	
14	319008261	A. Pedda Obulesu	
15	319008262	D. Sasi Kumar	
16	319008269	U. Someswara	
17	319008273	T. Sreekanth Reddy	
18	319008274.	T. Vikay	
III B.Sc. , VI SEMESTER			
S.no	Hall Ticket No.	Name of the student	Group & Medium
1	318008146	S. Gangadar	MPC (EM&TM)
2	318008159	M. Kusuma Priya	
3	318008163	G. Mahesh	
4	318008161	C. Madhu Mohan	
5	318008265	K. Vamsi	MZC (EM)
6	318008201	C. Gnana Sekahr	
7	318008240	P. Sivakumar Reddy	

H.O.D of the Department
HEAD
 DEPARTMENT OF CHEMISTRY
 S. G. S. ARTS COLLEGE
 TIRUPATHI

To

1. All teaching staff of chemistry Department for circulation.
2. Department notice board

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TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

INSTITUTIONAL TIME - TABLE FOR REMEDIAL CLASS (Even semester)

S.G.S ARTS COLLEGE,TIRUPATI			
COURSE : B.A, B.Sc, B.Com		SEMESTER:II , IV & VI	
TIME:4.00PM-5.30PM		YEAR: I,II& III	
DAY	SUBJECT	SEMESTER	ROOM NO
MONDAY	CHEMISTRY	II	224
		IV	225
		VI	230
	STATISTICS	II	220
		IV	221
		VI	222
	COMPUTER SCIE4NCE	II	114
		IV	115
		VI	116
	POLITICAL SCIENCE	IV	210
		VI	212
		II	306
	COMMERCE	IV	307
		VI	308
		II	224
TUESDAY	CHEMISTRY	II	224
		IV	225
		VI	230

	STATISTICS	II	220
		IV	221
		VI	222
	COMPUTER SCIENCE	II	114
		IV	115
		VI	116
	POLITICAL SCIENCE	II	210
		IV	211
		VI	212
	COMMERCE	II	306
		IV	307
		VI	308
WEDNESDAY	MATHEMATICS	II	224
		IV	225
		VI	230
	ZOOLOGY	II	Z LAB
		IV	ZLAB-2
		VI	129
	BIOTECHNOLOGY	II	LAB
		IV	315
		VI	316
	HISTORY	II	210
		IV	211
		VI	212
	ACCOUNTS	II	306
		IV	307
		VI	308
THURSDAY	MATHEMATICS	II	224
		IV	225
		VI	230
	ZOOLOGY	II	LAB
		IV	Z LAB-2

		VI	129
	BIOTECHNOLOGY	II	LAB
		IV	315
		VI	316
	HISTORY	II	210
		IV	211
		VI	212
	ACCOUNTS	II	306
		IV	307
VI		308	
FRIDAY	PHYSICS	II	224
		IV	225
		VI	230
	BOTONY	II	B-LAB
		IV	B-LAB-2
		VI	129
	MICROBIOLOGY	II	LAB
		IV	314
		VI	315
	ECONOMICS	II	210
		IV	211
		VI	212
	TELUGU	II	203
		IV	204
		VI	205
SATURDAY	PHYSICS	II	224
		IV	225
		VI	230
	BOTONY	II	B-LAB
		IV	B LAB-2
		VI	129
	MICROBIOLOGY	II	LAB

	ECONOMICS	IV	314
		VI	315
		II	210
		IV	211
		VI	212
		II	203
	TELUGU	IV	205
		VI	206



[Signature]
PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

“OM NAMO VENKATESAYA”

**TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY**

**Sri D.PARAMESWARA
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HOD OF CHEMISTRY**

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Mail.Id : parameswara1964@gmail.com**

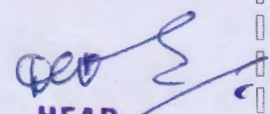
Departmental Time-Table for remedial classes

S.G.S. ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2016-17			
SEMESTERS - II, IV & VI		DATE : 23/01/2017 to 11/04/2017	
Name Of The Lecturer	MONDAY	TUESDAY	TIME
SRI D. PARAMESWARA	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. V. Venkata Lakshmi	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	II - SEMESTER	II - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2017-18			
SEMESTERS - II, IV & VI -		DATE : 06/12/2017 to 03/04/2018	
Name Of The Lecturer	MONDAY	TUESDAY	TIME
SRI D. PARAMESWARA	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. V. Venkata Lakshmi	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	II - SEMESTER	II - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2018-19			
SEMESTERS - II, IV & VI		DATE : 12/11/2018 to 11/02/2018	
Name Of The Lecturer	MONDAY	TUESDAY	
SRI D. PARAMESWARA	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. V. Venkata Lakshmi	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	II - SEMESTER	II - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2019-20			
SEMESTERS - II, IV & VI		DATE : 09/12/2019 to 03/03/2020	
Name Of The Lecturer	MONDAY	TUESDAY	TIME
SRI D. PARAMESWARA	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. V. Venkata Lakshmi	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	II - SEMESTER	II - SEMESTER	4.00 TO 5.30
Smt. K. REVATHI	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30


 HEAD
 DEPARTMENT OF CHEMISTRY
 S. G. S. ARTS COLLEGE
 TIRUPATHI



TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Attendance list of slow learners of all programmes with chemistry combination

Semester: II, IV&VI

Date: 09-12-2019 to 03-03-2020

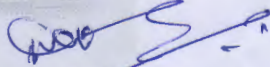
Name of the student		Date											
Roll No.	Name	09-12-19	10-12-19	11-12-19	12-12-19	13-12-19	14-12-19	15-12-19	16-12-19	17-12-19	18-12-19	19-12-19	20-12-19
P.P.C. (S.N.A.T.M.)													
2249	A. Venkata Dargam	P	P	P	P	P	P	P	P	P	P	P	P
2250	A. Chandra Prakash	P	P	P	P	P	P	P	P	P	P	P	P
C.C.S. (M.T.)													
2007	B. Divya Lakshmi	P	P	P	P	P	P	P	P	P	P	P	P
2008	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2009	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2010	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2011	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2012	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2013	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2014	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2015	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2016	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2017	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2018	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2019	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2020	C. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
M.S.S. (L.T.T.)													
2214	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2215	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
S.S.C. (L.T.T.)													
2216	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2217	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2218	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2219	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2220	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2221	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2222	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2223	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2224	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2225	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2226	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2227	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2228	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2229	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P
2230	A. Anitha	P	P	P	P	P	P	P	P	P	P	P	P

Name of the Institute:

II BSc IV Sem

SL NO.	ADMISSION NO.	NAME	CASTE	1	2	3	4	5	6	7	8	9	10	11
				7/12	10/12	16/12	17/12	23/12	24/12	30/12	31/12	7/1	20/1	21/1
		<u>CBZ (TH)</u>												
31900	8036	T. Asha		P	P	P	P	P	P	P	P	P	P	A
	8044	K. Himabindhu		P	P	P	P	A	P	P	P	P	P	P
	8047	A Manendra		P	A	P	P	P	P	P	P	A	P	P
		<u>BBC (VET)</u>												
8239		K. Banghav		P	P	P	P	P	P	A	P	P	P	P
8241		B. Chinna		P	P	P	A	P	P	P	P	P	P	P
8242		S. Chaudaiah		P	A	P	P	P	P	P	P	P	A	P
8246		B. Haibabu		P	P	P	P	P	A	P	P	P	P	P
8247		K. Jagadeeswara.		P	P	P	A	P	P	P	P	P	P	P
8248		H. Chandra		A	P	P	P	P	P	P	P	P	P	A
8250		T. Jayaram		P	P	P	P	A	P	P	P	P	P	P
8254		S. Hanjunadh		P	P	P	P	P	P	P	A	P	P	P
8257		V. Nagendra prasad		P	P	A	P	P	P	P	P	P	P	P
8258		P. Narasimha		A	P	P	P	P	P	P	P	P	P	P
8261		M. Naveen		P	P	P	P	P	A	P	A	P	P	P
8262		A. Perla Subela		P	P	P	P	P	P	A	P	P	P	P
8269		D. Sasi Kumar		P	P	A	P	P	P	P	P	P	P	P
8273		U. Somaswara		P	P	P	P	P	A	P	P	P	P	P
8274		T. Sreekanth Reddy		P	P	P	P	P	P	A	P	P	P	P
8280		T. Vijay		P	P	P	P	P	P	P	A	P	P	P
				16/12	17/12	18/12	19/12	20/12	21/12	22/12	23/12	24/12	25/12	26/12

SL. NO.	ADMISSION NO.	NAME	CASTE	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
		<u>MPC (ENTH)</u>																															
31800	8146	S. Ganga dar		P	D	P	D	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	
8159		M. Kuberma priya		P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	
8163		G. Mahesh		P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	
8161		C. Madhu Mohan		P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	
		<u>MZC (ENT)</u>																															
8265		K. Vamsi		P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	
		<u>BBC (ENT)</u>																															
8201		C. Gnana Sekhar		P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	
8240		P. Sivakumar Reddy		P	P	P	P	P	A	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	
				P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Students engaged in Study hours conducted for slow learners by the Department of Chemistry.







Students engaged in peer learning

Department of Microbiology

1. Remedial Coaching

Slow learners require additional encouragement and support for better academic performance. The college regularly conducts remedial coaching for such students. This led to improved performance of the students in the university examinations.

Special programmes conducted for slow learners

Phone: (0877) 2264599

E-mail: sgsartscollegettd@gmail.com

Web: sgsac.edu.in



Sri V. VENKATALAKSHMI
M.Sc., Mphil., (Ph.D)
HOD OF MICROBIOLOGY

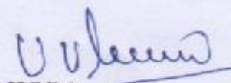
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

ODD SEMESTERS
CIRCULAR

Date: 11-06-2019

As Per the instructions of the principal we are planning to conduct remedial classes for slow learners to improve their performance in the university examinations. The classes may be conducted out of the college hours i.e., before 10 A.M /after 4 P.M . The time tabel of remedial classes for slow learners is enclosed.


(HOD)

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

“OM NAMO VENKATESAYA”
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI
DEPARTMENT OF MICROBIOLOGY

Dr V. Venkatalakshmi
M.Sc.,(Ph.D)
HOD OF MICROBIOLOGY

Phone No: 9492075161

NOTICE

Date: 25-06-2019

SGSAC/Microbiology dept/Remedial classes- 2019/HOD/01

All the slow learners in Microbiology of III&V semesters as enclosed in the list are here by intimated that remedial classes are planned to conduct from 28-06-2019 in Room numbers 314& 315 from 4.00Pm to 5.30 P.M. They are further instructed that attendance is mandatory. Stringent action will be taken against the students who will be absent for the remedial classes.

III - SEMESTER :

S. No	Hall Ticket No.	Name of the student	Group & Medium
1	0319008001	A Balaji	MZC(EM)
2	0319008005	D Ganesh	
3	0319008014	E Naveen	
4	0319008018	G Pranay kumar	
5	0319008026	D somasekhar	
6	0319008027	P subramanyam	
7	0319008030	G swarna latha	
8	0319008034	B vijay kumar	
9	0319008035	N vinod kumar	
10	0319008033	N venkatesh	

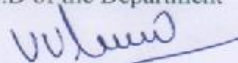
V - SEMESTER :

S. No	Hall Ticket No.	Name of the student	Group & Medium
1	0318008258	B anwarlal hussain	MZC (EM)
2	0318008267	Sk mazeed babu	

To

1. All teaching staff of microbiology
2. Department for circulation.
3. Department notice board

H.O.D of the Department



HEAD

DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
T. 517501

“OM NAMO VENKATESAYA”
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI
DEPARTMENT OF MICROBIOLOGY

Dr V. Venkatalakshmi
M.Sc.,(Ph.D)
HOD OF MICROBIOLOGY

Phone No: 9492075161

Departmental Time-Table for remedial classes

S.G.S. ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI			
DEPARTMENT OF MICROBIOLOGY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2019-20			
SEMESTERS - III & V		DATE : 28/06/2019to 31/10/2019	
Name Of The Lecturer	MONDAY	TUESDAY	TIME
Dr. K. SRIDEVI	III - SEMSTER	III - SEMSTER	4.00 TO 5.30
Dr. K. SRIDEVI	V - SEMESTER	V - SEMESTER	4.00 TO 5.30

"OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI
DEPARTMENT OF MICROBIOLOGY

Dr. V. VENKATALAKSHMI

Phone No: 9492075161

M.Sc.,(Ph.D)

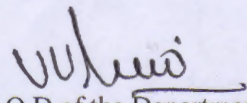
HOD OF MICROBIOLOGY

NOTICE

Date: 03-12-2019

SGS AC /Microbiology dept/Remedial classes- 2019/HOD/01

All the slow learners in Microbiology of II, IV&VI semesters as enclosed in the list are here by intimated that remedial classes are planned to conduct from 06-12-2019 in Room numbers lab, 314& 315 from 4.00P.M to 5.30 P.M. They are further instructed that attendance is mandatory for the students mentioned below. Stringent action will be taken against the students who will be absent for the remedial classes.


H.O.D of the Department

HEAD

DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

To

1. All teaching staff of microbiology Department for circulation.
2. Department notice board

“OM NAMO VENKATESAYA”

**TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI
DEPARTMENT OF MICROBIOLOGY**

Sri V. Venkata Lakshmi

M.Sc.,Mphil.,(Ph.D)

HOD OF MICROBIOLOGY

Phone No: 9492075161

Mail.ID: parameswara1964@gmail.com

Departmental Time-Table for remedial classes

S.G.S. ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI			
DEPARTMENT OF Microbiology			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2019-20			
SEMESTERS - II, IV & VI		DATE : 06/12/2019 to 07/03/2020	
Name Of The Lecturer	MONDAY	TUESDAY	TIME
SRI K. SRIDEVI	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. K. SRIDEVI	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. K. SRIDEVI	II - SEMESTER	II - SEMESTER	4.00 TO 5.30
Dr. K. SRIDEVI	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30



TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
 Tiruchanur Road, TIRUPATI-517501

Attendance list of slow learners of all programmes with Microbiology combination

Semester: III&V

Date: 28-06-2019 To 19-10-2019

2019-

Name of the Office

Attendance

Register

Month
Year

S. No.	NAME	Designation	1	2	3	4	5	6	7	8	9	10
III Year MZC V Sem (19-10-2019 To 28-06-2019)												
HTNO Name												
1.	31200849 B. Anuradha Lakshmi	V.10	A	A	A	A	A	A	A	A	A	A
	8267 S.K. Nagaraj Babu	A	A	A	A	A	A	A	A	A	A	A

II Year III Sem (I year II Sem 19-19)												
HTNO Name												
1.	31200801 A. Balaji	3.0	A	A	A	A	A	A	A	A	A	A
2.	8005 B. Ganesh	4.3	A	A	A	A	A	A	A	A	A	A
3.	8014 E. Naveen	3.7	A	A	A	A	A	A	A	A	A	A
4.	8018 G. Pranay Kumar	3.9	A	A	A	A	A	A	A	A	A	A
5.	8026 P. Somasekhara	3.0	A	A	A	A	A	A	A	A	A	A
6.	8027 P. Subramanyam	3.9	A	A	A	A	A	A	A	A	A	A
7.	8030 G. Suresh Lakshmi	2.4	A	A	A	A	A	A	A	A	A	A
8.	8034 B. Vijay Kumar	2.4	A	A	A	A	A	A	A	A	A	A
9.	8035 N. Vinod Kumar	2.5	A	A	A	A	A	A	A	A	A	A
10.	8032 N. Venkatesh	A	A	A	A	A	A	A	A	A	A	A

11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	Remarks
29/10/19	30/10/19	31/10/19	01/11/19	02/11/19	03/11/19	04/11/19	05/11/19	06/11/19	07/11/19	08/11/19	09/11/19	10/11/19	11/11/19	12/11/19	13/11/19	14/11/19	15/11/19	16/11/19	17/11/19	18/11/19	
H																					
A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	
A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	A	



TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Attendance list of slow learners of all programmes with Microbiology combination
Semester: II, IV & VI Date: 09-12-2019 to 03-03-2020

Name of the Office			Attendance										Register																														
			Month Year																																								
S. No.	NAME	Designation	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	Remarks									
III year VI Sem			(19-20 VI Sem)										29/12/14/15/21/22/29/31/1/3																														
HTNo Name			6/12/13/20/27/28/29/31/1/2																																								
1.	NIL																																										
II year IV Sem			(19-20 IV Sem)																																								
HTNo Name			II year 19-20 IV Sem																																								
1.	0319008007 T. Gowthami	A	//	//	//	//	A	//	//	//	//	//	//	//	//	A	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//											
2.	8030 G. Subimalata	2.2	//	//	//	//	//	A	//	//	A	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//											
3.	8035 M. Vinod Kumar	3.2	//	//	//	A	//	//	//	//	A	//	//	//	//	A	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//											
4.	8029 B. Sankar	A	//	A	//	//	//	//	//	//	A	//	//	//	//	A	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//											
5.	8026 D. Somasekhar	A	//	//	//	//	A	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//											
I year MZL II Sem			(I year I Sem 19-20)																																								
HTNo Name																																											
1.	0320008007 M. Murali Prasad	3.4	A	//	//	//	//	//	//	//	A	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//											
2.	8012 A. Ramakrishna	2.9	//	//	A	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//											
3.	8021 C. Surekh Babu	3.7	//	//	//	A	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//	//											



Phone: (0877) 2264599

-mail:sgsartscollegett@gmail.com

Web:sgsac.edu.in

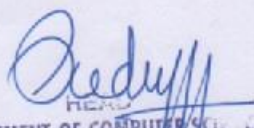
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501



LIST OF SLOW LEARNERS

II B.Sc. , III SEMESTER

S.no	Hall Ticket No.	Name of the student	Group&Medium
1	0318008075	SANAMBATLA HIMAVANTH	MSCs-EM


HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Phone: (0877) 2264599

E-mail:sgsartscollegeittd@gmail.com

Web:sgsac.edu.in



"OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S. ARTSCOLLEGE, TIRUPATI
AIDED (FOUNDED IN 1969)



TIME-TABLE FOR THE ACADEMIC YEAR - 2018-19

DATE : 11/6/2018 to 11/9/2018

SEMESTERS - III

NAME OF THE LECTURER	MONDAY	TUESDAY	TIME
B.TRIVENI N.JAYALAKSHMI	III - SEMESTER	III - SEMESTER	4.00 TO 5.30

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S.ARTS COLLEGE, TTD, TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

III & V SEMESTERS ATTENDANCE REGISTER FOR SLOW LEARNERS :: 2018-2019

S.No	Name of the Student	11/06	12/6	18/6	19/6	25/6	26/6	28/6	5/6	7/6	9/6	10/6	16/6	23/6	24/6	30/6	31/6	6/8	7/8	13/8	14/8	20/8	21/8	27/8
II YEAR III SEMESTER MSCS[EM]																								
1	SANAMBATLA HIMAVANTH	1	1	1	A	1	1	1	1	A	1	1	1	1	1	1	A	1	1	1	1	1	1	A

[Signature]

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Phone: (0877) 2264599

-mail:sgsartscollegettd@gmail.com

Web:sgsac.edu.in

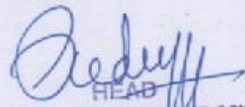
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501



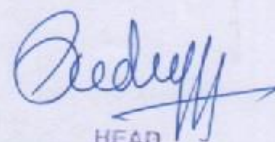
CLASS TEST MARK SHEET FOR SLOW LEARNERS

II B.Sc. III SEMESTER

S.no	Hall Ticket No.	Name of the student	Group& Medium	TEST-1 (25M)	TEST-2 (25M)
1	0318008075	SANAMBATLA HIMAVANTH	MSCs-EM	20	22


HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE,
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE					
PROGRESS REPORT OF SLOW LEARNERS IN COMPUTER SCIENCE AFTER REMEDIAL COACHING					
LIST OF SLOW LEARNERS					
S.NO	NAME OF THE STUDENT	STATUS OF STUDENTS (SLOW)	Group & Medium	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER
1	SANAMBATLA HIMAVANTH	SLOW LEARNERS	II YEAR MSCS-EM	42	44



HEAD

DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

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Web: sgsac.edu.in

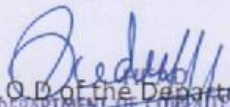
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501



Date:- 08/11/2018

NOTICE

All the slow learners in Computer Science of IV Semester as enclosed in the list are here by intimated that remedial classes are planned to conduct from 12/11/2018 in Room numbers 114 from 4.00P.M to 5.30 P.M. They are further instructed that attendance is mandatory. Stringent action will be taken against the students who will be absent for the remedial classes.


H.O.D. of the Department
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

To

1. All teaching staff of Computer Science Department for circulation.
2. Department notice board

Phone: (0877) 2264599

E-mail:sgsartscollegettd@gmail.com

Web:sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501



LIST OF SLOW LEARNERS

II B.Sc. , IV SEMESTER

S.no	Hall Ticket No.	Name of the student	Group&Medium
1	0318008118	G VAMSI KRISHNA	MSCs-EM
2	0318008099	KESANI SAI KUMAR	
3	0318008075	SANAMBATLA HIMAVANTH	

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Phone: (0877) 2264599

E-mail:sgsartscollegegettd@gmail.com

Web:sgsac.edu.in



"OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S. ARTSCOLLEGE, TIRUPATI
AIDED (FOUNDED IN 1969)

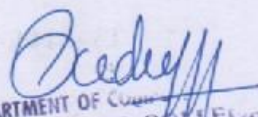


TIME-TABLE FOR THE ACADEMIC YEAR :: 2018-19

DATE : 12/11/2018 to 11/02/2018

SEMESTER – IV

NAME OF THE LECTURER	MONDAY	TUESDAY	TIME
B.TRIVENI N.JAYALAKSHMI	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30


DEPARTMENT OF COMMERCE
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S.ARTS COLLEGE, TTD, TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

II, IV & VI SEMESTERS ATTENDANCE REGISTER FOR SLOW LEARNERS :: 2018-2019

S.No	Name of the Student	12/11	13/11	19/11	20/11	26/11	27/11	3/12	4/12	10/12	11/12	17/12	31/12	7/1	21/1	22/1	28/1	29/1	4/2	5/2	11/2
II YEAR IV SEMESTER MSCS[EM]																					
1	G VAMEI KRISHNA	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
2	KISANI SAI KUMAR	P	P	P	A	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P
3	SANAMBATLA HIMAVANTH	A	P	P	A	P	P	P	A	A	P	P	P	A	P	P	P	P	P	P	A

[Signature]

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

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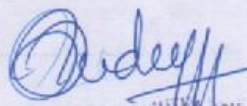
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501



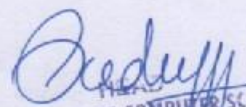
CLASS TEST MARK SHEET FOR SLOW LEARNERS

II B.Sc. , IV SEMESTER

S.no	Hall Ticket No.	Name of the student	Group& Medium	TEST-1 (25M)	TEST-2 (25M)
1	0318008118	G VAMSI KRISHNA	II YEAR IV	20	21
2	0318008099	KESANI SAI KUMAR	SEM	18	20
3	0318008075	SANAMBATLA HIMAVANTH	MSCs-EM	15	AB


DEPARTMENT OF COMPUTER
S.G.S. ARTS COLLEGE
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE					
PROGRESS REPORT OF SLOW LEARNERS IN COMPUTER SCIENCE AFTER REMEDIAL COACHING					
LIST OF SLOW LEARNERS					
S.NO	NAME OF THE STUDENT	STATUS OF STUDENTS (SLOW)	Group & Medium	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER
1	G VAMSI KRISHNA	SLOW LEARNERS	II YEAR IV SEM MSCS	20	66
2	KESANI SAI KUMAR			18	41
3	SANAMBATLA HIMAVANTH			22	AB


 HEAD
 DEPARTMENT OF COMPUTER SCIENCE
 S.G.S. ARTS COLLEGE
 TIRUPATI

2. Rubrics

A rubric is an evaluation tool which promotes consistent application of learning expectations and learning standards in the class room. The college conduct a good number of classes to solve previous examination question papers. This programme helps the slow learners to augment their subject knowledge and ensures better scores in the university examinations.

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[Total No. of Pages : 3]

THREE YEAR B.Sc. DEGREE EXAMINATION, Sept./Oct. - 2014
FIRST YEAR EXAMINATION
Part II - Chemistry
Paper I - ORGANIC, INORGANIC, PHYSICAL AND GENERAL
CHEMISTRY
(Revised from 2009-2010)

Time : 3 Hours **Max. Marks : 100**

Answer any ten questions, choosing atleast two questions from each unit. Each question carries 10 marks.

ఏదేని పది ప్రశ్నలకు సమాధానములు వ్రాయుము. ప్రతి యూనిట్ నుండి కనీసము రెండు ప్రశ్నలను ఎంపిక చేయండి. ప్రతి ప్రశ్నకు 10 మార్కులు.

Unit - I
(Inorganic chemistry)

1. What are oxides. Discuss their classification on the basis of. (2)
ఆక్సైడ్లు అనగా నేమి? వాటి వర్గీకరణ క్రింది వాని ఆధారంగా చర్చింపుము.
a) oxygen content. (4)
ఆక్సిజన్ పరిమాణము.
b) chemical behaviour. (4)
రసాయన స్వభావము.
2. Draw the molecular orbital diagram of CO and O₂ molecules. Find the bond order of them. (10)
CO మరియు O₂ అణువుల అణు ఆర్బిటల్ శక్తి పటములను వ్రాయుము వాటి బంధ క్రమమును కనుగొనుము.
3. Explain general characteristics of group 16 elements. (10)
గ్రూపు 16 మూలకాల సాధారణ లక్షణాలను వివరించండి.
4. Write preparation and applications of graphitic compounds. (10)
గ్రాఫైట్ సమ్మేళనాల తయారీ మరియు ఉపయోగములను వ్రాయుము.

S-160 **(1)** **[P.T.O.]**

5. Write the applications of organometallic compounds. (10)

సంగ్రహించు రసాయన పరివర్తనలు యొక్క ఉపయోగాలను సమీక్షించుకోవడం ద్వారా.

Unit - II

(Organic Chemistry)

6. Explain with mechanism for the following reactions of benzene.
బెంజీన్ పై ఈ క్రింది చర్యలను వివరించు మరియు వాటిని వివరించు.

a) Nitration (5)

నైట్రేషన్

b) Friedel crafts alkylation reaction. (5)

ఫ్రీడెల్ - క్రాఫ్ట్ ఆల్కైలేషన్ చర్య.

7. What is meant by mesomeric effect? How it explains acidity of phenol and acidity of carboxylic acid. (10)

మీసమర్పక ప్రభావం అంటే ఏమిటి? ఫీనాల్ యొక్క ఆమ్లత్వమును మరియు కార్బాక్సిలిక్ ఆమ్ల ఆమ్లత్వమును, దీని ద్వారా ఎలా వివరించవచ్చును.

8. Explain orientation of aromatic substitution. (10)

ఎరోమాటిక్ ప్రతిక్షేపణ యొక్క స్థాన నిర్దేశకతను వివరించు.

9. Explain the following.

ఈ క్రింది వాటిని వివరించు.

a) Hyper conjugation. (5)

అతి సంయుగ్మము.

b) Diels-older reaction. (5)

డీల్స్ - ఆల్డర్ చర్య

10. Explain inductive effect with examples. How it effect Basicity of amines and acidity of carboxylic acid. (10)

ప్రేరేపక ప్రభావమును ఉదాహరణలతో వివరించు మరియు అది ఎమైన్ల బేసికతను మరియు ఆమ్ల ఆమ్లత్వమును ఏవిధంగా ప్రభావితం చేయును?

Unit - III

(Physical chemistry)

11. State and explain Nernst distribution Law. Explain any four applications of it (10)

నెర్స్ట్ విభజన నియమమును తెల్పి, వివరించు. దాని యొక్క నాలుగు అనువర్తనములను తెలుపు.

12. Deduce the values of critical constants in terms of vander waals constants. (10)

వాండర్ వాల్ స్థిరాంకముల పరంగా సందిగ్ధ స్థిరాంకములను ఉత్పాదించు విధానము తెల్పుము.

13. What is critical solution temperature? Explain the phenol water system. (10)

సందిగ్ధ ద్రావణ ఉష్ణోగ్రత అనగా నేమి? ఫినాల్ - నీటి వ్యవస్థను వివరింపుము.

14. What is Joule Thomson effect? Explain. (10)

జౌల్ - థామ్సన్ ప్రభావము అనగా నేమి? వివరింపుము.

15. Derive Bragg's equation. (10)

బ్రాగ్ సమీకరణమును ఉత్పాదించుము.

Unit - IV

(General chemistry)

16. What is meant by R, S - configuration? Explain with examples. (10)

R, S విన్యాసము అనగా నేమి? ఉదాహరణలతో వివరింపుము.

17. Write the difference between physical and chemical adsorption. (10)

అనురూపాత్మక సాదృశ్యము అనగా నేమి? n - బ్యూటీన్ యొక్క అనురూపాత్మక సాదృశ్యాన్ని వివరించుము.

18. Explain the following

ఈ క్రింది వాటిని వివరింపుము.

a) Compton effect (5)

కాంప్టన్ ఫలితము

b) Photo electric effect. (5)

ఫోటో ఎలక్ట్రిక్ ప్రభావము.

19. What is conformational isomerism? Explain the conformational isomerism of n-butane. (10)

అనురూపాత్మక సాదృశ్యము అనగా నేమి? n - బ్యూటీన్ యొక్క అనురూపాత్మక సాదృశ్యాన్ని వివరించుము.

20. What is meant by a colloid? Explain optical and electrical properties of colloids. (10)

కొల్లాయిడు అనగానేమి? దృశ్య మరియు విద్యుత్ ధర్మాలను వివరింపుము.

THREE YEAR B.Sc. DEGREE EXAMINATION — MARCH/APRIL 2014
 THIRD YEAR EXAMINATION
 Part II — Chemistry
 Paper III — INORGANIC, ORGANIC AND PHYSICAL CHEMISTRY
 (Revised from 2010-2011)

Time : 3 hours

Max. Marks : 100

Answer any TEN questions, choosing not less than THREE from each Section.

All questions carry equal marks.

ప్రతి సెక్షన్ నుండి మూడు ప్రశ్నలకు రెక్కువ కాకుండా వదిలి ప్రశ్నలకు సమాధానములు వ్రాయుము.

అన్ని ప్రశ్నలకు మార్కులు సమానుము.

(Marks: 10 × 10 = 100)

SECTION - I

సెక్షన్ - I

(INORGANIC CHEMISTRY)

What is crystal field theory? Discuss splitting of d-orbitals in case of tetrahedral and octahedral complexes. (10)

స్పటిక క్షేత్ర సిద్ధాంతమును తెలుపుము? అష్ట, పంచ, మరియు చతుర్ముఖ సమ్మేళనాలలో d-ఆర్బిటల్ విభజనను వివరించుము.

Explain the determination of composition of a complex by mole-ratio method. (10)

సంశ్లేష్ట సమ్మేళనాల సంఘటనమును మోల్-నిష్పత్తి పద్ధతి ద్వారా ఎట్లు నిర్ణయించెదరో తెలుపుము.

Explain SN^1 and SN^2 reaction mechanism in coordination compounds. (10)

సంశ్లేష్ట సమ్మేళనాలలో SN^1 మరియు SN^2 చర్యల చర్యా విధానమును వ్రాయుము.

Explain electronic absorption spectrum of $[Ti(H_2O)_6]^{3+}$. (10)

$[Ti(H_2O)_6]^{3+}$ అయాన్ యొక్క ఎలక్ట్రానిక్ కోషుల వర్ణవలమును వివరించుము.

(a) Explain biological significance of Na^+ and K^+ ions. (5)

జీవ వ్యవస్థలో Na^+ మరియు K^+ అయాన్ల ప్రాముఖ్యతను తెలుపుము.

(b) Explain biological significance of copper. (5)

జీవ వ్యవస్థలో కాపర్ యొక్క ప్రాముఖ్యతను తెలుపుము.

Explain the geometrical isomerism in complexes. (10)

సంశ్లేష్టాలలోని క్షేత్ర సార్వత్రికమును వివరించండి.

SECTION — II

పక్షన్ - II

(ORGANIC CHEMISTRY)

7. What are amines? Discuss the action of nitrous acid on primary, secondary and tertiary amines. (10)

ఎమిన్స్ అంటే ఏమిటి? సైమరీ, సెకండరీ మరియు టర్షరీ ఎమిన్స్ లో నైట్రస్ ఆమ్లం చర్య ఏ విధంగా ఉంటుందో తెలపండి.

8. What are heterocyclic compounds? Write any two methods of preparation of pyrrole. Explain its acid character. (10)

హెటరీయ వలయ సమ్మేళనాలు అనగా నేమి? పిర్లోల్ తయారు చేయుటకు రెండు పద్ధతులను వ్రాయండి. పిర్లోల్ ఆమ్ల గుణత్యాన్ని వివరించండి.

9. What are Amino Acids? Discuss their classification and explain (10)

(a) Zwitter ion structure and

(b) Isoelectric point.

ఎమైన్స్ ఆమ్లాలు అనగా నేమి? వాటి వర్గీకరణ గూర్చి వ్రాసి

(a) జ్యైట్టర్ ఆయాన్ నిర్మాణము మరియు

(b) సమ విద్యుత్ స్థానములను గురించి వివరించండి.

10. Write the structural elucidation of glucose. (10)

గ్లూకోజ్ నిర్మాణమును గురించి వివరించండి.

11. Write notes on the following :

(a) Hoffman's bromamide reaction with mechanism (5)

హోఫ్మన్ బ్రోమైమైడ్ చర్య మరియు చర్యా విధానము

(b) Carbylamine reaction. (5)

కార్బైల్ ఎమైన్ చర్యలను గూర్చి గూర్చి క్లుప్తముగా వివరించండి

12. Write the preparation of pyridine and its nucleophilic substitution reactions.

పిరిడిన్ తయారు చేయు పద్ధతిని తెల్పి, దాని న్యూక్లియోఫిలిక్ ప్రతిక్షేపణ చర్యలను వ్రాయుము.

SECTION – III
(PHYSICAL CHEMISTRY)

13. Derive an equation for the rate constant of a first order reaction and write its characteristics. (10)

ప్రథమ క్రమాంక చర్య రేటు స్థిరాంక సమీకరణమును ఉత్పాదించుము మరియు దాని లక్షణాలను తెలుపుము.

14. Give the various definitions for the first law of thermodynamics. Define C_p and C_v and derive the relation between them. (10)

ఉష్ణ గతిక శాస్త్ర ప్రథమ నియమాన్ని తెలిపే వివిధ నిర్వచనాలను వ్రాయుము. C_p మరియు C_v లను నిర్వచించి వాటి మధ్యగల సంబంధమును రాబట్టుము.

15. State and explain Einstein's Law of photochemical equivalence. (10)

ఐన్స్టీన్ కాంతి రసాయన తుల్యతా నియమమును నిర్వచించి వివరించుము.

16. Derive and explain Kirchoff's equation and explain the Carnot cycle. (10)

కిర్కాఫ్ సమీకరణమును ఉత్పాదించి వివరించుము. మరియు కార్నాట్ చలయమును గురించి వివరించుము.

17. State and explain quantum yield. Explain high quantum yield in $H_2 - Cl_2$ reaction. (10)

క్వాంటమ్ ప్రాప్తి గూర్చి తెల్పి వివరించుము. $H_2 - Cl_2$ చర్య యొక్క అధిక క్వాంటమ్ ప్రాప్తి గూర్చి వివరించండి.

18. Describe various methods for determining the order of a chemical reaction. (10)

చర్య క్రమాంకాలను కనుగొను పద్ధతులను గూర్చి వ్రాయుము.

(23051)

THREE YEAR B.Sc. DEGREE EXAMINATION, APRIL 2014

SECOND YEAR EXAMINATION

Part II - Chemistry

Paper II — INORGANIC, ORGANIC, PHYSICAL AND GENERAL CHEMISTRY

Time : 3 Hours

Max. Marks : 100

Answer any TEN questions choosing not less than TWO from each Section.

All question carry equal marks.

ప్రతి సెక్షన్ నుండి కనీసము రెండు ప్రశ్నలను ఎంచుకొను ఏ వది ప్రశ్నలకైనా సమాధానములు వ్రాయుము.

అన్ని ప్రశ్నలకు మార్కులు సమానము.

(Marks : 10 × 10 marks = 100 marks)

SECTION - I

సెక్షన్ - I

(INORGANIC CHEMISTRY)

(మూలక రసాయన శాస్త్రము)

Explain the colour and magnetic behaviour of transition elements.

పరివర్తన మూలకాల రంగు మరియు ఆయస్కాంత ధర్మములను వివరింపుము.

Give an account of the following

- (a) Lanthanide contraction
- (b) Oxidation states of f-block elements.

ఈ క్రింది వాటిపై వ్యాఖ్యానించుము.

- (a) లాంథనైడ్ సంకోచము.
- (b) f-బ్లాకు మూలకాల ఆక్సీకరణ స్థితులు.

Explain the free electron theory and give any two of its limitations.

స్వేచ్ఛా ఎలక్ట్రాన్ సిద్ధాంతమును వివరించుము. ఆ సిద్ధాంతమును రోహితును ఏదేని రెండింటిని తెలుపుము.

Explain the band theory of conductors, insulators and semiconductors.

వాహకాలు, అర్ధ వాహకాలు మరియు నిరోధకాలను రోహితు వల్ల సిద్ధాంతపరంగా వివరించండి.

[P.T.O]

5. Discuss the structures and shapes of $[\text{Ni}(\text{Co})_4]$ and $[\text{Fe}(\text{Co})_5]$.

$[\text{Ni}(\text{Co})_4]$ మరియు $[\text{Fe}(\text{Co})_5]$ ల నిర్మాణాలను, ఆకారములను వివరింపుము.

SECTION - II

పెక్షన్ - II

(ORGANIC CHEMISTRY)

(కర్తవ్య రసాయన శాస్త్రము)

6. Give the mechanism of SN^2 reaction. Explain the stereochemistry of SN^2 reactions example.

SN^2 చర్యల చర్యా విధానమును తెలుపుము. SN^2 చర్యల త్రిమితీయ రసాయన శాస్త్రమును సాదాసాదా వివరింపుము.

7. Write a note on the following

- (a) Synthesis of alcohols from grignard reagents
- (b) Pinacol-pinacolone rearrangement

ఈ క్రింది వాటిపై చర్యా వ్యాఖ్యను వ్రాయుము.

- (a) గ్రిగ్నార్డ్ కారకాల నుండి ఆల్కహాల్ సంశ్లేషణ
- (b) పినకాల్-పినకోలన్ పునఃఅమరిక

Write the mechanisms of the following reactions.

- (a) Aldol condensation
- (b) Perkin reaction.

ఈ క్రింది వాటికి చర్యా విధానాలను వివరింపుము.

- (a) ఆల్డోల్ సంఘనన చర్య
- (b) పెర్కిన్ చర్య.

Explain the following reactions.

- (a) Huns - Dieckmann reaction
- (b) HVZ-reaction

ఈ క్రింది చర్యలను వివరింపుము.

- (a) హన్స్-డికర్ప్ చర్య
- (b) HVZ-చర్య.

How do you synthesise the following from aceto acetic ester.

- (a) *n*-Butyric acid
- (b) Succinic acid
- (c) Crotonic acid
- (d) Heterocyclic compound.

అసిటో, అసిటేట్ నుండి ఈ క్రింది వాటిని ఎలా విశ్లేషిస్తారు.

- (a) ఎన్-బ్యూటీరిక్ ఆసిడ్
- (b) సుసమిక్ ఆసిడ్
- (c) క్రోనోమిక్ ఆసిడ్
- (d) హెటిరో సైకిలిక్ సమ్మేళనం.

SECTION - III

సెక్షన్ - III

(PHYSICAL CHEMISTRY)

(భౌతిక రసాయన శాస్త్రము)

11. Define phase rule. Draw and explain the phase diagram of NaCl - H₂O system.

ప్రాచుర్య నియమాన్ని నిర్వచించుము. NaCl - H₂O వ్యవస్థ యొక్క ప్రాచుర్య చిత్రాన్ని గీచి వివరించుము.

12. Define osmotic pressure. How will you determine the osmotic pressure of a solution experimentally?

ప్రవాహనరణ పీడనము అనగా నేమి? ప్రయోగ పూర్వకంగా ఒక ద్రావణం యొక్క ప్రవాహనరణ పీడనాన్ని ఎ నిర్ణయిస్తారు.

13. State and explain Kohlrausch's law. Briefly explain any two of its applications.

కోల్ రాష్ నియమాన్ని తెలిపి, ఏదేని రెండు అనువర్తనాలను వివరించుము.

14. Write notes on the following

- (a) Ostwald's dilution law
- (b) Calomel electrode.

ఈ క్రింది వాటిని వివరించుము.

- (a) ఆస్వాల్డ్ విలీన ద్రావణాల నియమము
- (b) కాలోమెల్ ఎలక్ట్రోడ్.

15. Define EMF of a cell. How EMF of a cell can be experimentally measured.

ఘటక EMF అనగా నేమి? ఘటక EMF ను ప్రయోగపూర్వకంగా ఎలా కనుగొంటారు.

SECTION - IV

పెక్షన్ - IV

(GENERAL CHEMISTRY)

(సాధారణ రసాయన శాస్త్రము)

16. Define symmetry. Explain the different elements of symmetry with examples.

సౌష్ఠ్యత అనగా నేమి? వివిధ సౌష్ఠ్యవాంశాలను ఉదాహరణలతో వివరింపుము.

17. Briefly explain the principle involved in complexometric titration's.

సంక్లంఘితృక అంశ మాపనాల సిద్ధాంతాన్ని క్లుప్తంగా వివరింపుము.

18. Define and explain "Significant figures" and its importance.

సార్థక సంఖ్యలను నిర్వచించి వివరింపుము. వాటి ప్రత్యేకతను వివరింపుము.

19. What is asymmetric synthesis? Give any two methods of asymmetric synthesis.

అసౌష్ఠ్య సంశ్లేషణ అనగా నేమి? అసౌష్ఠ్య సంశ్లేషణ పద్ధతులను ఏదేని రెండింటినీ తెలుపుము.

20. Write a note on acid base indicators.

ఆమ్లక్షార సూచికలను వివరింపుము.

3. Simplified study material

Slow learners are those students' whose learning pace is slower than their peers. The respective class teachers provide simplified study notes to enhance the confidence in the subject and augment their learning peace.

Photochemistry

కాంతి రసాయన శాస్త్రము

కాంతి శక్తి వల్ల ప్రేరితమయ్యే రసాయన చర్యలను కాంతి రసాయన చర్యలు అనీ, కాంతి రసాయన చర్యలను వివరించే రసాయన శాస్త్ర విభాగాన్ని కాంతి రసాయన శాస్త్రము అంటారు.

కాంతి రసాయన చర్యలకు, ఉష్ణ రసాయన చర్యలకు మధ్య భేదాలు:

కాంతి రసాయన చర్యలు	ఉష్ణ రసాయన చర్యలు
1. కాంతి రసాయన చర్యలు కాంతి శక్తి శోషణం చెందుతుంది.	1. ఉష్ణ రసాయన చర్యలు ఉష్ణ శక్తి వెలువడుతుంది (అని) గ్రహించబడుతుంది.
2. కాంతి సమక్షంలో మాత్రమే చర్య జరుగుతుంది.	2. ఈ చర్య జరుగుటకు కాంతి అవసరం లేదు.
3. చర్యలేదు మీద ఉష్ణోగ్రత ప్రభావం ఉండదు. అయితే చర్యలేదు కాంతి తీవ్రతకు అనులోమానుపాతం ఉంటుంది.	3. చర్యలేదు మీద ఉష్ణోగ్రత ప్రభావం చాలా ఎక్కువగా ఉంటుంది. ఉష్ణోగ్రతను 10° పెంచినప్పుడు రేటు రెండు (అని) మూడు రెట్లు పెరుగుతుంది.
4. అన్ని కాంతి రసాయన చర్యలకు AG ఋణాత్మకంగా ఉండకపోవచ్చు.	4. అన్ని ఉష్ణ రసాయన చర్యలకు AG ఋణాత్మకంగా ఉంటుంది.

కాంతి రసాయన శాస్త్ర నియమాలు :

గ్రోథీస్ - డేవీస్ నియమం మరియు స్టార్క్ - వినస్టీన్ నియమం కాంతి రసాయన చర్యలను వివరించును. ఈ నియమాలను కాంతి రసాయన శాస్త్ర నియమాలు అంటారు.

1. గ్రోథీస్ - డేవీస్ నియమం :-

ఒక చర్యావృణస్థైత్యం కాంతి పతనం చెందినప్పుడు దానిలో కొంతభాగం వృణస్థైత్య శోషణం చెందుతుంది, కొంతభాగం వృణస్థైత్య సూర్య ఉపరస్థ చెందుతుంది, మిగిలినది వృణస్థైత్య చర్య సరావధికి చెందించబడుతుంది.

గ్రోథీస్ - డేవీస్ నియమం ద్వారా "చర్యావృణస్థైత్య శోషణం చెందిన కాంతి మాత్రమే దానిలో రసాయన చర్యలకు కలుగజేస్తుంది".

కానీ చర్యావృణస్థైత్యం కాంతి శోషణం చెందిన ఉత్తేజానికి దానిలో రసాయన చర్య జరుగుతుంది. అన్ని భావించకూడదు. వృణస్థైత్య రసాయన చర్య జరగడానికి అనుకూల పరిస్థితి లేనప్పుడు శోషణి కాంతి వలన ఎలాంటి రసాయన చర్య జరుగదు. చర్యావృణస్థైత్య రసాయన చర్య జరిగింది అంటే పూర్తిగా శోషణి కాంతి వల్లనే జరుగుతుంది.

2. స్టార్క్ - వినస్టీన్ నియమం (అని) కాంటి-శక్తి తెల్లతా నియమం :-

కాంతి రసాయన చర్యల పూర్ణత చర్య అనువు ఒక క్షయం శక్తిని (అని) ఒక ఫోటాన్ శక్తిని శోషించుకొంటుంది. ఈ నియమాన్ని స్టార్క్ - వినస్టీన్ నియమం అంటారు.

చర్మావ్యవస్థల శోషణం చెందిన కాంతి పోషకపునాం 'ν' అనుకుంటే ②
కాంతి రసాయన చర్మాల పాల్గొనే ఒక అణువు శోషించుకొన్న శక్తి

$$E = h\nu$$

$$E = h \frac{c}{\lambda}$$

'N' అణువుల (ఒక మోల్) శోషించుకొన్న శక్తి

$$E = N h \frac{c}{\lambda}$$

$$E = \frac{2.86}{\lambda} \text{ cal/mole}$$

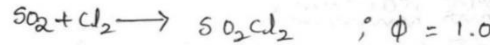
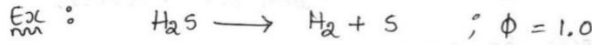
కాంతి రసాయన చర్మాల పాల్గొనే ఒక మోల్ అణువుల శోషించుకొన్న శక్తి పరిమాణాన్ని ఒక "బెన్ థాన్" అంటారు.

క్యాంటం (పోటీ (0.1) క్వాంటం ద్వారా :

ఒక కాంతి రసాయన చర్మాల పాల్గొనే అణువుల సంఖ్యకు మరియు చర్మావ్యవస్థల శోషణం చెందిన క్వాంటం ల సంఖ్యకు మధ్యగల నిష్పత్తినే, ఓ కాంతి రసాయన చర్మాలయొక్క క్వాంటం ప్రాప్తి అంటారు. దీనిని ϕ అనే సంకేతం చేత సూచిస్తారు.

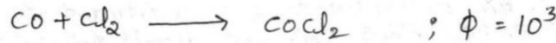
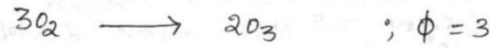
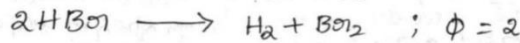
$$\text{క్వాంటం ప్రాప్తి, } \phi = \frac{\text{చర్మాల పాల్గొన్న అణువుల సంఖ్య}}{\text{చర్మావ్యవస్థల శోషణం చెందిన క్వాంటంల సంఖ్య}}$$

* చర్మావ్యవస్థల శోషణం చెందిన క్వాంటంల సంఖ్య మరియు చర్మాల పాల్గొన్న అణువుల సంఖ్య సమానంగా ఉంటే అప్పుడు ఓ చర్మాలకు క్వాంటం ప్రాప్తి ϕ ఒకటి ఉంటుంది. దీనిని "సాధారణ" క్వాంటం ప్రాప్తి అంటారు.

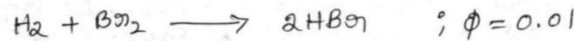


* అసాధారణ క్వాంటం ప్రాప్తి :

చర్మాల పాల్గొనే అణువుల సంఖ్య, శోషణం చెందిన క్వాంటం సంఖ్య కన్న ఎక్కువగా ఉంటే దానిని "అధిక క్వాంటం ప్రాప్తి" అంటారు.



* చర్మాల పాల్గొనే అణువుల సంఖ్య, శోషణం చెందిన క్వాంటం సంఖ్య కన్న తక్కువగా ఉంటే దానిని "అల్ప క్వాంటం ప్రాప్తి" అంటారు.



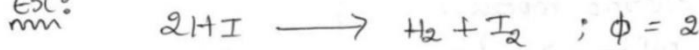


3

అధిక స్థాంతుల ప్రాథమిక కారణాలు :

చర్బామలతమలని ఒక అణువు స్థాంతుల శక్తిని శోషించుకొని ఉత్తేజితం చెందుతుంది. ఈ విధంగా ఉత్తేజితం చెందిన అణువు తర్వాత శుంఖల చర్బాను కొనసాగిస్తుంది. చర్బాలు పొల్లనే అణువుల సంఖ్య శోషణం చెందిన స్థాంతుల సంఖ్య కన్నా అధికంగా ఉంటుంది. ఫలితంగా స్థాంతుల దక్షత "1" కన్నా ఎక్కువగా ఉంటుంది.

ఉ:



చర్బావిధానం :



పై చర్బాలు ఒక స్థాంతుల శోషణం చెందుట వలన రెండు అణువుల చర్బాలు పొల్లనును. కనుక ఈ చర్బా యొక్క స్థాంతుల దక్షత "2"

అల్ప స్థాంతుల ప్రాథమిక కారణాలు :

అల్ప స్థాంతుల ప్రాథమిక కారణాలను క్రింది విధంగా వర్ణిస్తూ చూద్దాము.

- i. చర్బావృష్టల శోషణం చెందిన కాంతి శక్తి వానిలో రసాయన చర్బాను కలుగజేయడానికి తగినంతగా తీసుకురావడం దక్షత తక్కువగా ఉంటుంది.
- ii. చర్బావృష్టలని ఒక అణువు స్థాంతుల శక్తిని శోషించుకొని ఉత్తేజితం చెందుతుంది. ఈ విధంగా ఉత్తేజితం చెందిన అణువు చర్బాలు పొల్లడానికి ముందే అణుత్రిజితం చెందుతుంది. అలాంటి పరిస్థితులలో స్థాంతుల దక్షత "1" కన్నా తక్కువగా ఉంటుంది.
- iii. చర్బావృష్టలని ఒక అణువు స్థాంతుల శక్తిని శోషించుకొని పరమాణువులుగా అవి రాడికల్స్ గా విడిపోతుంది. ఇవి వృష్టలని ఇతర అణువులతో చర్బాజరిపి ఉత్పన్నాలను ఏర్పరచుటకు బదులుగా ఒక వానితోనాకటి కలసిపోయి పూర్ణపు అణువును ఏర్పరచును. ఇలాంటి పరిస్థితులలో స్థాంతుల దక్షత "1" కన్నా తక్కువగా ఉంటుంది.

ఉ:

H_2 మరియు Br_2 ల మధ్య జరిగే కాంతి రసాయన చర్బా.

1. హైడ్రోజన్ మరియు క్లోరిన్ మధ్య జరిగే కాంతి రసాయన చర్బా అధిక స్థాంతుల ప్రాప్తి అని, H_2 మరియు Br_2 ల మధ్య జరిగే కాంతి రసాయన చర్బా అల్పస్థాంతుల ప్రాప్తి కల్గివుండుటకు గల కారణాన్ని వాటి చర్బా

విధానాల పరంగా వివరించండి?

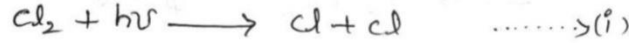
(4)

Ans. H_2 మరియు Cl_2 ల మధ్య జరిగే కాంతి రసాయన చర్య అధిక స్థాంతుం ప్రాప్తిని కలిగి ఉంటుంది. $H_2 + Cl_2 \longrightarrow 2HCl$; $\phi = 10^6$

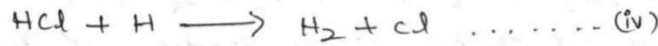
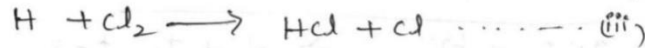
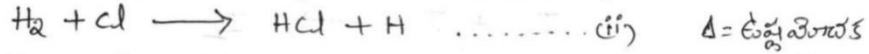
ఈ చర్య అధిక స్థాంతుం ప్రాప్తిని చర్చావిధానం పరంగా వివరించవచ్చు.

Stage I :- ప్రాథమిక ప్రక్రియ :

చర్చామిశ్రమంలోని ఒక Cl_2 అణువు స్థాంతుం శక్తిని శోషించుకొని పరమాణువులుగా వియోగం చెందును.



Stage II :- పై విధంగా విచ్ఛేదనం జరిగిన పరమాణువు చర్చామిశ్రమంలోని ఉండే తెలియని గొణి చర్చాలను కలుగజేయును.

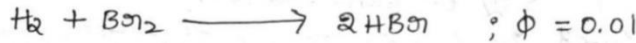


ద్వితీయ ప్రక్రియలలో (ii) & (iii) చర్చలు పలుమార్లు జరిగి చర్చామిశ్రమంలోని మొత్తం H_2 మరియు Cl_2 అణువులు ఉత్పన్నమయ్యాయి. ఈ విధంగా ఈ కాంతి రసాయన చర్య "శృంఖల చర్చ"గా నిర్వచించబడింది. కనుక ఇది "అధిక స్థాంతుం ప్రాప్తిని" కలిగి ఉంటుంది. చర్చామిశ్రమంలో శోషణం చెందే ఒకే స్థాంతుం శక్తి వలన అనేక అణువులు చర్చలు పొందును.

ఈ చర్చ శృంఖల చర్చగా కొనసాగడానికి కారణం, ద్వితీయ ప్రక్రియలలోని గొణి చర్చ (ii) ఉష్ణమోచక చర్చ కావడమే.

(ii) H_2 మరియు B_{12} ల మధ్య జరిగే కాంతి రసాయన చర్చ :

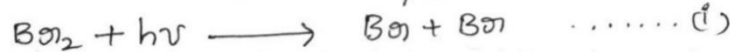
ఈ చర్చ అల్ప స్థాంతుం ప్రాప్తిని కలిగి ఉంటుంది.



ఈ చర్చ అల్ప స్థాంతుం ప్రాప్తిని వాని చర్చా విధానం పరంగా వివరించవచ్చును.

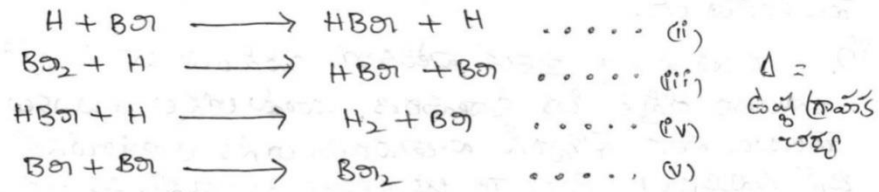
Stage I :- ప్రాథమిక ప్రక్రియ :

చర్చామిశ్రమంలోని ఒక B_{12} అణువు స్థాంతుం శక్తిని శోషించుకొని 'B' పరమాణువులుగా విచ్ఛేదనం చెందును.



Stage II :- ద్వితీయ ప్రక్రియ :

పాథమిక ప్రక్రియల విచ్ఛేదన 'B' పరమాణువుల చర్చా
మీశ్రమంల క్రింద తెలిపిన గొణ చర్చలను కలుగజేయవచ్చును.



$\Delta =$
ఉష్ణ గ్రాహక
చర్చ

H_2 మరియు O_2 ల మధ్య జరిగే కాంతి రసాయన చర్చవలె ఈ
చర్చలూ కూడా గొణ చర్చలు గ్రి (ii) & (iii) పలుమార్లు జరిగి ఈ కాంతి
రసాయన చర్చ కూడా అత్యధిక స్థాయిం పోయి వలగి ఉండవచ్చును.
కానీ ఇది అల్ప స్థాయిం పోయి కలిగి ఉండును. దీనికి కారణాన్ని ఇలా
వివరించవచ్చును.

గొణ చర్చ గ్రి ఉష్ణ గ్రాహక చర్చ. ఈ చర్చ జరుగుటకు అధిక ఉత్త-
పిత శక్తి అవసరం అవుతుంది. కనుక సాధారణ పరిస్థితుల వద్ద ఈ చర్చ
నెమ్మదిగా జరుగుతుంది. కనుక గొణ చర్చలు గ్రి (ii) & (iii) కూడా నెమ్మ-
దిగా జరుగును. కనుక ద్విత్వీయ ప్రక్రియలని ఈ నాలుగు చర్చలలో
చర్చ V ప్రధానంగా జరుగుతుంది. ఇది పాథమిక చర్చకు విరుద్ధ చర్చ.
అనగా పాథమిక చర్చల విచ్ఛేదన B పరమాణువుల ఉత్పన్నలను
విచ్ఛేదించుటకు బదులుగా తీరిగి యదార్థపు అణువును విచ్ఛేదించును. అం-
దుచేత ఈ కాంతి రసాయన చర్చ తక్కువ స్థాయిం పోయి కలిగి
ఉండును.

ఈ విధంగా ఈ కాంతి రసాయన చర్చ తక్కువ స్థాయిం పోయి
కలిగి ఉండుటకు ప్రధాన కారణం ద్విత్వీయ ప్రక్రియలని గొణ చర్చ
గ్రి ఉష్ణ గ్రాహక చర్చ కావడమే.

6. క్రింది వాటికి లభ్య వ్యాఖ్యలు (పాఠంబి)?

- 1) రసాయన బిష్టి
- 2) యతి బిష్టి
- 3) స్ఫుర బిష్టి

1. రసాయన బిష్టి :

సాధారణ ఉష్ణగ్రత వద్ద జరిగే ఒక రసాయన చర్చల వెల-
వడ కాంతిని "రసాయన బిష్టి" అని అంటారు. ఇటువంటి కాంతిని
"చల్లటి కాంతి" అని కూడా అంటారు.

క్రింది వాటిని రసాయన బిష్టికి ఉదాహరణలుగా గ్రహించవచ్చు.

- 1) మీనుగురు పురుగు ఉష్ణగ్రత చేయి కాంతి, రసాయన బిష్టికి ఒక
ఉదాహరణ. మీనుగురు పురుగు శరీరం నుండి "బ్లూ ఫ్లూయిడెన్స్" అనే

ఒక $proton$ వెలువడుతుంది. ఇది నాతావరణంగా $electron$ విస్తరించి సమయం అధికరణం చెందుతుంది. ఈ అధికరణ ప్రక్రియలు కాంతి వెలువడుతుంది.

6

i). కాంతి వేగంగా శక్తిగా వాటికలు కనిపించే అనుభవం కాంతి రసాయన దీప్తికి ఒక ఉదాహరణ. భూమి లోపల ఎలక్ట్రాన్ వీడ్చి తీయబడుతున్న ఫోటోన్ నాతావరణంగానికి ప్రవేశించినప్పుడు అధిక-జన్ సమయం P_{20} గా అధికరణం చెందును. ఈ అధికరణ ప్రక్రియలు కాంతి వెలువడుతుంది. దీనిని రసాయన దీప్తికి ఉదాహరణగా గ్రహించవచ్చు.

ii). సముద్రాలలో ఒక రకమైన "బెల్లీ చాప" చనిపోయి వీడ్చి తీయబడే ప్రక్రియలు కాంతి వెలువడుతుంది.

iv). ఒక రకమైన బాల్బ్ రియో కటిగిన కొద్దిగా వీడ్చి తీయబడే ప్రక్రియలు కూడా కాంతి వెలువడుతుంది.

2. ప్రతిదీప్తి :

ఒక పదార్థం కాంతిని శోషించుకొని వెలువెంటనే కొంత భాగాన్ని వీరు తరంగదైర్ఘ్యం గల కాంతి రూపంగా ఉద్ఘాటం చెందుతున్న "ప్రతిదీప్తి" అంటారు. ఈ దృగ్విషయాన్ని వేదవర్ణించు పదార్థాన్ని "ప్రతిదీప్తి జనకం" అంటారు. పదార్థం కాంతి శోషణం చెందుతున్న నేలపై పడిన వెలువెంటనే ఉద్ఘాటకాంతి కూడా అగిపోతుంది.

శోషక కాంతి తరంగదైర్ఘ్యం ఉద్ఘాటకాంతి తరంగదైర్ఘ్యం సమానంగా ఉంటే ఆ ప్రక్రియను "రెట్రో స్కాటర్లింగ్" అంటారు.

ప్రతిదీప్తికి కారణాలు :

పదార్థం కాంతి శోషణం చెందిన కాంతి శక్తి దానిలోని పరమాణువులు లేదా అణువుల నుండి లోలను తొలగించడానికి తగినంతగా తొలగించబడుతుంది. ఈ లోలను తొలగించడానికి స్థాయిలో నుండి అధిక శక్తి స్థాయిలోనికి ప్రవేశించును. ఈ లోలను తీరిగి భూస్థాయిలోనికి ప్రవేశించినప్పుడు అది శోషించుకొన్న శక్తిని కాంతి రూపంగా ఉద్ఘాటం చెందును.

ఉదా: యురేనియం, సెషియం, క్షిప్రిట్, Hg భాష్పాలు, బిర్రు, మొదలైనవి మొదలైనవి ప్రతిదీప్తిని వేదవర్ణించును.

ఉపయోగాలు :

1). అధిక పదార్థాల స్థితిని తెలుసుకొనుటకు.

2). సురిష్టుగులు (Ring waxes) ను గుర్తించుటకు.

3). Lightening ప్రక్రియలో ప్రతిదీప్తి ఉపయోగపడుతుంది.

3). స్ఫురదీప్తి :

ఒక పదార్థం కాంతిని శోషించుకొని దానిలో కొంతభాగాన్ని వెనువెంటనే ఉద్ఘాటం చేస్తుంది. అయితే పదార్థం మీద కాంతి పతనం చెందుటను నీలపుచేసిన తర్వాత కూడా ఇంతకాలం కాంతిని ఉద్ఘాటం చేస్తూనే ఉంటే ఆ పుష్కీషయాన్ని "స్ఫురదీప్తి" అంటారు.

"స్ఫురదీప్తిని నిదానంగా చురిగి యేదీప్తి అనికూడా అంటారు."

స్ఫురదీప్తికి కారణం :

ఒక పదార్థంలో శోషణం చెందిన కాంతిశక్తి, దానిలో రసాయన బంధాలను కలుగజేయడానికి తగినంతగా తీరేనపుడు అందు భూస్థాయిలలోనికి ప్రవేశించినపుడు అవి శోషించుకొన్న శక్తిని కాంతి రూపంలో ఉద్ఘాటం చేయును. అయితే అది అధిక స్థాయిల నుండి నిదానంగా భూస్థాయిలలోనికి ప్రవేశించును. క్లిష్టమైన నీర్మాణం కల్గిన పదార్థాలు ఈ పుష్కీషయాన్ని ప్రదర్శించును.

Ex :-

భారలోహ సల్ఫైడులను మలినాలుగా కల్గిన జింక్ సల్ఫైడ్, టైరలోహ మరియు టైరమ్యూక్రిక లోహ సల్ఫైడులు ఈ పుష్కీషయాన్ని ప్రదర్శించును.

ఉపయోగాలు :

4). ఎలక్ట్రిక్ స్క్రీన్ల మీద, వాచ్ డయల్స్ మీద painting చేయుటకు ఉపయోగించే Paints లో స్ఫురదీప్తి పదార్థాలను ఉపయోగిస్తారు.

కిరణజన్య సంయోగ క్రియ

పచ్చటి మొక్కలు CO_2 & H_2O ను ఉపయోగించుకొని క్లోరోఫిల్ మరియు గూర్భకాంతి సమక్షంలో కార్బోహైడ్రేట్లను తయారు చేసుకొను ప్రక్రియను "కిరణజన్య సంయోగ క్రియ" అంటారు. ఈ ప్రక్రియలో O_2 విడుదల అవుతుంది. కిరణజన్య సంయోగ క్రియ ప్రక్రియలో గూర్భకాంతి వికిరణ శక్తి చురుగుటకు రసాయన శక్తిగా మారుతుంది. ఈ కిరణజన్య సంయోగ క్రియ చురుగుటకు అవసరమైన కాంతి తరంగదైర్ఘ్యం "610 nm - 700 nm".

10
 కిరణజన్య సంయోగ క్రియలు CO_2 క్షయకరణం చెంది
 కాబోహైడ్రేట్లుగాను, నీరు బిల్డింగ్ కరణం చెంది బిల్డింగ్ గాను మారును.

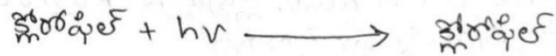
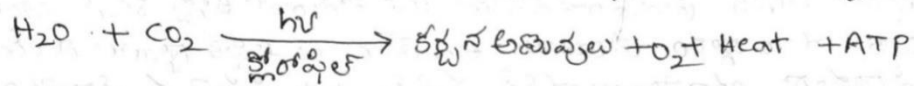
ఈ చర్యలకు జరుగుతున్న రెండు రకాల చర్యలు తోడ్పడును.

- 1). కాంతి చర్యలు
- 2). నిష్కాంతి చర్యలు

①. కాంతి చర్యలు :

క్లోరోఫిల్ లు ఉండే క్లోరోప్లాస్ట్ లు పెర తమమైన విధానంగా
 సూర్యకిరణాలను శోషించుకొని H_2O & CO_2 లకు అందించి తద్వారా
 కాబోహైడ్రేట్లను O_2 & ATP లను ఉత్పాదించుచున్నాయి.

కిరణజన్య సంయోగ క్రియను క్రింది విధంగా సూచించవచ్చును.



ఇక్కడ CO_2 గ్లూకోజ్ గానూ మరియు కాబోహైడ్రేట్ ల గానూ
 మారుతున్నాయి కావలసిన శక్తిని ATP అణువులు అందించును. పెర
 గ్లూకోజ్ అణువు సంశ్లేషణ చెందుతున్నప్పుడు 18 ATP అణువులు
 అవసరం.

HETERO CYCLIC COMPOUNDS

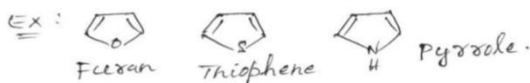
HETERO CYCLIC COMPOUNDS:

compounds which contain at least one atom other than carbon in their ring systems, are called heterocyclic compounds and the atom other than carbon that is involved in ring formation is known as hetero atom. The commonly found hetero atoms are N, O, S, P etc.

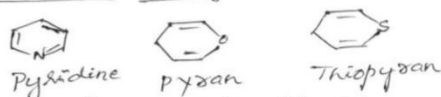
Ex: Furan, Pyrrole, furan etc.

Classification: Based on ring size heterocyclic compounds may be classified as

(i) Five membered rings:



(ii) Six membered rings:

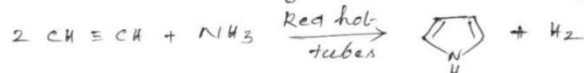


(iii) Fused heterocyclic rings:

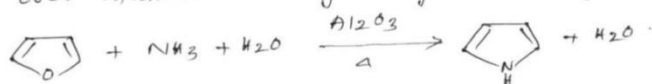


Preparation:

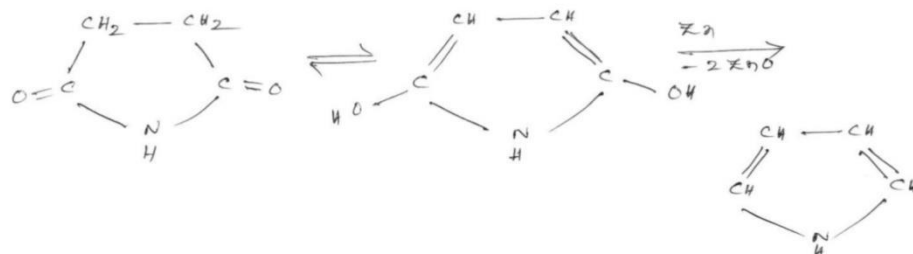
PYRROLE: (1) A mixture of acetylene and ammonia when passed through red hot tubes, pyrrole is formed.



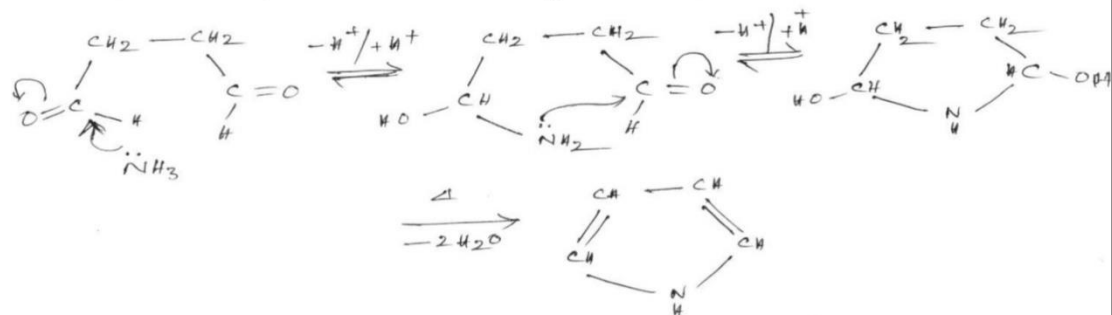
(2) A mixture of furan, ammonia and steam when passed over alumina at high temperature, pyrrole is formed.



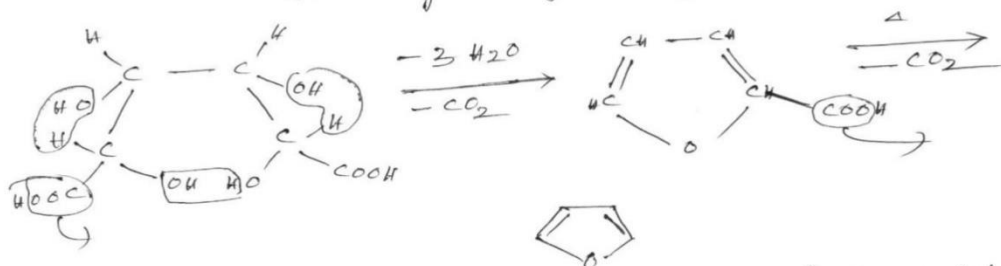
(3). Succinimide when heated with zinc dust forms pyrrole.



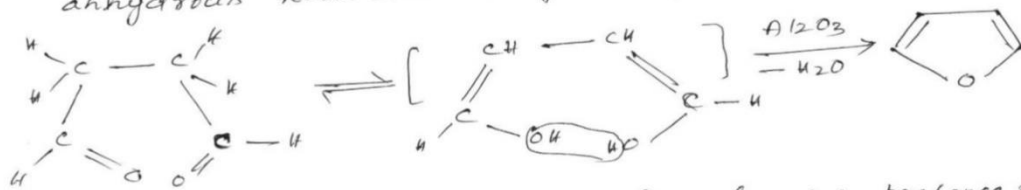
(4) By Paul-Knorr synthesis: 1,4-Dicarbonyl compound when heated with ammonia forms Pyrrrole.



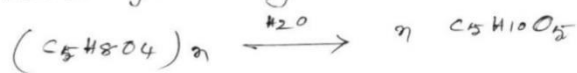
FURAN: (1) Mucic acid when subjected to dry distillation forms furic acid which on decarboxylation forms furan.

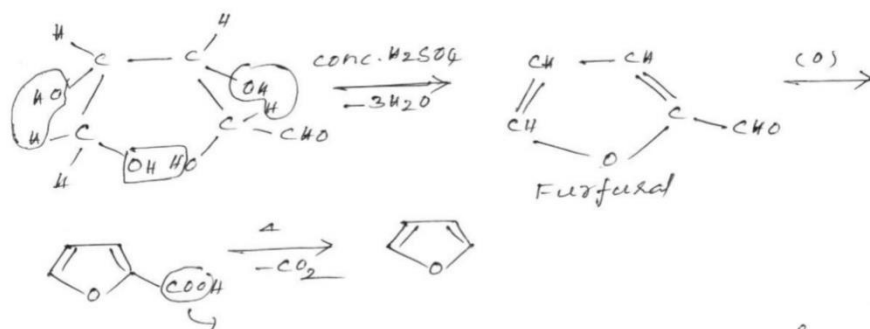


(2) By Paul-Knorr synthesis: Succinaldehyde, a 1,4-dicarbonyl compound when heated with Al_2O_3 (or) anhydrous zinc chloride forms furan.

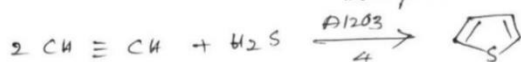


(3) A C_5 -Polysaccharide when hydrolysed in presence of an acid forms a pentose which on dehydration forms furfural. Furfural on oxidation followed by decarboxylation forms furan.

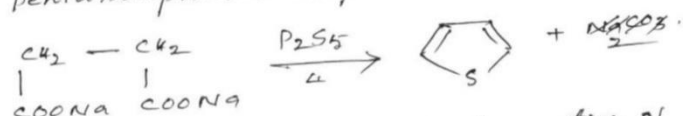




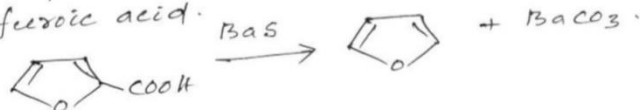
THIOPHENE : (1) A mixture of acetylene and hydrogen sulphide when passed over Al_2O_3 at high temperature, thiophene is formed.



(2). Sodium Succinate when heated with phosphorous-pentasulphide, thiophene is formed.



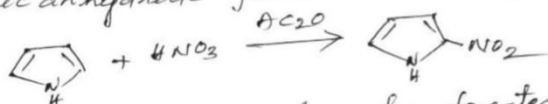
(3). Thiophene is prepared by the action of Bas on furoic acid.



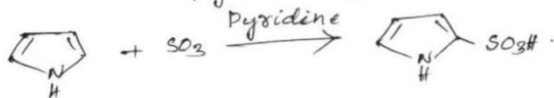
Properties :

PYRROLE : Pyrrole undergoes electrophilic substitution reactions preferably at C2.

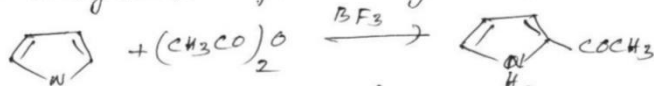
1. Nitration : Pyrrole when treated with HNO_3 in presence of acetic anhydride forms 2-nitropyrrole.



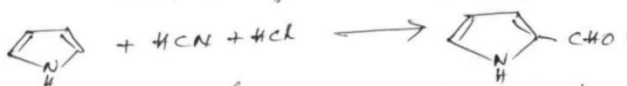
2. Sulphonation : Pyrrole when treated with SO_3 in presence of Pyridine forms Pyrrole-2-sulphonic acid.



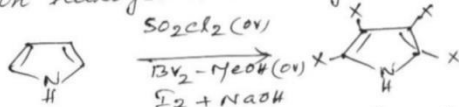
3. Friedel-Crafts acylation: Pyrrole when treated with acetic anhydride in presence of BF_3 forms 2-acetylpyrrole.



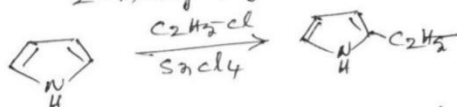
4. Formylation: Pyrrole when treated with a mixture of HCN and HCl forms 2-formylpyrrole.



5. Halogenation: Pyrrole is more reactive towards halogenation. on halogenation it gives 2,3,4,5-tetrahalopyrrole.

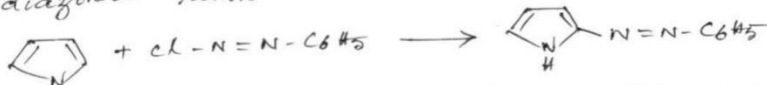


6. Alkylation: Pyrrole when treated with alkyl halide forms 2-Alkylpyrrole.

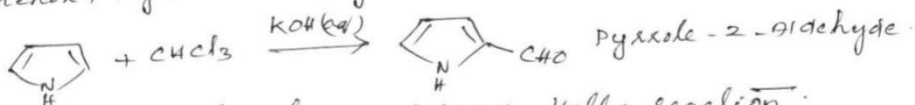


COMPARISON OF PYRROLE AND PHENOL:

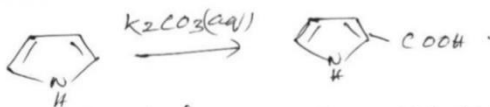
(i) Pyrrole like phenol undergo coupling reaction with diazonium salts.



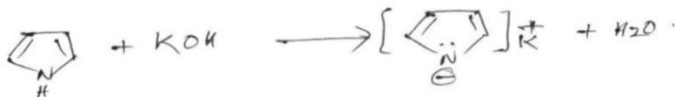
(ii) Phenol, like Pyrrole undergoes Reimer-Tiemann reaction



(iii) Pyrrole, like phenols undergoes Kolbe reaction.



(iv) Pyrrole behaves as a weak acid. It reacts with strong alkalis to form salts.



COMPARISON OF PYRROLE WITH AMINES :

5

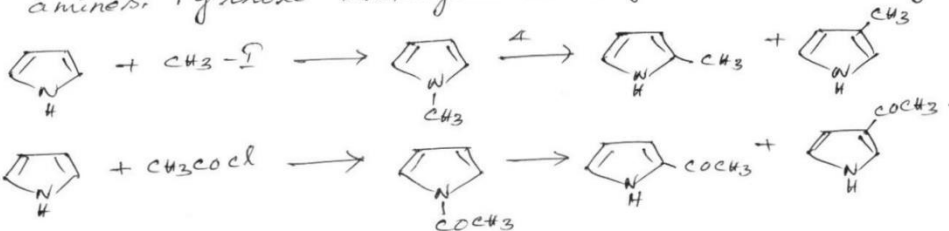
1. Like aromatic secondary amines, pyrrole is a weak base. It reacts with strong mineral acids to form salts.



2. Like 2°-amines, it reacts with nitrous acid to form N-nitroso amine.

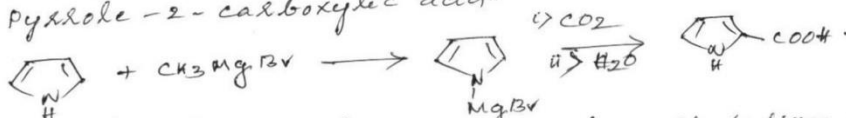


3. Like amines, pyrrole undergoes N-alkylation and N-acylation.

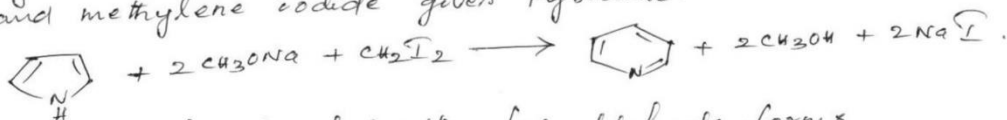


Other Properties of Pyrrole :

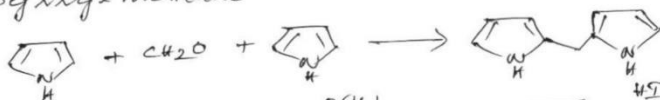
1. Reaction with Grignard reagent: Pyrrole on reaction with Grignard reagent forms pyrrole magnesium bromide which on reaction with CO_2 followed by hydrolysis forms pyrrole-2-carboxylic acid.



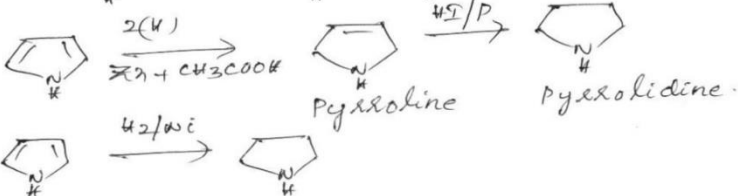
2. Pyrrole when treated with a mixture of sodium methoxide and methylene iodide gives pyridine.



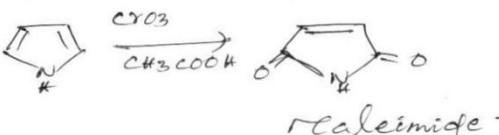
3. Pyrrole when treated with formaldehyde forms dipyrromethane.



4. Reduction :



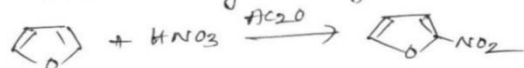
5. Oxidation :



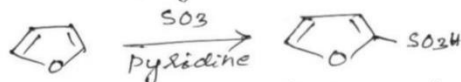
Properties of Furan :

6

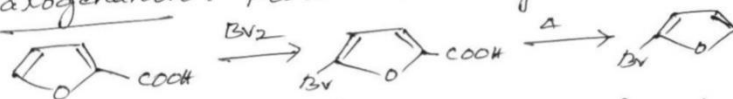
1. Nitration : Furan when treated nitric acid in presence of acetic anhydride forms 2-nitrofuran.



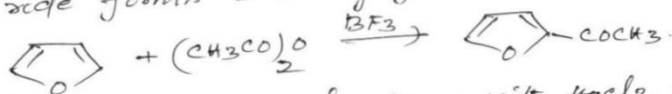
2. Sulphonation : Furan react with SO_3 in presence of pyridine to form Furan-2-sulphonic acid.



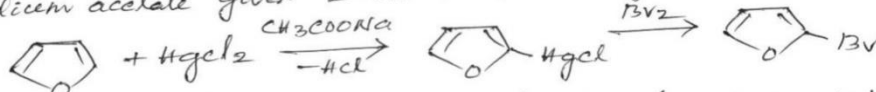
3. Halogenation : Furan is halogenated indirectly as follows



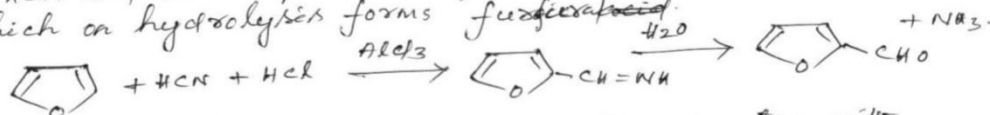
4. Friedel craft reaction : Furan when treated with acetic anhydride forms 2-acetyl furan.



5. Mercuriation : Furan on heating with HgCl_2 and aqueous sodium acetate gives 2-chloromercuric furan



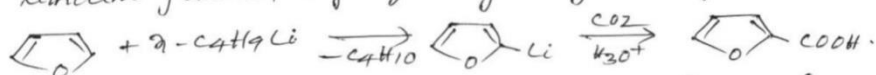
6. Gattermann reaction : Furan on treatment with a mixture of HCN and HCl in presence of AlCl_3 forms an imine which on hydrolysis forms furoic acid.



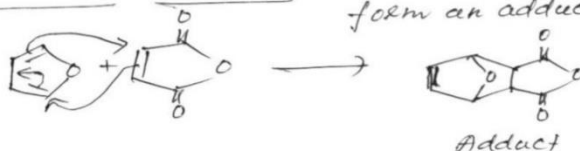
7. Gomberg reaction : undergoes coupling reaction with diazonium salts.



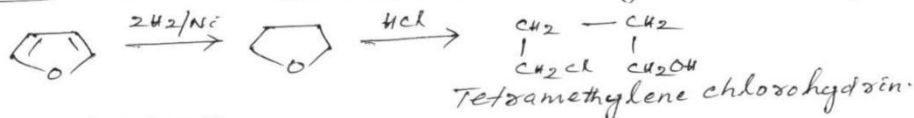
8. Reaction with alkyl lithium : Furan on reaction with n-butyl lithium followed by hydrolysis forms furoic acid.



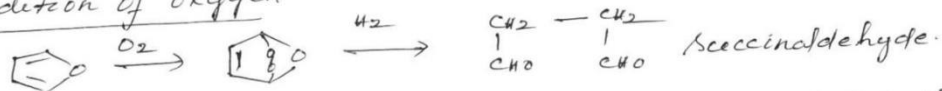
9. Diels-Alder reaction : It react with maleic anhydride to form an adduct.



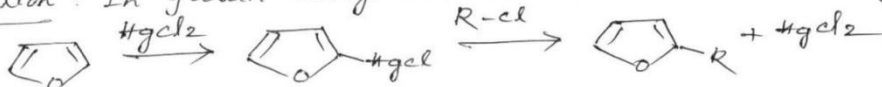
10. Reduction: Furan when reduced with Ni forms tetrahydrofuran.



11. Addition of oxygen

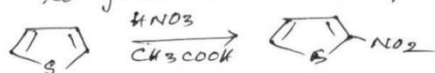


12. Alkylation: In furan alkylation is carried out indirectly

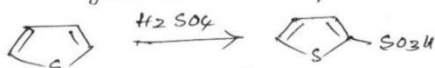


Properties of Thiophene:

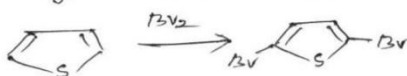
1. Nitration: Thiophene reacts with HNO_3 in presence of acetic acid to form 2-nitrothiophene.



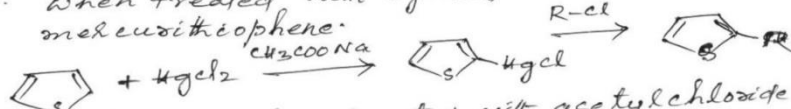
2. Sulphonation: Thiophene when treated with cold conc. H_2SO_4 forms 2-thiophene sulphonic acid



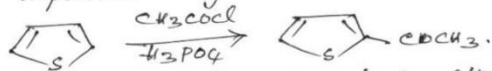
3. Halogenation:



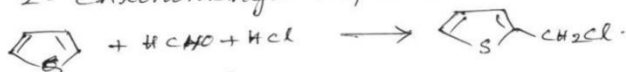
4. Mercuriation: When treated with HgCl_2 forms 2-chloro-mercuriothiophene.



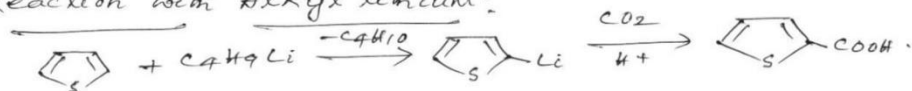
5. Friedel craft acylation: When treated with acetyl chloride in presence of H_3PO_4 forms 2-acetylthiophene.



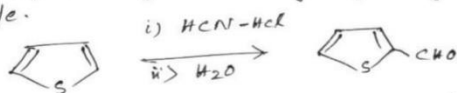
6. Chloromethylation: When treated with a mixture of HCHO and HCl forms 2-chloromethyl thiophene.



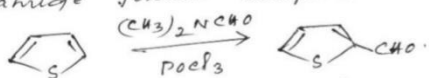
7. Reaction with Alkyl lithium:



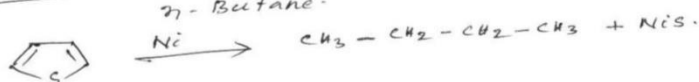
8. Gattermann reaction: Thiophene when treated with a mixture of HCN and HCl followed by hydrolysis gives thiophene-2-aldehyde.



9. Wittig reaction: Thiophene when treated with dimethylformamide forms thiophene-2-aldehyde.



10. Reduction: Thiophene when reduced with Raney Ni gives n-Butane.



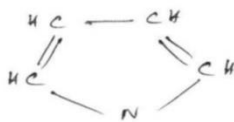
Structure Elucidation of Pyrrole:

- From analytical data the molecular formula of pyrrole was found to be $\text{C}_4\text{H}_5\text{N}$.
- Pyrrole add one two moles of hydrogen to form an addition product pyrrolidine.
- $\text{C}_4\text{H}_5\text{N} + 2\text{H}_2 \xrightarrow{\text{Ni}} \text{C}_4\text{H}_9\text{N}$

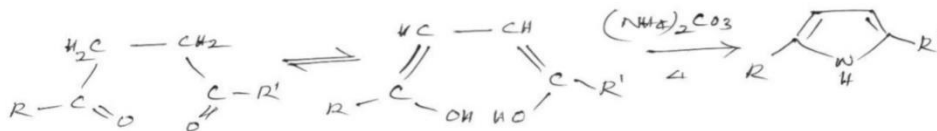
This reaction shows the presence of two double bonds in pyrrole.

- Pyrrole, like benzene undergoes electrophilic substitution reactions such as nitration, halogenation, sulphonation etc. These reactions indicate 'aromatic nature of pyrrole'.

From the above properties, the structure of pyrrole is shown as



- Confirmation by synthesis: The above structure of pyrrole is confirmed by its synthesis.



Limitations of diene structure: The above diene structure fails to explain the following properties.

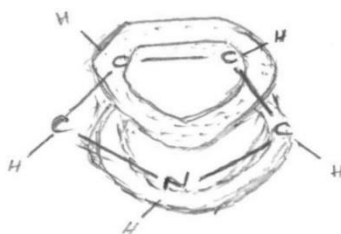
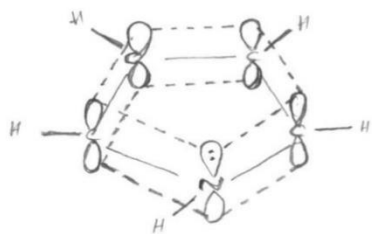
- They behave like benzene and undergo electrophilic substitution reactions.

- (ii) Like secondary amines, pyrrole lacks of basic character. 9
 (iii) Pyrrole is a flat molecule.

ORBITAL STRUCTURE: The above limitations are explained in terms of orbital structure of pyrrole.

In pyrrole all the four carbon atoms and nitrogen atom are sp^2 hybridised. These hybrid orbitals of 'C' and 'N' atoms form C-H, C-C, C-N & N-H sigma bonds.

The four carbon atoms contain ~~an~~ a unhybridised p-orbital containing an unpaired electron. Nitrogen atom contains a unhybridised orbital containing a lone pair of electrons. These unhybridised p-orbitals overlap sidewise to form a cyclic delocalised π -molecular orbital containing six electrons embracing all the five atoms of the ring. The delocalised orbital would make up two π -electron clouds, one above and one below the plane of the ring.



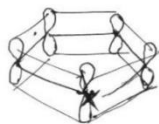
Explanation for the unique behaviour of Pyrrole:

1. As both the p-electrons on N-atom are involved in the formation of delocalised π -electron cloud, hence it lacks basic character.
2. The presence of electron cloud of six delocalised electrons makes the compound aromatic.
3. Pyrrole is a flat molecule due to involvement of sp^2 hybrid orbitals.
4. It undergoes addition reactions with difficulty, because the double bonds in pyrrole are not true double bonds due to resonance.
5. Pyrrole is a resonance hybrid of the following forms.



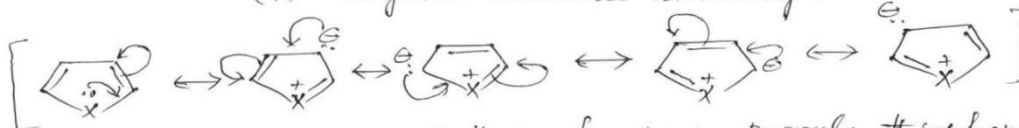
COMPARISON OF AROMATIC CHARACTER OF PYRROLE, FURAN & THIOPHENE

1. The three heterocyclic compounds are flat and planar rings. 10
2. Each one of these compounds possess 4π electrons and a lone pair of electrons on the hetero atom. These six electrons form an aromatic sextet.



$$X = N(0\pi) O(0\pi) S(0\pi)$$

- (3) The three heterocyclic compounds satisfy Huckel's rule i.e. they have $(4n+2)\pi$ electrons.
- (4) They have resonance stability.

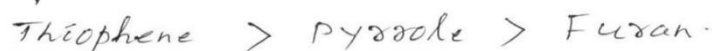


Because of all these factors, pyrrole, thiophene and furan exhibit aromatic character. But they have different degree of aromaticity.

Oxygen atom being more electronegative less readily contributes its lone pair of electrons in the formation of aromatic sextet. Hence furan is less aromatic than pyrrole and thiophene.

Sulphur being less electronegative, it can readily donate its lone pair of electrons in the formation of aromatic sextet. Hence thiophene is more aromatic than furan & pyrrole.

Therefore the order of aromatic character of these three heterocyclic compounds is shown as



COMPARISON OF ORDER OF REACTIVITY:

The reactivity of furan, pyrrole and thiophene follows the following order: $\text{Pyrrole} > \text{Furan} > \text{Thiophene} > \text{Benzene}$.

The reactivity order of the above three heterocyclic compounds depends upon the relative abilities of the hetero atoms to accommodate the positive charge in the intermediate carbonium ion.

Nitrogen atom accommodates the positive charge more readily than oxygen atom and sulphur atom. Hence pyrrole is more reactive than furan and thiophene.

The sulphur atom of thiophene though less electronegative than oxygen, its 3p-orbital overlap less efficiently

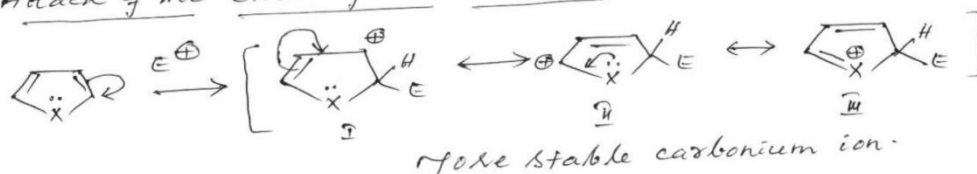
With the 2p-orbital of the aromatic π -electron system. Hence thiophene is less reactive than furan. //

Electrophilic Aromatic Substitution:

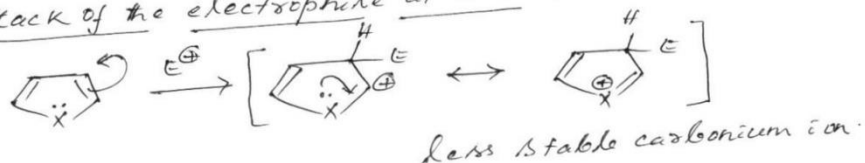
pyrrole, furan, thiophene are aromatic. Hence they undergo aromatic substitution reactions. They undergo electrophilic substitution predominantly in the position-2 than position-3 of the ring.

The carbonium ion formed from the attack of the electrophile at C-2 has more important resonance structures and therefore more stable than the carbocation resulting from attack at C-3. Hammond's postulate suggests that, the reaction involving the more stable intermediate will be faster.

Attack of the electrophile at C-2:



Attack of the electrophile at C-3:



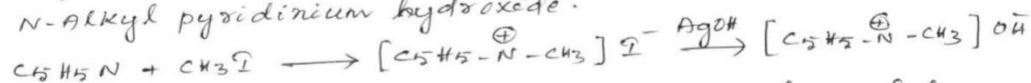
Therefore electrophilic substitution takes place at position-2 because it results in more stable carbocation.

PYRIDINE: Isolation from coal tar: Furan is present in the light oil fraction of coal tar. When light oil fraction is treated with $\text{dil. H}_2\text{SO}_4$, pyridine and all other basic substances pass into solution as their sulphates. Acid layer is separated and treated with NaOH solution to regenerate pyridine and other bases. Pyridine is obtained from the mixture of bases by fractional distillation.

Properties of Pyridine:

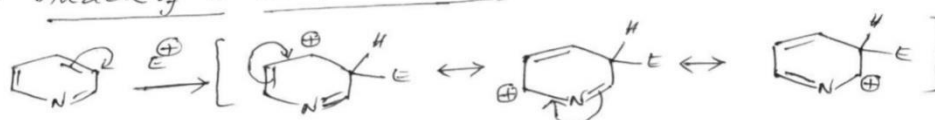
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1. Basic nature: It is a strong base. When treated with alkyl halide in presence of moist silver oxide yield N-alkyl pyridinium hydroxide.



2. Electrophilic Substitutions: It undergoes electrophilic substitutions preferably at position-3 because attack of electrophile at position-3 result in a more stable carbonium ion.

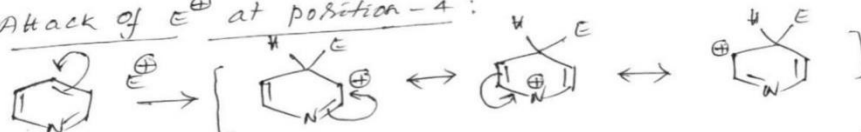
(i) Attack of E^+ at position-3



more stable carbonium ion

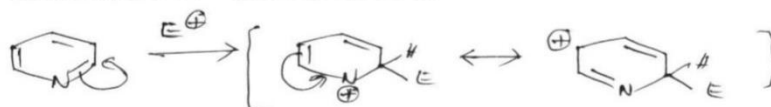
But attack of electrophile at position-4 (or) position-2 result in a less stable carbonium ion.

(ii) Attack of E^+ at position-4:



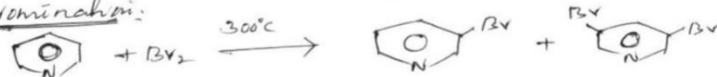
less stable carbonium ion because 'N' atom less readily accommodate +ve charge.

(iv) Attack of E^+ at position-2:

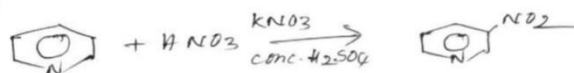


less stable carbonium ion because 'N' atom carries +ve charge.

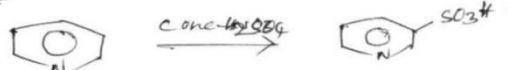
Example 1: Bromination:



Nitration:



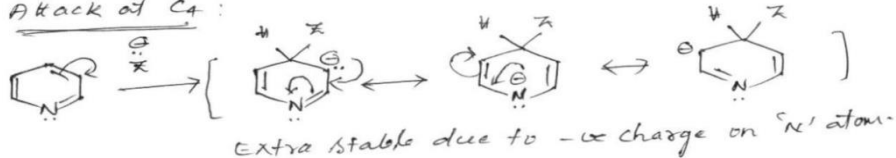
Sulphonation:



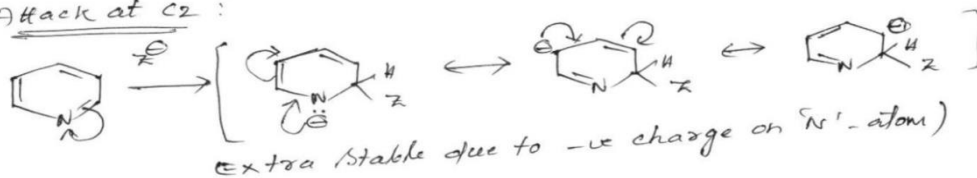
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Nucleophilic Substitution reactions: The electron withdrawing nature of nitrogen makes pyridine less reactive towards electrophilic substitutions and more reactive towards nucleophilic substitutions. Nucleophilic substitutions occur at positions 2 & 4 because the resultant carbanions are resonance stabilised. The negative charge on the nitrogen atom makes the carbanion extra stable.

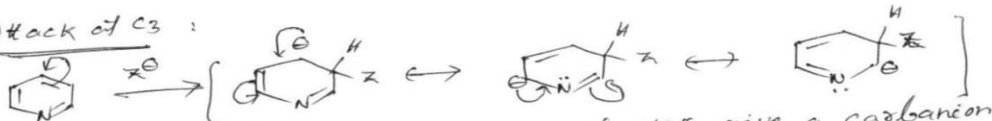
Attack at C4:



Attack at C2:

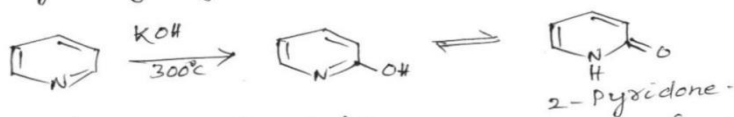


Attack at C3:

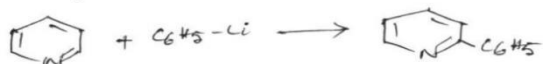


\therefore Attack of Nu^- at position -3 doesn't give a carbanion that carry $-ve$ charge on the nitrogen atom.

Examples: (1) Reaction with KOH: on heating with KOH 2-Hydroxy pyridine is formed.



(2) Reaction with Phenyl lithium: pyridine when treated with Phenyl lithium forms 2-Phenyl Pyridine.



3. Reaction with sodamide: pyridine when treated with sodamide forms 2-Aminopyridine. This reaction is called chichibabin reaction.



Mechanism:



4. Question bank

The respective class teachers provide the slow learners chapter wise important question bank which is useful an in-depth analysis of questions.

“OM NAMO VENKATESAYA”
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
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1. Unit wise important questions provided for slow learners

I B.Sc – I SEMESTER - PAPER – I
(INORANIC & ORGANIC CHEMISTRY)

UNIT – I

Chemistry of p-Block Elements – I

1. Define Borane. Write the synthesis, structure, properties and uses of diborane.
2. Write the synthesis and structure of boron-nitrogen compounds
3. What are silicones? How are they classified?
4. Write the preparation and reaction of
 - a. hydrazine
 - b. hydroxyl amine
5. What are silanes? Write two preparation methods of silanes.
6. Write the general characteristics of 13 group (Boron family) elements

UNIT – II

Chemistry of p-Block Elements – II

1. Write the classification of oxides based on
 - i. Chemical behaviour ‘
 - ii. Oxygen content
2. What are inter halogen compounds? How they are classified?
Give their general preparation
3. Write about pseudohalogens and psuedohalides and its uses?

Organo Metallic Chemistry

1. Write the Classification of Organo Metallic compounds.
2. Write the preparation, properties and applications of Alkyls of lithium
3. Write the preparation, properties and applications of Alkyls of Magnesium

UNIT – III

Structural Theory in Organic Chemistry

1. Write explanatory notes on:
 - i. Inductive effect
 - ii. Mesomeric effect?
2. What do you understand by the term resonance? What are the different conditions for writing resonance structures?
3. Use mesomeric effect to explain
 - a. Acidity of phenol
 - b. Acidity of carboxylic acid.
4. What is carbene? How we explain their stability?
5. What are the three major types of organic reactions? Explain giving examples in each case.
6. What types of bond fission possible for organic compounds?

UNIT – IV

Cyclic Hydro Carbons:

1. Suggest a possible reason why peroxide effect is for HBr and not for HCl.
2. Compare and explain the relative rates of addition HCl, HBr and HI to alkenes.
3. What do you know about Anti-Markovnikov effect?
4. Write a brief note on Diels-Alder reaction.
5. Why are conjugated dienes more stable than non-conjugated dienes?
6. Give two properties each of electrophilic and nucleophilic additions to acetylene along with their mechanisms.
7. Explain: Acetylene is more acidic than ethane.
8. Acetylene forms metal acetylides but dimethyl acetylene does not form such derivatives.

Cyclo Alkanes:

1. What factors make the cyclohexane ring the most stable, abundant and important ring structure in organic chemistry? Cyclopropane undergoes ring opening reactions. Explain why?
2. Cyclopropane undergoes ring opening reactions. Explain why?

3. How will you calculate the percentage of chair and boat forms?
4. What are applications and limitations of Baeyer's strain theory?

Benzene and its reactivity

1. State Huckel's rule. Illustrate with suitable examples.
2. Write the mechanism of nitration of benzene.
3. Discuss the orbital structure of benzene.
4. Describe the Friedel-Crafts acylation. Give its mechanism
5. Define Huckel's rule. What do you understand by the term aromaticity? Give examples.
6. State Huckel's rule and explain the stability of:
 - i. tropyllium cation
 - ii. Cyclopentadienyl anion.
 - iii. What is the cause of aromatic character of benzene in terms of Huckel's rule?
7. Discuss the mechanism of electrophilic substitution of benzene.
8. What is the cause of aromatic character of benzene in terms of Huckel's rule?
9. What is aromaticity?

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I B.Sc – II SEMESTER - PAPER – II
(PHYSICAL & GENERAL CHEMISTRY)

UNIT – I

Solid State:

1. Explain the laws of crystallography.
2. Define space lattice, lattice point and lattice plane.
3. What is Unit cell? Explain the types of unit cells.
4. What are Crystal systems? Explain.
5. Explain Bragg's equation.
6. State and Explain the law of symmetry.
7. What are Miller indices? Give the Miller indices of different planes in a simple cube.
8. Discuss about the application of Bragg's equation in studying the crystal structure.
9. Define Crystal defect. Explain the various types of crystal defects.

UNIT- II

Gaseous State:

1. Derive vander Waal's equation of state.
2. Define critical constants T_c , P_c and V_c
3. Derive the relationship between critical constants and van der Waal's constants.
4. Explain the law of corresponding state.
5. Explain the structural differences between solids, liquids and gases.

Liquid State:

6. Write the properties of liquid crystal.
7. Write the applications of liquid crystals.

UNIT – III

Solutions:

1. Explain Henry's law.
2. Discuss with examples, the solutions which show positive and negative deviation from Raoult's law.

3. Define critical solution temperature (CST). Write the classification of partially miscible liquid pairs depending on their CST values.
4. Discuss the principle of steam distillation.
5. Explain Nernst distribution law.

UNIT – IV

Colloids & Surface Chemistry:

1. Write the preparation of colloids.
2. How colloids can be purified?
3. Stability of colloids.
4. Write a short note on zeta potential.
5. What are the applications of colloids.
6. Electro osmosis
7. Derive the Langmuir adsorption isotherm and mention its applications.
8. Explain Freundlich adsorption isotherm.
9. Explain the factors influencing the adsorption of gases on solids.
10. Write the postulates of Valence bond theory.

Chemical Bonding:

1. Explain the valence bond theory.
2. Write a short note on Dipole movement.
3. Write molecular orbital diagrams for homo nuclear molecules, Nitrogen and oxygen
4. Write molecular orbital diagrams for hetero nuclear molecules, Carbon monoxide and nitric oxide.
5. Write the postulates of LCAO methods

UNIT – V

Stereo Chemistry of Carbon Compounds:

1. Write to molecular representations.
2. Write about stereo isomerism.
3. Define asymmetric (or) chiral carbon. What are the conditions required for a carbon compound to show optical isomerism?
4. What are the differences between enantiomers and diastereomers?
5. Write the optical isomers of a) glyceraldehydes b) lactic acid and c) alanine.
6. Discuss the optical isomerism exhibited by tartaric acid.
7. Explain R, S configuration with example.
8. Explain D, L configuration with examples.

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DEPARTMENT OF CHEMISTRY

**II B.Sc – III SEMESTER - PAPER – III
(INORGANIC & ORGANIC CHEMISTRY)**

UNIT – I

Chemistry of d-Block Elements – I

1. Write briefly about characteristics of d-block elements with special reference to electronic configuration
2. Write about oxidation states of d-block elements
3. Write about magnetic properties of d-block elements
4. Write about catalytic property and complex formation of d-block

Theories of Bonding in Metals

5. Explanation of properties of metals by valence bond theory
6. Write briefly about band theory of metals
7. Write a short note on free electron theory of metals?
8. Explain the properties of metals by VB Theory.

UNIT – II

Metal carbonyls and related compounds

1. What are metal carbonyls write about classification of metal carbonyls
2. Write the preparation, properties and structure of nickel tetra carbonyl
3. Write the preparation, properties and structure of iron penta carbonyl
4. Define and explain effective atomic number rule with examples
5. Write the preparation, properties and structure of chromium hexa carbonyl
6. Write the preparation, properties and structure of $[\text{Mn}_2\text{CO}_{10}]$
7. Write the preparation, properties and structure of $[\text{Fe}_2\text{CO}_9]$
8. Write the preparation, properties and structure of $[\text{Co}_2\text{CO}_8]$
9. Write the preparation, properties and structure of $[\text{Fe}_3\text{CO}_{12}]$

Chemistry of f-Block Elements – I

1. What are inner transition elements? Write the electronic configuration of lanthanides and actinides?

2. What are Lanthanides? What is meant by Lanthanide contraction? What is the cause of Lanthanide contraction?
3. What are the similarities and differences of Lanthanides and Actinides? Write the comparison of Lanthanides and Actinides.

UNIT -III

I. Halogen Compounds

1. SN^1 and SN^2 reaction mechanism with examples
2. Differences between SN^1 and SN^2 reactions

II. Hydroxy Compounds

3. Grignard synthesis of alcohols
4. *Special reactions of Phenols*
 - (i) Bromination
 - (ii) Azocoupling
 - (iii) Pinacol-Pinacoline rearrangement
 - (iv) Esterification

Name reactions

- (i) Kolbe – Schmidt
- (ii) Riemer – Tiemann reaction
- (iii) Fries rearrangement
- (iv) Schotten Baumann reaction

UNIT - IV

Carbonyl Compounds

5. Nucleophilic addition reactions
6. (i) Aldol condensation
 - (ii) Cannizzaro's reaction
 - (iii) Perkin reaction
 - (iv) Benzoin condensation
 - (v) Haloform reaction
 - (vi) Knoevenagel reaction
 - (vii) Baeyer-villager oxidation
 - (viii) Clemmenson reduction, wolf-Kishner reduction,

MPV reduction, reduction with LiH_4 and NaBH_4 .

7. 2, 4 DNP Test, Tollen's test, Fehling test, Schiff's test and Haloform test

UNIT - V

Carboxylic acids and derivatives

8. (i) Hydrolysis of esters (ii) Carbonation of Grignard reagents

9. (i) Hydrogen bonding (ii) Relative differences in the acidities of aromatic and aliphatic acids

10. Huns Diecker reaction, Decarbonylation by Schmidt reaction, Arndt – Eastert synthesis, Hell-Vollhard-Zelinsky reaction

Active Methylene Compounds.

(i) Malonic ester and Acetoacetic ester preparation and Synthetic applications

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II B.Sc – IV SEMESTER - PAPER – IV

SPECTROSCOPY AND PHYSICAL CHEMISTRY

UNIT – I

Electronic Spectroscopy :

1. Explain the difference types of Molecular Spectra
2. Write the difference types of electronic transition
3. Write a short note on Chromophore auxochrome

Infrared Spectroscopy:

1. Write a note on modes of vibrations in diatomic molecules
2. Write the interpretation of Spectra-Alcohols

Proton Magnetic Resonance Spectroscopy

1. Write a note on equivalent and non equivalent protons
2. Explain the principles nuclear magnetic resonance spectroscopy
3. Explain the chemical shift with examples
4. Write a short on the following
 - a. Splitting of signals
 - b. Spin spin coupling
 - c. Coupling constant
5. Explain the NMR data for acetaldehyde and ethyl acetate

UNIT – II

Spectrophotometry

1. Write a short on the following
 - a. Transmittance
 - b. Absorbance
 - c. Molar Absorptivity
2. Explain single beam spectrophotometer
3. Explain double beam spectrophotometer

4. Explain Berlambeer-lamberts law and its limitations
5. Explain the application of beer lambers law for quantitative analysis of chromium in potassium dichromate solution
6. Explain the application of beer lambers law for quantitative analysis of chromium in Manganese in mangamous sulphate

UNIT –III

Dilute Solutions :

1. Explain roults law for lowering of vapour pressure how it is useful for determination of the molecular weight of a non volatile solute
2. What are the colligative properties explain each of them with suitable examples
3. Described cottrele’s method of determining molecular weight of non volatile solute by using elevation of boiling point
4. What is osmosis define osmotic pressure, explain the determination method for osmotic pressure
5. How is molecular weight of a non volatile solute determined by using Ostwald dynamic method on basis of relative lowering
6. Define an ideal solution and dilute solutions
7. Explain vanthoffs laws of osmotic pressure
8. Write a short note on
 - a. Semi permeable membrane
 - b. Abnormal colligative property

UNIT – IV

Electro Chemistry - I

1. What are transport numbers, describe hittarfs method for the determination of transport number
2. State and explain kohlrauschs law
3. What is Conductometric Titrations, explain various types of Conductometric Titrations
4. Write a short note on
 - a. Specific resistance
 - b. Specific Conductance

- c. Equi valent conductance
- d. Molar conductance
- 5. Explain Debye-Huckel-Onsagars theory of strong electrolytes
- 6. Discuss about Nernst equations
- 7. State and explain Ostwald's dilution law
- 8. Explain the reversible and irreversible cells
- 9. Explain Arhenius theory of electrolytes
- 10. How do you determine the EMF of a cell write any one application of EMF measurements

UNIT – V

Electro Chemistry – II

- 1. Discuss about Potentiometric titrations and give their advantages
- 2. What single electrode potential write the construction and working of standard hydrogen electrode
- 3. How Potentiometric titrations used to find the P^H of acid solution
- 4. What is standard electrode potential explain calomel electrode with cell reactions
- 5. Write the principles and applications of indicator electrode
- 6. Discuss about electrolytes

Phase Rule

- 1. State and explain the gibbs phase rules , explain the phase diagram of Ag-Pb system
- 2. Explain the phase diagram of single component system
- 3. Explain the phase diagram of solid liquid system
- 4. Explain the phase diagram of Magnesium-Zinc system
- 5. Write short note on
 - a. Freezing mixture
 - b. Triple point
 - c. Eutectic point
 - d. Terms in phase rule
 - e. Condensed system
 - f. In congruent melting point

S.G.S. ARTS COLLEGE : TIRUCHANOOR ROAD : TIRUPATI

DEPARTMENT OF CHEMISTRY

III B.Sc – V SEMESTER - PAPER – V

UNIT - I

Coordination Chemistry

1. Explain Warner's theory of complex compounds
2. Explain valence bond theory in complex compounds
3. Explain high spin and low spin complexes with an example each
4. Discuss the salient features of crystal field theory
5. Write notes on the stereoisomerism of complex compounds with co-ordination number 4.
6. Explain the EAN rule. Give any two complexes which do not obey this rule.
7. Explain the formation of the following complexes basing on valence Bond theory.
 - i. $[\text{CoF}_6]^{3-}$
 - ii. $[\text{Zn}(\text{HN}_3)_4]^{2+}$
 - iii. $[\text{Ni}(\text{CO})_4]$
 - iv. $[\text{Cu}(\text{NH}_3)_4]$
8. What is meant by Crystal Field Stabilisation Energy
9. Define coordination number in a complex. What is the coordination number of
 - i. Cobalt in $[\text{Co}(\text{NH}_3)_4\text{ClBr}]$
 - ii. Nickel in $[\text{Ni}(\text{DMG})_2]$
10. What is ligand? Give examples
11. Calculate the EAN of the central Atom the following:
 - i. $[\text{Co}(\text{NH}_3)_6]^{3+}$
 - ii. $[\text{Cr}(\text{CO})_6]$
 - iii. $\text{K}_6[\text{Fe}(\text{CN})_6]$
12. Apply the concept of EAN rule on the complexes $[\text{Fe}(\text{CN})_6]^{4+}$ and $[\text{Fe}(\text{CN})_6]^{3-}$. What conclusion you draw about the validity of the EAN rule from these two well-known complexes?

UNIT – II

Special and magnetic properties of metal complexes stability of metal complexes

1. Describe how magnetic moment can be determined by Gouy's balance
2. What is stability constant? What are the factors that determine the stability of a complex
3. Describe Job's method of determining the composition of a metal complex.
4. Explain the term Chelates.
5. Explain the kinetic stability and thermo-dynamic stability of a complex.

UNIT – III

Nitro Hydrocarbons

1. Distinguish between alkyl nitrites and nitroalkanes.
2. How is nitrobenzene prepared? Give its reducing properties
3. Write note on:
 - a. Mannich reaction
 - b. Michael addition
 - c. Michael Reduction
4. Explain Nef reaction
5. Explain the acidic nature of nitro alkanes.

UNIT IV

Nitrogen Compounds

1. a. What are amines" Give general methods of preparation of primary amines.
b. Give a brief account of the properties of amines
2. Describe a method of separation of primary, secondary and tertiary amines in a mixture.
(Hinsberg Method)
3. Write about the Basic characters of amine
4. Write about Hoffman's bromamide reaction
5. Write about the aromatic electrophilic substitution.

UNIT V

Thermodynamic

1. a. State the first law of thermodynamics in its different forms. Give the mathematical expression of the law.
b. Calculate the work done in calories when 3 moles of an ideal gas at 25°C expands isothermally and reversibly from a volume of 10 litres to a volume of 50 litres.
2. Derive the relationship between $H\Delta$ & $E\Delta$
3. Explain Joule –Thomson effect.
4. Distinguish between thermodynamically Reversible and Irreversible process.
5. a. Derive an expression for the work done in the reversible isothermal expansion of an ideal gas.
b. Calculate the work done when 10gms. of an ideal gas of molecular weight 44 expands isothermally and reversibly from a volume of 5 litres to 10 litres at 27°C .
6. Which of the following is function of state. Why?
 - i. Internal energy.
 - ii. Work done by a gas.
7. What is a spontaneous process? Explain with examples.
8. State and explain Carnot theorem

V YEAR – SEMESTER-V – PAPER-VI

UNIT - I

Reactivity of Metal Complexes

1. What are labile and inert complexes? Explain with necessary examples.
2. How many types are nucleophilic ligand substitution reaction in complexes? Explain them with examples and necessary mechanism.
3. What are the theories and application of Trans effect in metal complexes.
4. Name the factors which influence the rate of square planar substitution reaction.

Bio-inorganic Chemistry

1. What are the elements required for the body? Name them give their percentages.
2. Write about the structure and functions of hemoglobin.
3. What is the role of Mg^{2+} ions in biochemical processes and in photosynthesis?
4. What are metal porphyrins?

UNIT – II

Chemical Kinetics

1. Derive the equations for the specific rate 'k' of a second order reactions $2A \rightarrow P$
2. What is Arrhenius equation? Explain the concept of activation energy of reaction.
3. Give the integrated equation for the rate constant of a first order reaction. Give two examples of first order reactions.
4. What is half-time ($t_{1/2}$) of a reaction? Give equations for its value for first and second order reactions.

Photo Chemistry

1. What is Grothus-Draper law? Explain with examples
2. State and explain Stark Einstein's law of photo chemical equivalence.
3. What are the factors that influence the ϕ value? Explain in detail the reasons for higher values.
4. What are fluorescence and phosphorescence? Explain with the help of Jablonsky diagram.

5. What is Jablonsky diagram? Illustrate it briefly.
6. What is Grothus-Draper law?
7. What is Stark-Einstein's law of phot chemical equivalence? Explain.
8. What is fluorescence? Ecplain
9. What is phosphorescence? Explain.
10. What are photo-sensitized reaction? Give examples.

UNIT – III

Heterocyclic Compounds

1. Write about the aromatic character of pyrrole, furan and thiophene in detail.
2. Discuss in detail the acidic character of pyrrole.
3. Explian Diels-Alder reaction of furan with an example.
4. Give two methods of preparation of pyridine with equations.
5. What are heterocyclic compounds? Give examples.
6. Explain the aromatic character of pyrrole.
7. Explain the aromatic character of furan.

UNIT - IV

Carbohydrates

1. What are carbohydrates? How are they classified? Give examples.
2. Discuss the cyclic structure of glucose.
3. Discuss the structure of Fructose.
4. How are the following conversions effected?
 - i. Aldo hexose to Aldo pertose
 - ii. Aldo pentose to Aldo hexose
 - iii. Keto hexose to Also hexose
 - iv. Also hexose to Keto hexose
5. What are carbohydrates? Give two examples
6. What is mutarotation? Explain.
7. How do you prove that there are 5 —OH groups in Glucose? Explain.

UNIT – V

Amino Acids and Proteins

1. What are amino acids? How are they classified? Give examples.
2. Discuss the structure of proteins with necessary explanations.
3. How are amino acids separated from proteins and identified?
4. How are amino acids prepared from α -halogen acids?
5. What is Zwitter ion? Explain.
6. What are the different structures of protein?

S.G.S. ARTS COLLEGE : TIRUCHANOUR ROAD : TIRUPATI

DEPARTMENT OF CHEMISTRY
III B.Sc – VI SEMESTER - PAPER – VI(A)

ANALYTICAL METHODS IN CHEMISTRY

UNIT - I

Quantitative Analysis:

1. Write the principles of Volumetric analysis.
2. Write the Theories of Acid base Titrations.
3. Write the Theories of Redox Titrations.
4. Write the Theories of Complexometric Titrations.
5. Write the Theories of Iodometric Titrations.
6. Write the Theories of precipitation Titrations.
7. Write a short note on
 - a. Indicators
 - b. Precipitation
 - c. Coagulation
 - d. Peptization
 - e. Co-precipitation
 - f. Post precipitation
 - g. Digestion
 - h. Filtration
 - i. Drying and Ignition
8. Explain the steps involved in Chemical Analysis

UNIT – II

Treatment of Analytical Data:

1. Write the types of errors.
2. Write a short note on
 - a. Significant figures.
 - b. Accuracy.
 - c. Precision.
 - d. Standard Deviation.
 - e. Confidence Limit.

UNIT – III

Separation Techniques of Chemical Analysis

1. Explain solvent extraction.
2. Explain batch extraction.
3. Explain continuation extraction.
4. Explain counter current extraction.
5. Write a short note on synergism.
6. Write the action of Ion exchange resins.
7. Write the separation of Inorganic mixtures.

UNIT – IV

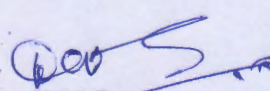
Chromatography

1. Give the principle involved in chromatography.
2. Define the term 'Eluent'
3. What do you understand by the term ' R_f value'?
4. What is the advantage of descending paper chromatography over ascending paper chromatography?
5. What do you understand by the term 'Retention factor'? How does it help in the identification of components in a mixture?
6. Explain the term 'Paper' and how it acts as a chromatographic medium in paper chromatography.
7. a. Explain the principle of column chromatography.
b. How will you proceed to separate a mixture of glycine and alanine?

UNIT – IV

Thin layer Chromatography

1. Write the basic principles and application of TLC.
2. Write the basic principles and application of HPLC
3. Write a short note on R_f value.
4. Adsorbents.
5. Detection of the spots.
6. Preparation of plates.
7. Development of Chromatogram


HEAD
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S. G. S. ARTS COLLEGE
TIRUPATHI

5. Additional assignments and Tests

As a part of remedial work the students are given additional tests and assignments to reinforce the concept learnt in the class and improve performance in the university examinations.

Tests conducted for slow learners

Class test for slow learners conducted on 23-12-2019

SEMESTER-II

Max Marks : 25

Section – A

Answer any one question (1x5=5 Marks)

1. Explain shottky and Frenkel defects.
2. Write notes on Tyndall effect.

Section – B

Answer all Questions (2x10=20 Marks)

3. a. Derive Bragg's equation. Discuss its application in studying NaCl crystal structure.
or
b. Write a brief note on Langmuir adsorption isotherm.
4. a. What are differences between solid and liquid crystals. Write the application of liquid crystals.
or
b. Discuss the application of Nernst Distribution Law.

S.G.S ARTS COLLEGE: TIRUCHANOOR ROAD: TIRUPATI :

DEPARTMENT OF CHEMISTRY

ACADEMIC YEAR: 2019-20

Tests conducted for slow learners

Class test for slow learners conducted on 02-03-2020

SEMESTER-IV

Max Marks : 25

Section – A

Answer any one question (1x5=5 Marks)

1. Discuss selection rules of IR spectra.
2. Write the basic principle of NMR spectroscopy.

Section – B

Answer all Questions (2x10=20 Marks)

3. a. Explain different modules of vibrations of non linear molecules in IR with two examples.

or

- b. Write a brief note on NMR splitting of signals.
4. a. i) Explain the finger print region of IR spectra.
ii) Explain the IR spectra of an harmonic oscillator.

or

- b. What is chemical shift? How to measure it? What are the factors affecting chemical shift.

S.G.S ARTS COLLEGE: TIRUCHANOOR ROAD: TIRUPATI :

DEPARTMENT OF CHEMISTRY

ACADEMIC YEAR: 2019-20

Tests conducted for slow learners

Class test for slow learners conducted on 02-03-2020

SEMESTER-VI

Max Marks : 25

Section – A

Answer any one question (1x5=5 Marks)

1. Briefly explain a. Precipitation and b. Co-precipitations
2. Explain the theories of Acid base Titrations.

Section – B

Answer all Questions (2x10=20 Marks)

3. a. Write the classification of errors.

or

b. Explain the following i) Accuracy ii) Precision and iii) Standard deviation

4. a. Write the principles of Volumetric analysis.

or

b. Write brief notes on i) Indicators ii) Coagulation

Semester: II

 S.A. ARTS COLLEGE / S.P.M. COLLEGE / S.S.S. ARTS COLLEGE Name of the Institution: _____ Date: 23/11/2023 Name of the Student: A. Srinivasan _____ Roll No: 2023 Name of the Teacher: Dr. P. Srinivasan _____ Page: 1 of 1 Teacher's Signature: _____	
1.	Answer any one question (100% correct)
2.	Explain crystal and liquid states
3.	Write note on liquid crystal
4.	Answer all questions (20% = 20 marks)
5.	a. Explain Gay's equation, discuss its application in crystal and liquid states
6.	b. Write short note on liquid crystal
7.	c. What are differences between solid and liquid crystal, write one application of liquid crystal
8.	d. Discuss the application of van der Waals law
9.	<u>Liquid crystal</u> :-
10.	Liquid is a state of matter that when a beam of light is passed through a solution it cannot be seen unless it is placed directly in the line of vision of the observer.
11.	The liquid crystal is only in the solid state of light from the surface of crystal

4 - The appearance of dust particles in a semi-darkened room, when a sun beam enters are the well known examples of Tyndall effect. The dust particles are large enough to scatter light which renders the path of light visible.

∴ There are following the of particles of character of the particles of the dispersed phase must be no much smaller than wavelength of the light used. The latter do not exhibit the Tyndall effect.

ii. The difference b/w the refractive indices of the dispersed and the dispersion medium must be appreciably high. Colours have 'is lyophobic sol' the difference of refractive indices is very small and hence the Tyndall effect is very weak.

Siedentz and Zsigmondy applied the principles of Tyndall effect in designing the ultra-microscope.

ultra-microscope:- It shines by white light from an arc lamp powerful source is condensed by a system of convex and concave lenses and focused through with the help of a cell.

Rayleigh equation:- According to Rayleigh's equation could be employed not only as transmission grating but also as a reflection grating, when X-rays, an incident on a crystal face, this

in successive planes. called on n and n' . The angle of refraction between the wavefronts (2) The distance (2) between the successive atomic planes and the angle of refraction. The relation is known as Bragg's equation.

Fig 13.10 depicts a diffraction of monochromatic x-rays striking the parallel equidistant planes of atoms (or crystal) with interplanar distance d . The incident wavefronts in plane 1 at C_1 striking the successive planes, there are reflected such that angle of refraction angle at incidence θ .



Fig. 13.10 Reflection of monochromatic x-rays by a crystal.

$$n\lambda = C_2 + 2D$$

where $n, \lambda = 2$, whole numbers.

Since the triangles OC_2 and ODC_2 are congruent

$$C_2 = 2D$$

$$n\lambda = 2C_2$$

Now, in the right angle triangle OC_2

$$\sin \theta = \frac{C_2}{D_2}$$

$$C_2 = D_2 \sin \theta$$

$$C_2 = D \sin \theta$$

$$C_2 = d \sin \theta$$

Substituting the value of C_2 from equation

$$n\lambda = 2d \sin \theta$$

$$n = 1, 2, 3 \dots$$

where

The equation is known as Bragg's equation. $n=1$ is called first order reflection. The reflection corresponding to $n=2$ is the second order reflection and so on. Thus by measuring θ and the incident angle θ , we can know d .

$$d/d = \frac{a}{\sin \theta}$$


from this d can be calculated if λ is known and vice versa. In x-ray reflection ' n ' is usually set as equal to 1. Thus, Bragg's equation can alternatively be written as

$$\lambda = 2d \sin \theta$$

where d is the perpendicular distance between adjacent planes with indices h, k, l .

For experimental arrangement the opposite way, employed is Bragg's x-ray spectrometer and determination is done by this method. Bragg's method is by passing monochromatic beam so details about the Bragg's method are described below. A beam of x-ray is generated from x-ray tube.

Semester-IV

 Tirumala Tirupati Devasthanams Degree & PG Colleges, Tirupati
S.V. ARTS COLLEGE / S.P.W. COLLEGE / S.G.S. ARTS COLLEGE

Name of the Examination: Revised class test Date: 2-3-2019
Name of the Student: D. Sai Kumar Roll No.: 8269
Dept: Scs Grad: BSc Medium: _____ Subject: Chemistry
Paper: II Q. Code: _____

No. of Additional Sheets used: ☐

Questions

Section - A

1. Write the basic principles of NMR spectroscopy.
2. Discuss selection rules of IR spectra.

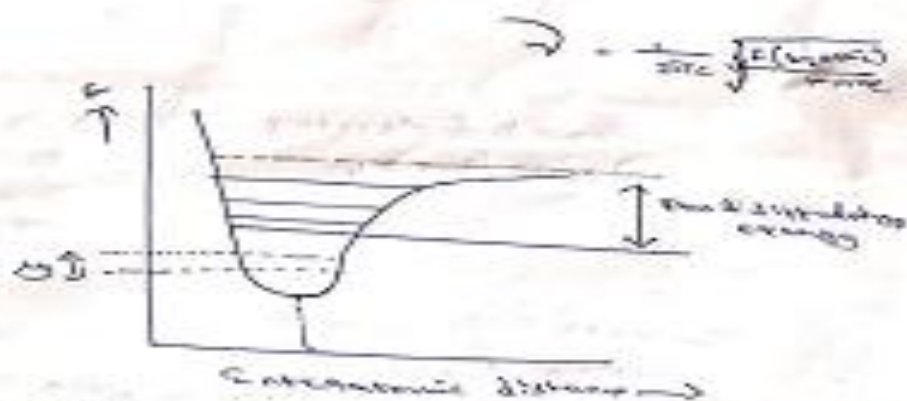
Section - B

3. Differentiate wavenumbers & frequencies of molecules in IR with two examples.
(5)
4. Write brief note on NMR splitting of signals.
5. Explain the fingerprint region of IR spectra.
a. harmonic oscillator.
(5)

Chemical shift? how to measure it? factors affecting chemical shift.

Answers:

1. Rules of IR spectra! - The selection rule says that vibrations are only IR active (or allowed) if the molecule dipole moment changes during the vibration. A diatomic molecule with the same atom cannot be excited to vibrate because no dipole moment is present.



eg: CO₂

① Degenerated vibrations: Symmetrically equivalent vibrations having same wave number

eg: CO₂ bending vibrations

2. Coupled vibrations: Bands with slightly more or less splitting can occur due to the rotation of molecules in a structure.

3. Overtone bands: Due to combination of 2 $h\nu$, 3 $h\nu$ etc. lower intensity than the fundamental vibration band.

4. Combination vibrations: Sum of difference of two vibrations. They can be helpful to determine the expectation & molecule is couple.

9. harmonic oscillator: In classical mechanics a harmonic oscillator is a system that when displaced from its equilibrium position experiences a restoring force. F proportional to the displacement x .
 $F = -kx$
 where k is a positive constant.

→ If F is the only force acting on the system the system is called simple harmonic oscillator. It undergoes simple harmonic motion. Simple harmonic oscillation about the equilibrium point with a constant amplitude & a constant frequency (which doesn't depend on the amplitude).

→ Oscillate with a frequency lower than F the undamped oscillator (undamped oscillator).

→ Decay to the equilibrium position, with oscillations (overdamped oscillator).

→ It can explain examples include pendulums (with small angles of displacement) masses connected to spring & oscillated by

→ Day is an external time-dependent force & present the harmonic oscillator is described as a system.

$$\Rightarrow \frac{d^2\theta}{dt^2} + \frac{g}{L} \sin\theta = 0$$

or $\theta \approx 0$, indeed condition the equation

$$\frac{d^2\theta}{dt^2} + \frac{g}{L} \theta = 0$$

$$\theta(t) = r \cos(\sqrt{\frac{g}{L}} t + \phi)$$

③ Non-linear molecule!

Acetylene, carbon dioxide, hydrogen cyanide etc... are some examples of linear molecule whenever branched & cross linked polymer such as vulcanized rubber are examples of non-linear molecules.

⇒ The key difference between linear non-linear molecules is that linear molecules have a chemical structure.

⇒ non linear molecule have a zig-zag or cross-linked chemical structure.

* All the molecules we know can be divided into two types as linear and non-linear depending on the shape of the molecule. If the chemical structure of a molecule has linear geometry which appears to be a straight line.

* All theory molecules are categorized as non-linear molecules.

Semester-VI

14
25

Tirumala Tirupati Devasthanams Degree & PG Colleges, Tirupati.
S.V. ARTS COLLEGE / S.P.W. COLLEGE / S.G.S. ARTS COLLEGE

Name of the Examination: Remedial class test Date: 22-03-20
Name of the Student: C. Ananasekhare Roll No.: 8201
Class: VI Sem Group: BBC Medium: ET Subject: Chemistry
Paper: VI
No. of Additional sheets used: ☐ Q. Code: _____

Section-A

Answer any One question (1x5 = 5 Marks)

1. Briefly explain a. precipitation and b. co-precipitations
a. Explain the theories of Acid base Titrations

Section-B

Answer all Questions (2x10 = 20 Marks)

2. a. Write the classification of Errors
(or)
b. Explain the following: (i) Accuracy (ii) Precision and (iii) Standard deviation.

3. a) Write the principles of volumetric Analysis
(or)
b) Write brief note on (i) Indicators (ii) Coprecipitation.

Answers

2. Acid base titration:-
The chemical reaction involved in acid base titration is known as neutralisation reaction. It involves the combination of H_3O^+ ions with OH^- ions to form water. In acid-base titrations, solutions of alkali are titrated against standard acid solutions. The estimation of an alkali solution using a standard acid solution is called acidimetry. Similarly, the estimation of an acid solution using a standard alkali

Solution is called alkrimetry

→ Acid base titration determines the Concentration of an acid or base by Exactly neutralising it with an acid (or) base of known Concentration.

→ Equivalence point the point at which an added amount's moles are stoichiometrically equal to the moles of acid / base in the sample.

→ Analyte's unknown solution whose Concentration is being determined in the titration.

Section - B

3. a) Classification of Errors

Errors are classified into two types - Systemic (Determinate) and Random (Indeterminate) Errors.

Systemic (Determinate) Errors:

Errors which can be avoided (or) whose magnitude can be determined is called as Systemic Errors. It can be determinable and presumably can be either avoided or corrected. Systemic Errors further classified as:

- Operational and personal Error.
- Instrumental Error.
- Errors of method
- Additive or proportional Error.

Operational and personal Error:

Errors for which the individual analyst is Responsible and are connected with the method is called as Personal Error. Ex: mistake in reading.

When Errors occur during operation is called as operational Error Eg: Transfer of Solution, overweighing, Incomplete drying.

Instrumental and Reagent Errors:-

Errors occur due to faulty instrument (or) Reagent containing impurities Eg:- uncalibrated balance, un-calibrated burette, pipette and measuring flask.

Errors of Method.

When Error occur due to method, it is difficult to correct. In gravimetric analysis, Error occurs due to Involubility of precipitates, co precipitates, post precipitates, decomposition.

Additive (or) proportional Errors:-

Additive Error does not depend on instrument present in the determination Eg: loss in weight of a crucible in which a precipitate is ignited.

Proportional Error depends on the amount of the constituent Eg: Impurities in standard compound.

Random Errors:-

It occurs accidentally (or) randomly So called as Intermediate or accidental or random Error. Analyst has no control in this Error. It follows a random distribution and a mathematical law of probability can be applied.

6. Yoga & Meditation

Regular yoga & meditation enhances mental and physical health and concentration in the work. The college organizes yoga and meditation sessions to ensure the students better performance in the academics and examinations.







Students engaged in Yoga & Meditation

7. Progress report

The various programmes organized in respect of slow learners have given positive results which can be witnessed in the improved performance of the slow learners in the end semester examinations.

“OM NAMO VENKATESAYA”

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc., Mphil., (Ph.D)
HOD OF CHEMISTRY

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Mail.ID: parameswara1964@gmail.com

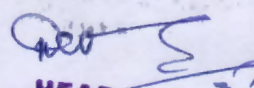
Progress Report Of Slow Learners In Chemistry After Remedial Coaching I.BSc: I & II Semesters
2016 – 2017

S.NO	Name of The Student	STATUS OF STUDENTS (SLOW)	SUBJECT	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER (After Remedial Coaching)
1	A. SAI KUMAR	SLOW LEARNERS	CHEMISTRY MPC(T.M)	38	54
2	B. RAJESH			35	22
3	C. CHANDU			31	52
4	C. PRAVEEN KUMAR			0	1
5	M. HARI KRISHNA			18	53
6	M .MOHAN			17	0
7	M. PRASANNA			34	65
8	N. SAI GANESH			37	53
9	P .RAM MAHESH BABU			31	0
10	R. MANOHAR			21	20
1	A. AKIL KUMAR	SLOW LEARNERS	CHEMISTRY MPC(E.M)	3	0
2	A. VEERANJANEYULU			33	37
3	B. PRAVEEN KUMAR			0	0
4	C. KIRAN KUMAR			33	50
5	C. SURI BABU			26	48
6	B. RAVI TEJA REDDY			36	AB
7	G. LALITH KISHORE			22	21
8	K. MEGHANA			35	59
9	M. VISHNU			22	51
10	P. THRIVEN KUMAR			30	42
11	P .SAI KRISHNA			4	42
12	P. LAKSHMAIAH			27	42
13	R. ASHOK KUMAR			31	66

1	B. POOJA	SLOW LEARNERS	CHEMISTRY CBZ(T.M)	25	57
2	B. NAVEEN KUMAR			35	57
3	C. UPPUTURIAIAH			29	54
4	C. SAI			AB	52
5	C. NAGARJUNA			32	54
6	D. KRISHNA MOORTHY			39	24
7	G. AMMULU			30	AB
8	G. HAREESH			35	54
9	J. SANJEEVA RAYUDU			34	65
10	K. MOHAN			26	63
11	K. SANGEETHA			28	88
12	K. PRIYANKA			34	91
13	M. SUNIL			30	73
14	M .BUJJAMMA			29	86
15	M .JANARDHAN GOUD			38	75
16	R. PAVAN KALYAN			15	29
17	S. BHARGAVI			27	43
18	S. HARI KRISHNA			29	58
19	S. RAMESH			26	68
20	Y .MOULI			21	41
21	Y. GURU VISHNU			24	64
22	P. CHANDRA SEKHAR REDDY			40	70
1	B. POORNIMA	SLOW LEARNERS	CHEMISTRY MZC(E.M)	32	30
2	D. SREENATH			38	60
3	G. SIVA KUMAR			27	50
4	G. GOWTHAM SAI BABU			28	35
5	H .KEERTHI KUMAR			23	50
6	K .KOUSHIK			15	20
7	K. NAVEEN			19	51
8	K. GNANA TEJA			40	58
9	M .VENKATA SURENDRA			29	AB
1	A. UMAPATHI	SLOW LEARNERS	CHEMISTRY BBC(E.M)	34	22
2	C. PRAKASH			40	34
3	G .ANIL KUMAR			36	33
4	K. NAGARAJU			26	36
5	K. NAVEEN			25	51
6	M .HEMANTH SURYA			AB	AB

LIST OF SLOW LEARNERS: II B.Sc III & IV SEMESTERS 2016-2017					
S.NO	NAME OF THE STUDENT	STATUS OF STUDENTS (SLOW)	SUBJECT	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER (After Remedial Coaching)
1	C. ROHITH KUMAR	SLOW LEARNERS	CHEMISTRY MPC(T.M)	25	52
2	G.SUNIL KUMAR			32	52
3	M.AKHIL KUMAR			32	50
1	A. SUDHAKAR	SLOW LEARNERS	CHEMISTRY MPC(E.M)	32	68
2	B. HARISH KUMAR			31	57
3	B.V.L. SIVA			25	52
4	B.JANAKIRAMUDU			38	69
5	B. THIMMAPPA			33	54
6	C. PITHAMBERAM REDDY			28	AB
7	D.SREENIVASULU			28	AB
8	K. RAJ KUMAR			39	57
9	M.MANI			40	55
10	M. SAI KUMAR			22	50
11	M. SUNIL NAIK			31	96
12	Y. RAMANJANEYULU			24	38
13	G. CHANDU			2	20
1	A. ANIL KUMAR	SLOW LEARNERS	CHEMISTRY CBZ(T.M)	29	70
2	B. THULASI			23	AB
3	B. SUPRIYA			26	54
4	B. SAI RAM			34	25
5	B. VENKATESH			26	33
6	D. PRADEEP			28	53
7	G. SARITHA			38	36
8	K. SHOBA			31	55
9	K. KOSOGAIAH			24	31
10	P. HARI KRISHNA			25	AB
11	SARASWATHI			37	AB
12	T. SAGAR			18	AB
13	M. HARI NAIK			19	24
1	D. DAMODARAM	SLOW LEARNERS	CHEMISTRY MZC(E.M)	39	20
2	M.C. VINAYAKA			29	AB
3	M. SANDEEP			39	51
4	S. PURUSHOTHAM NAIK			37	AB

5	T. NAVEEN KUMAR			24	43
6	Y. LAVANYA			37	50
7	B. VEERANJANEYULU			35	39
8	K. RAGAVENDRA			33	32
9	N. ANJI NAIK			24	50
10	T. VARUN TEJA			36	20
1	C. SWATHI	SLOW LEARNS	CHEMISTRY BBC(E.M)	40	AB
2	K. MADHU			38	38
3	M.JYOTHI REDDY			29	46
4	M.VIMAL PRABHU			29	52
5	R. MUNI PRAKASH			29	54
6	U. DENIAL			29	27
7	A. GIRI			38	26
8	K. RAJESH			40	35
9	M.CHANDRA SEKHAR NAIK			18	AB


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DEPARTMENT OF CHEMISTRY
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DEPARTMENT OF CHEMISTRY
Progress Report Of Slow Learners In Chemistry After Remedial Coaching
I .BSc : I & II SEMESTERS : 2017 – 2018

S.NO	Name Of The Student	Status Of Students (SLOW)	Programme	Marks In Odd Semester	Marks In Even Semester (After Remedial Coaching)
1	M. PRASANTH	SLOW LEARNERS	MPC	40	57
2	Y. HANUMANTHU			39	68
3	Y. HEMANTH KUMAR			39	69
4	B. HARI PRASAD			38	76
5	T.CHIRANJEEVI			38	54
6	M.RAKESH			37	51
7	V. NAGESWARA			36	52
8	P.SIVA KUMAR REDDY			32	53
9	P. SUDHEER			27	56
10	P. SRAVYA			25	AB
11	R. GUNA SEKHAR			23	52
12	K. CHANDU			22	51
13	C. GNANA SEKHAR			21	25
1	V. VENKATESH	SLOW LEARNERS	CBZ	39	55
2	B. SRAVANI			32	61
1	P. RAKESH NAIK	SLOW LEARNERS	MZC	40	41
2	G. PANDU RANGA			40	59
3	U. ANANDA			39	54
4	D. NAVEEN			38	21
5	N. PRAVEEN NAIK			34	38
6	M. SREEDHAR			29	30
7	K. VAMSI			28	56
8	T. AKASH			27	62

9	K. SANDEEP KUMAR			23	36
10	P. SRINIVASULU			22	26
11	M. RAMESH			22	33
12	G. RAJ KUMAR			21	AB
13	B. GOVARDHAN			16	18
14	A. SAMUYELL			15	AB
15	T. NAVEEN KUMAR			0	AB

1	J. VENKATESH			39	52
2	A. JAGADEESWARA REDDY			38	38
3	J. VENKATESH			35	54
4	G. KIRAN KUMAR			34	74
5	S. GIRISH KUMAR			31	43
6	E. GOPINATH GOUD			30	56
7	M. GAYATHRI			30	54
8	K. SIVA			29	28
9	K. UMA			27	35
10	S. GANGADHAR			26	43
11	K. KUSUMA PRIYA			26	40
12	P. NITHISH KUMAR			21	41
13	K. AJITH			20	AB
14	T. NAVEEN KUMAR			AB	AB

LIST OF SLOW LEARNERS: II B.Sc III&IV SEMESTERS 2017 – 2018

S.NO	Name Of The Student	Status Of Students (SLOW)	SUBJECT	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER (After Remedial Coaching)
1	C. PRAVEEN KUMAR	Slow Learners	MPC-TM	AB	AB
2	N. SAI GANESH			39	AB
3	N. CHANDRA SEKHAR			38	61
4	M. SUDHAKAR			29	27
5	A.B. LOKESH			22	AB
6	R. MANOHAR			22	20
7	B. RAJESH			19	AB
1	C. SURI BABU	Slow Learners	MPC-EM	35	39
2	A. VEERANJANEYULU			32	39
3	G. LALITH KISHORE			28	40
4	R. ASHOK KUMAR			25	AB
5	P. SAI KRISHNA			18	18
6	M. VISHNU			17	18
7	T.B. BALARAMUDU			55	74
8	P. LAKSHMAIAH			39	20
9	K. VINAY			37	57
10	A. AKIL KUMAR			6	AB
1	K. SANGEETHA	Slow Learners	CBZ-TM	40	
2	B. GAJENDRA			39	54
3	D. KRISHNA MOORTHY			37	19
4	R. INDRASENA REDDY			36	21
5	C. UPPUTURAI AH			34	26
6	Y. MOULI			30	65
7	M. SUNIL			28	49
8	C. SAI			26	17
9	B. POOJA			24	26
10	M. VAMSI			24	21
11	R. PAVAN KALYAN			22	25
1	S. VENKATA NAGATEJASREE	SLOW LEARNERS	BBC	39	38
2	B. POORNIMA			36	21

3	G. Gowtham Sai Babu			28	21
4	D. SREENATH			24	56
5	K. GNANA TEJA			23	21
6	M. RAVI NAIK			23	74
7	K.KOUSHIK			20	AB
1	T.Y. MANOJ KUMAR	SLOW LEARNERS	MZC	39	68
2	P. LAKSHMI PRASAD			33	AB
3	K. REDDI BABU			28	39
4	M. HEMANTH SURYA			28	62
5	G. ANIL KUMAR			26	63
6	T.SAI KIRAN			24	41
7	C. PRAKASH			22	30

LIST OF SLOW /ADVANCED LEARNERS: III B.Sc V & VI SEMESTERS: 2017 – 2018

S.NO	Name Of The Student	Status Of Students (SLOW)	Subject	Marks In Odd Semester	Marks In Even Semester (After Remedial Coaching)
1	C. ROHITH KUMAR	SLOW LEARNERS	MPC-TM	37	51
2	R. VINOD KUMAR			22	23
3	O. ASHOK KUMAR			20	23
1	B. HARISH KUMAR	SLOW LEARNERS	MPC-EM	33	34
2	K. SUJITH KUMAR			32	41
3	K.ASHOK KUMAR REDDY			29	51
4	M. SAI KUMAR			22	18
5	K. YELLAPA			21	33
6	Y. RAMANJANEYULU			18	7
7	D. SREENIVASULU			0	0
1	B. VENKATESH	SLOW LEARNERS	CBZ-TM	37	54
2	K. KOSOGAIAH			27	53
3	M. HARI NAIK			22	23
4	HARI KRISHNA			20	23
1	M.C. VINAYAKA	SLOW LEARNERS	MZC	40	42
2	T. AMARNADH			38	AB
3	K. RAGAVENDRA			38	40
4	R. JAGADESH BABU			37	61
5	M. SANDEEP			29	AB
6	VARUN TEJA			23	31
7	N. ANJI NAIK			20	41
1	Y. VIKAS REDDY	SLOW LEARNERS	BBC	39	35
2	K.RAJESH			38	AB
3	M. JYOTHI REDDY			37	63
4	U. DENIAL			37	54
5	A. GIRI			31	AB
6	S.S. ANOOF			23	53

DEPARTMENT OF CHEMISTRY

Progress Report Of Slow Learners In Chemistry After Remedial Coaching

I .BSc: I & II Semesters: 2018 – 2019

S.NO	NAME OF THE STUDENT	Status Of Students (SLOW)	Subject	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER (After Remedial Coaching)
1	O. ARAVIND	SLOW LERNERS	CHEMISTRY (MPC)	3.7	5.4
2	A. HEMANTH KUMAR			3.3	2.2
3	B.SIVARAMUDU			3.3	5.2
4	B. DATHATHREYUDU			3.1	5.4
5	K.NITHISH			3.1	
6	P. MUNI SEKhar			2.8	5.6
7	S. NANDA KUMAR			2.1	6.6
1	B. GANGADHAR	SLOW LERNERS	CHEMISTRY (CBZ)	3.7	5.6
2	A. MANENDRA			3.7	5.4
3	K. HIMA BINDU			3.3	
4	K. VENKATESH			3.2	
5	B. MOHAN NAIK			3.1	6.1
6	D. SAI KUMAR			2.8	
7	M. HARI PRAKASH NAIK			2.7	3
8	M. NARESH BABU			2.7	
9	K. KALPANA			2.5	AB
10	C. CHANDU			2.4	AB
11	B. CHANDRA SEKhar NAIK			2.2	5.7
12	K.CHENCHIAIAH			2.2	AB
13	G. NAGARAJU			2	5.5
14	G. MAHENDRA			1.9	
1	K. SANTHOSH KUMAR	SLOW LERNERS	CHEMISTRY (MZC)	AB	AB
2	S. SOMA SEKhar			4	5.3
3	E. NAVEEN			3.9	3.3
4	G. PRANAY KUMAR REDDY			3.5	5.7
5	E. REDDY SEKhar			3.5	5.2
6	C. RAJA SEKhar			3.2	
7	B. RANGADU			3.2	2.5
8	N. VINOD KUMAR			3.2	3.9

9	P. SAI THARUN			2.8	
10	A. BALAJI			2.7	6.3
11	T.GOWTHAM			2.7	5.5
12	G. SWARNA LATHA			2.7	2
13	B. GANESH			2.6	3.2
14	B. VIJAY KUMAR			2.6	2.5
15	G. BHEEMUDU			2.5	5.3
16	T. MOUNIKA			2.4	5.3
17	Y. DEESNESH BABU			2.3	5.8
1	M. NARASIMHULU	SLOW LERNERS	CHEMISTRY (BBC)	3.7	6.4
2	K. BHARGAV			4	5.6
3	B. CHINNA THIMMAPPA			4	5.7
4	S. SAI SREE			4	4.2
5	P. VANDHANA			4	6.2
6	M. SATISH KUMAR			3.6	6.4
7	K. RAMANJINEYULU			3.5	5.3
8	P. NARASHIMHA REDDY			3.2	3.3
9	U. SOMESWARA RAO			3.1	5.9
10	R. TEJASREE			2.8	3.2
11	K.DEENA			2.6	3.2
12	M.NAVEEN			2.5	6.1
13	T.VIJAY			2.4	3.2
14	B. HAREESH			2.2	3.4
15	P. JAYCHANDRA			1.6	5.4
16	A. PEDDA OBULESU			AB	AB
17	T.SUJAN			AB	AB

LIST OF SLOW LEARNERS: II B.Sc : III & IV SEMESTERS : 2018 - 2019

S.No	Name of The Student	Status Of Students (SLOW)	Subject	Marks in odd semester	Marks In Even Semester (After Remedial Coaching)
1	P. SIVA KUMAR REDDY	SLOW LERNERS	CHEMISTRY (MPC)	21	5.1
2	P. MAHESH			31	2.4
3	P. SUDHEER			33	6.3
4	N. SUMAN			37	5
5	M. RAKESH			39	AB
1	B. SRAVANI	SLOW LERNERS	CHEMISTRY (CBZ)	44	6.5
2	S. BALAJI NAIK			43	3.4
3	T. REVATHI			47	4.4
1	T. AKASH	SLOW LERNERS	CHEMISTRY (MZC)	AB	6.5
2	P. SRINIVASULU			17	AB
3	SK. MAZEED BABU			19	AB
4	M. SREEDHAR			20	6.5
5	P. RAKASH NAIK			25	6.8
6	B. YUGANDHAR			25	5.2
7	N. PRAVEEN NAIK			27	6.2
8	V. PARSARAM			30	5.5
9	A. BALAJI			35	3.7
1	P. NITHISH KUMAR	SLOW LERNERS	CHEMISTRY (BBC)	19	2.2
2	G. KIRAN KUMAR			29	7.4
3	S. GANGADHAR			33	5.4
4	C. LOKESH			35	6.6
5	M. GAYATHRI			36	6.2
6	E. GOPINATH GOUD			38	8
7	J. VENKATESH			39	2.7

LIST OF SLOW LEARNERS: III B.Sc V & VI SEMESTER 2018 - 2019

S.NO	NAME OF THE STUDENT	Status Of Students (SLOW)	Subject	Marks in odd semester	Marks In Even Semester (After Remedial Coaching)
1	M. HARI KRISHNA	SLOW LERNERS	CHEMISTRY (MPC)	17	68
2	M. SUDHAKAR			19	68
3	M. VISHNU			21	AB
4	P. LAKSHMAIAH			21	22
5	K. SACHIN CHOWDARY			30	71
6	G. MIDHILESH			39	65
		SLOW LERNERS	CHEMISTRY (CBZ)	NIL	
	NIL				
1	G. GOWTHAM SAI BABA	SLOW LERNERS	CHEMISTRY (MZC)	17	22
2	K. NAVEEN			22	37
3	S. VENKATANAGA TEJSREE			34	53
4	M. BHANU PRAKASH			40	59
		SLOW LERNERS	CHEMISTRY (BBC)		
1	C. PRAKASH			20	27
2	B. SUNIL KUMAR NAIK			21	50
3	P. LAKSHMI PRASAD			21	53
4	A. HARI KRISHNA			28	33
5	N. MAHESH KUMAR			30	52
6	G. ANIL KUMAR			33	AB
7	T. SAI KIRAN			33	55
8	M. HEMANTH SURYA			36	35

Progress report of slow learners after remedial coaching

LIST OF SLOW LEARNERS: I B.Sc I & II SEMESTERS - 2019 – 2020

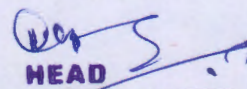
S.No	Name of the Student	Status Of Students (Slow)	SUBJECT	MARKS IN ODD SEM	MARKS IN EVEN SEM
1	A. RAGHU KUMAR	SLOW	MPC.T.M	3.8	2.5
2	A. VENKATAARJUN	LEARNERS	&E.M		3.2
3	A. CHANDU PRAKASH				3.6
1	A M. SAMEERA	SLOW	CBZ.T.M	1.9	2.7
2	B.DINESH NAIK	LEARNERS		2	3.3
3	C. LAKSHMI NARAYANA			2.2	3.5
4	C. HARINAD			2.2	5.3
5	C. HEMANTH KUMAR			2.4	5.3
6	C. CHANDRA SEKHAR			2.5	5.4
7	D. SURYA			2.7	5.4
8	D. CHANDRA			2.7	5.4
9	D.MUNI LATHA			2.8	5.4
10	E. SREE KANTH			3.1	5.5
11	VEERANJINEYULU			3.2	5.5
12	G. RAJASEKHAR			3.3	5.5
13	K. SASI KUMAR			3.7	5.9
14	K. USHA			3.7	5.9
1	A. GIRIRAMESH	SLOW	MZC E.M	2.2	5.1
2	A. GIRISHMA	LEARNERS		3.1	5.4
1	A. CHANDINI	SLOW	BBC E.M	2.4	2.5
2	A. POOJITHA	LEARNERS		2.8	5.2
3	D. GOUTHAM KRISHNA			2.8	5.4
4	VENKATASAI KUMAR			3.2	5.4
5	V. MANJUNADH NAIK			3.5	5.5
6	D. GOVARDHAN			3.7	5.6
7	V. JANSI LAKSHMI			3.8	5.6
8	C. BABU			3.8	5.8
9	C. KARUNAKAR			3.8	6
10	C. RAMAKANTH			3.9	6

LIST OF SLOW /ADVANCED LEARNERS: II B.Sc : III & IV SEMESTERS - 2019 - 2020

S.No	Name of the Student	STATUS OF STUDENTS (Slow)	SUBJECT	MARKS IN ODD SEM	MARKS IN EVEN SEM
1	NIL	NIL	MPC E.M &T.M	NIL	
1	A. BALAJI	SLOW LEARNERS	MZC(EM)	NIL	2.6
2	G.BEEMUDU				3
3	G.CHANDRA SEKAR				3.6
4	Y.DINESH BABU				3.8
1	T.ASHA	SLOW LEARNERS	CBZ(EM)	2.9	5.4
2	B.CHANDRASEKAHR			3.4	6
3	C.CHANDU			3.7	
1	T.BARATH KUMAR	SLOW LEARNERS	BBC.E.M	2.1	3.3
2	K. BARGAV			2.1	4.1
3	P. CHANDRA SHEKAR			2.1	5
4	B. CHINNA			2.2	5
5	S. CHOUDAIAH			2.2	5.2
6	K.DEENA			2.4	5.3
7	B.DILIP			2.5	5.3
8	B.HARI BABU			2.7	5.5
9	K.JAGADESWARA			2.7	5.6
10	M. CHANDRA			2.7	5.6
11	T.JAYARAM			2.8	6
12	B. LAKSHMI KANTH			2.9	6.2
13	K. LAKSHMI NARASIMHA			3	7
14	G. MALLIKARJUNA			3.1	7
15	S. MANJUNDH			3.1	7.1
16	M. MAREPPA			3.3	7.2

LIST OF SLOW LEARNERS: III B.Sc V&VI SEMESTERS - 2019 - 2020

S.No	Name of the Student	STATUS OF STUDENTS (Slow)	SUBJECT	MARKS IN ODD SEM	MARKS IN EVEN SEM
1	T.AVINASH	SLOW LEARNERS	MPC E.M	2.9	0
2	K.BAGYA LAKSHMI		&T.M	3.2	2.3
3	S.GANGADHAR			3.5	2.6
1	NIL	SLOW LEARNERS	CBZ(T.M)	NIL	
1	A.BALAJI	SLOW LEARNERS	MZC(E.M)	3.7	4.1
1	NIL	SLOW LEARNERS	BBC(E.M)	NIL	


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"OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY

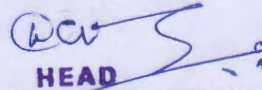
Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

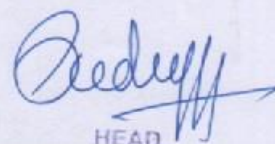
SEMESTER WISE NUMBER OF SLOW LEARNERS				
Batch „ 2016-19				
PROGRAMME NAME	PROGRAMME CODE	SEMESTER	TOTAL NUMBER OF STUDENTS	TOTAL NUMBER OF SLOW LEARNERS
B.Sc	MPC	II SEMESTER	44	23
		III SEMESTER	44	15
		IV SEMESTER	39	21
		V SEMESTER	35	14
		VI SEMESTER	31	8

PROGRAMME WISE SLOW LEARNERS PERFORMANCE				
Batch „ 2016-19				
PROGRAMME NAME	PROGRAMME CODE	SEMESTER	TOTAL NUMBER OF STUDENTS	TOTAL NUMBER OF SLOW LEARNERS
B.Sc	CBZ	II SEMESTER	24	22
		III SEMESTER	27	4
		IV SEMESTER	25	11
		V SEMESTER	26	12
		VI SEMESTER	24	2

The above data is a direct eye witness for the positive outcome of remedial coaching.


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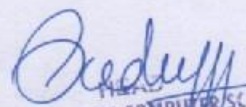
DEPARTMENT OF COMPUTER SCIENCE					
PROGRESS REPORT OF SLOW LEARNERS IN COMPUTER SCIENCE AFTER REMEDIAL COACHING					
LIST OF SLOW LEARNERS					
S.NO	NAME OF THE STUDENT	STATUS OF STUDENTS (SLOW)	Group & Medium	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER
1	SANAMBATLA HIMAVANTH	SLOW LEARNERS	II YEAR MSCS-EM	42	44



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DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE					
PROGRESS REPORT OF SLOW LEARNERS IN COMPUTER SCIENCE AFTER REMEDIAL COACHING					
LIST OF SLOW LEARNERS					
S.NO	NAME OF THE STUDENT	STATUS OF STUDENTS (SLOW)	Group & Medium	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER
1	G VAMSI KRISHNA	SLOW LEARNERS	II YEAR IV SEM MSCS	20	66
2	KESANI SAI KUMAR			18	41
3	SANAMBATLA HIMAVANTH			22	AB


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 DEPARTMENT OF COMPUTER SCIENCE
 S.G.S. ARTS COLLEGE
 TIRUPATI

1. Quiz competition

In order to boost self esteem and confidence in the students the college organized quiz competitions. They encourage students self awareness of progress and self assessment. This can help the students to identify the areas they need to develop. A sample copy is presented here.

Tirupati
Date: 26-08-2019

From
The HOD of CHEMISTRY
S.G.S. Arts College
Tirupati

To
The Principal
S.G.S. Arts College
Tirupati

Respected Sir

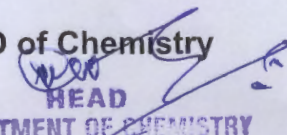
Sub:- Permission for conducting Quiz programme – Request –
Regarding.

-:oo0oo:-

We are conducting Quiz programme for I, II and III year students of
B.Sc. (MPC, MZC, BBC & CBZ) in Room No.225. So, kindly grant us permission
on 29-08-2019.

Thanking you Sir

Yours faithfully

HOD of Chemistry

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI



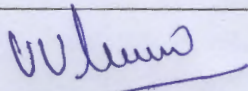
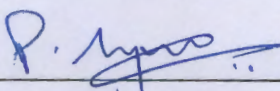
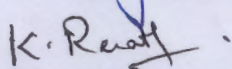
"OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOOR
ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY



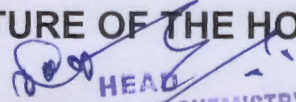
Sri D.PARAMESWARA
M.Sc., Mphil., (Ph.D)
HOD OF CHEMISTRY

Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

COMMITTEE FOR QUIZ PROGRAMME
ACADEMIC YEAR – 2019-2020

S.NO.	NAME OF THE COMMITTEE MEMBERS	SIGNATURE
1	Dr. V. Venkata Lakshmi, Coordinator	
2	Dr. P. Suguna, Member	
3	Smt. K. Revathi, Member	

SIGNATURE OF THE HOD


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

chem

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOUR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

TOPIC: PHARMACEUTICAL USES OF MEDICINAL CHEMISTRY

DATE: 29/08/2019

S.No	Programme name	Programme Code	Name Of The Student	Signature Of The Student	BATCH	Marks Awarded (30)	Remarks
1	II B.Sc	MPC	M.Tharunkumar	m.Tharunkumar	A	28	I PLACE
2			B.Sharmila	B.Sharmila			
3			B.Varaprasad	B.Varaprasad			
4			M.SandeepReddy	M.SandeepReddy			
5			B.Nagaraju naik	B.Nagaraju naik			
6			B.sri vijay	B.sri vijay			
7		MPC	C.Varaprasad	C.Varaprasad	B	25	II PLACE
8			C.Chandra uday kumar	C.Chandra uday kumar			
9			T.Vishnuvardhan	T.Vishnuvardhan			
10			M.Raja	M.Raja			
11			N.KanthaReddy	N.KanthaReddy			
12			D.Karthik	D.Karthik			
13		MZC	C.Chandu	C.Chandu	C	24	
14			P.Lokesh	P.Lokesh			
15			P.Vaishnavi	P.Vaishnavi			
16			R.Pavithra	R.Pavithra			
17			M.Kumar	M.Kumar			
18			N.Usman	N.Usman			
19		BBC	K.Venkatasuresh	K.Venkatasuresh	D	27	II PLACE
20			P. Praddeep Kumar	P. Praddeep Kumar			
21			K.Sunetha	K.Sunetha			
22			V. Manjunadh Naik	V. Manjunadh Naik			
23			K.RAJA	K.RAJA			
24			A.POOJITHA	A.POOJITHA			
25		BBC	M.Devendra	M.Devendra	E	23	
26			D. Goutham Krishna	D. Goutham Krishna			
27			B.Venkatasai Kumar Naik	B.Venkatasai Kumar Naik			
28			P. Govardhan	P. Govardhan			
29			S.Pravallika	S.Pravallika			
30			T. Sreelatha	T. Sreelatha			

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATI



"OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOOR
ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY



Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No:9490728655
Mail.ID: parameswara1964@gmail.com

QUIZ - DEPARTMENT OF CHEMISTRY

ACADEMIC YEAR ::2019-20						
S.NO	PROGRAMME NAME	PROGRAMME CODE	TITLE OF THE QUIZ PROGRAMME	DATE OF CONDUCTION	TOTAL NO. OF BATCHES	WINNING BATCH
1	II B.Sc.	MPC	PHARMACEUTICAL USES OF MEDICINAL CHEMISTRY	29/8/2019	5 (30 Nos.)	1(6Nos.) (MPC)
		CBZ				
		MZC				
		BBC				

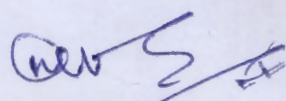
HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

S.G.S. ARTS COLLEGE, TTD, TIRUPATI

Department of CHEMISTRY

EVENT ORGANISED REPORT: 2019 – 20

1	Name of the Department	Chemistry
2	Name of Event Organized	QUIZ COMPETITION
3	Title of Event Organized	Pharmalogical uses of medicinal plants
4	Name of Co-ordinator of Event	Sri.D.Parameswara
5	Class of the Participants	B.Sc,MPC,BBC,MZC&CBZ
6	No. of Participants	30
7	Objective of the Event	➤ To encourage students to look beyond their textual knowledge and establish a relationship between theory and application of the learnt concepts
8	Outcome of the Event	<ul style="list-style-type: none">• Interactive platform• Changes scope of learning• Encourages team work• Bridges academia-industry gap


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI





**Presentation of certificate to the quiz winner A.Raghu kumar &M. Sandeep Reddy of
IIB.Sc : MPC -EM by HOD Sri.D.PARAMESWARA**

QUIZ FORMAT

TIRUPATI,
DATE : 24-08-2019

FROM,
The HOD of MICROBIOLOGY
S.G.S. ARTS COLLEGE,
TIRUPATI.

TO ,
THE PRINCIPAL
S.G.S. ARTS COLLEGE,
TIRUPATI.

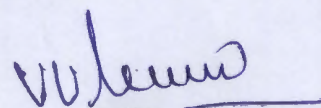
Respected Sir/ Madam,

Sub : Request for permission to conduct Quiz programme
Regarding

We are conducting Quiz programme for III year students of
MICROBIOLOGY in Room No 316 So, kindly grant us permission to
conduct Quiz programme on **26-08-2019**.

Thanking you Sir / Madam

Yours faithfully


HOD

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

DEPARTMENT OF MICROBIOLOGY
COMMITTEE FOR QUIZ PROGRAMME
ACADEMIC YEAR – 2019 – 2020

S.NO.	NAME OF THE COMMITTEE MEMBERS	SIGNATURE
1	Dr. K. Sridevi	<i>K. Sridevi</i>
2		

SIGNATURE OF THE HOD

[Signature]
HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

Phone: (0877) 2264599

E-mail: sgsartscollegettdd@gmail.com

Web: sgsac.edu.in



TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

QUIZ - DEPARTMENT OF MICROBIOLOGY

ACADEMIC YEAR 2019 -2020

S.NO	PROGRAMME NAME	PROGRAMME CODE	TITLE OF THE QUIZ PROGRAMME	DATE OF CONDUCTION	TOTAL NO. OF BATCHES	WINNING BATCH
1	QUIZ	III MZC	Food Microbiology	28-08- 2018	4	B

V. Srinivas
HOD HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

**S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD,
TIRUPATI**

Student supporting programs / Events conducted

Name of the department: MICROBIOLOGY

Academic year: 2019 - 2020

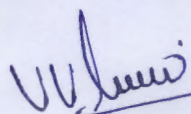
S.No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	Quiz Competition	26-08-2019	III MZC	12	

V. V. V.
HOD
HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

V. V. V.
HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

S.G.S. ARTS COLLEGE, TTD, TIRUPATI
Department of MICROBIOLOGY
EVENT ORGANISED REPORT: 2019 - 2020

1	Name of the Department	MICROBIOLOGY
2	Name of Event Organized	QUIZ
3	Title of Event Organized	<u>Food microbiology</u>
4	Name of Coordinator of Event	Dr. K. Sridevi
5	Class of the Participants	MZC
6	No. of Participants	12
7	Objective of the Event	To know the pathogenic and non pathogenic micro organisms that found in food and their importance in Fermented food , food borne disease etc.,
8	Outcome of the Event	It is useful for methods of isolation , detection, identification of food borne and spoiled food in food Microbiology

 **HOD**

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

S.G.S. ARTS COLLEGE, TTD, TIRUPATI
Department of MICROBIOLOGY
EVENT ORGANISED REPORT: (Year)

A quiz competition is conducted by the Department of
MICROBIOLOGY , in the view of Competitive awareness on
26-08-2019 .

Batch. No	Name of the Student	Class	Roll No	Student Signature	Marks Awarded	Total	Remarks
1	Chakrapani Tharun	MZC	18704	Chakrapani Tharun	20	8	
	D.Gawtham gowda	MZC	18711	D. Gawtham gowda		7	
	K.E.pavan kumar	MZC	18718	K.E. Pavan kumar		5	
2	E.Reddy sekhar	MZC	18724	E.Reddy sekhar	23	8	
	K.Sai krishna	MZC	18727	K.Sai krishna		8	
	P.Subramanyam	MZC	18730	P.Subramanyam		7	
3	R.Gayathri	MZC	18708	R. Gayathri	25	8	WIN
	M.Gurramma	MZC	18712	M. Gurramma		9	
	K.Kalyani	MZC	18713	K. Kalyani		8	
4	G.Nagarjuna	MZC	18716	G. Nagarjuna	19	7	
	E.Naveen	MZC	18717	E. Naveen		7	
	B.N.Rangadu	MZC	18722	B.N. Rangadu		5	

Organizing Committee:

1. Dr. K. Sri devi

K. Sri devi

*** Photo Evidence**



Tirupati

Date. 02-09-2019

From

The HOD of Computer Science
S.G.S. Arts College,
Tirupati.

To

The Principal,
S.G.S. Arts College,
Tirupati.

Respected Sir/Madam,

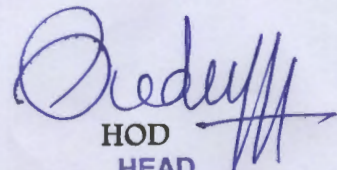
Sub. - Permission for conducting Quiz programme - Request - Regarding.

We are conducting Quiz programme for II & III year students of B.Sc(Mpcs & Mscs)

Em in Room No 214. So, kindly grant us permission on 04-09-2019.

Thanking you Sir/Madam,

Yours faithfully

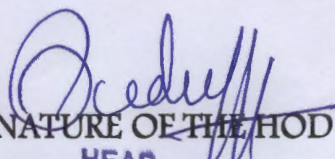


HOD
HEAD

DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE
COMMITTEE FOR QUIZ PROGRAMME
ACADEMIC YEAR (2019-2020)

S.NO.	NAME OF THE COMMITTEE MEMBERS	SIGNATURE
1	Smt. B.Triveni, Coordinator	B. Triveni
2	Smt N.Jayalakshmi	N. Jayalakshmi
3	Sri D.Chaithanya Kumar	D. Chaithanya Kumar


SIGNATURE OF THE HOD
HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

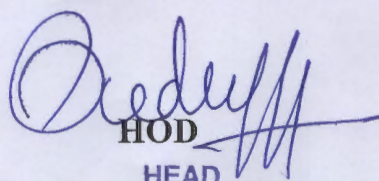
S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

Student supporting programs / Events conducted

Name of the department: Computer Science

Academic year: (2019-2020)

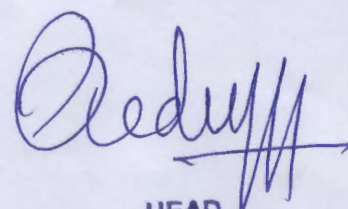
S.No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	QUIZ COMPETITION	04-09-2019	III B.Sc (Mpcs & Mscs)	20 (4 Batches)	


HOD
HEAD

DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S. ARTS COLLEGE, TTD, TIRUPATI
Department of Computer Science
EVENT ORGANISED REPORT: (2019-2020)

1	Name of the Department	Computer Science
2	Name of Event Organized	QUIZ COMPETITION
3	Title of Event Organized	Object Oriented Programming using Java
4	Name of Coordinator of Event	Smt B.Triveni
5	Class of the Participants	II & III B.Sc MPCS, MSCS [EM]
6	No. of Participants	20
7	Objective of the Event	<ul style="list-style-type: none"> To evaluate the knowledge of the participants within academics as well as beyond academics and to make them familiar with the prospects of quizzes and the objectivity of the questions.
8	Outcome of the Event	<ol style="list-style-type: none"> There is clear evidence from psychological experiments that practicing retrieval of something after learning it, for instance by taking a quiz or test, makes us more likely to retain it for the long term. Testing identifies gaps in knowledge. Essentially it reduces forgetting which makes the next related study area more productive.



HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Phone: (0877) 2264599

E-mail: sgsartscollegettd@gmail.com

Web: sgsac.edu.in



TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

QUIZ - DEPARTMENT OF COMPUTER SCIENCE
ACADEMIC YEAR. (2019-2020)

S.NO.	PROGRAMME NAME	PROGRAMME CODE	TITLE OF THE QUIZ PROGRAMME	DATE OF CONDUCTION
	II & III B.Sc	MPCS	Object Oriented Programming using Java	04-09-2019
		MSCS		

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI



TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

QUIZ - DEPARTMENT OF COMPUTER SCIENCE ACADEMIC YEAR (2019-2020)

S.NO.	PROGRAMME NAME	PROGRAMME CODE	TITLE OF THE QUIZ PROGRAMME		TOTAL NO. OF BATCHES	WINNING BATCH
1	III B.Sc	MPCS	17495	Bhargav Embeti	A	1 ST PRIZE
2			17496	Brahmaiah Sandu		
3			17505	Mahesh Avisineni		
4			17513	Prabhakar Badikela		
5			17526	Surekha Pulluru		
6		MSCS	17541	Allipeera Dudekula	B	
7			17545	Bhuvanesh Ragam		
8			17550	Ganesh Kumar Gujjala		
9			17513	Prabhakar Badikela		
10			17518	Ravali Baduru		
11	II B.Sc	MPCS	18522	Hemasri Kuppa	C	
12			18524	Jayalakshmi Kailasam		
13			18525	Mabbu Sharif Gajuguduru		
14			18543	Sandhya Putturu		
15			18558	Varshini Annem		
16		MSCS	18577	Deepa Kadirimangalam	D	2 ND PRIZE
17			18586	Indhumathi Lalapeta		
18			18589	Lakshmi Manne		
19			18609	Venkata Chanukya Kethineni		
20			18610	Venkat Nisanth Kanupuru		

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Photo Gallery



2. STUDENT SEMINAR

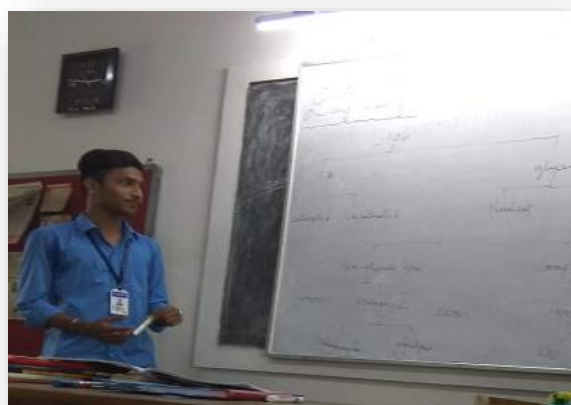
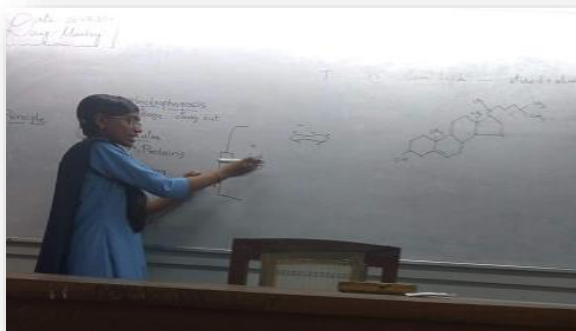
The college promotes the learning capabilities of the students by making them engaged in seminars on the topic of their choice. This ensures the students to learn about the latest information and new skills related to the concerned subjects.



Student seminars conducted in the department of sankrit for the academic year 2017-18 on the topic “UNNATHI” .



Student seminars conducted in the department of chemistry for the academic year 2019-20 on the topic “Factors affecting distributionlaw” .



Student seminars conducted in the department of Microbiology for the academic year 2019-20 on the topic "General properties of liquids" .



OM NAMO VENKATESAYA”

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

Sri Govindaraja Swamy Arts College

Department of Computer Science

STUDENT SEMINAR

Date : 25-09-2018

OBJECTIVES:

1. To progress the communication skills by motivating them to give seminars in different topics.
2. To make students Expertise in Verbal Communication.
3. To remove stage fear in students.
4. To explore expression of the students and also to gauge their awareness quotient.

CONVENER

- Sri P.Udaya, HOD, Department of Computer Science

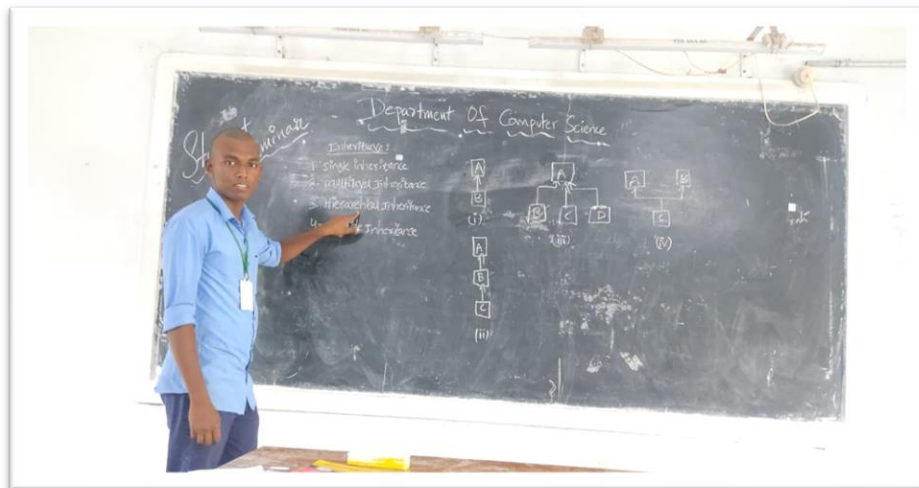
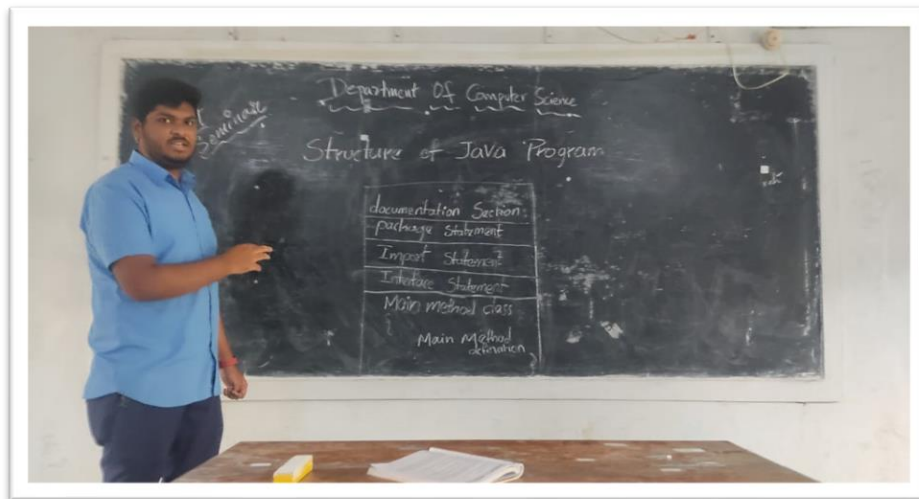
ORGANIZING MEMBERS

- SmtB.Triveni
- SmtN.Jayalakshmi
- Sri D.Chaitanya Kumar

TOPIC: Fundamentals of Java

- Fundamentals of OOP
- Structure of Java Program
- Control Structures in Java
- Types of Inheritance
- Interface
- Thread Life Cycle
- Applet Life Cycle

PHOTO GALLERY:



Attendance Sheet for IIB.Sc. (MPCS) EM

Sno	Roll No	Name of the Student	Signature
1	17491	Anjan KumarPottiramigalla	P. Anjan Kumar
2	17492	ArjunKarli	K. Arjun
3	17493	AshokKanimeni	K. Ashok
4	17500	JosephMerugumolu	M. Joseph
5	17501	JyothsnaKondru	K. Jyothsna
6	17503	LokeshKommala	K. Lokesh
7	17504	MadhuUndela	V. Madhu
8	17505	MaheshAvisineni	A. Mahesh
9	17506	ManoharBandari	B. Manohar
10	17509	MohiniIrri	T. Mohini
11	17510	Om PrakashAaramadaka	A. Om Prakash
12	17511	Pavan KumarAnderi	A. Pavan Kumar
13	17513	PrabhakarBadikela	B. Prabhakar
14	17514	Praneeth KumarUppuluru	V. Praneeth Kumar
15	17515	PrasanthMuppuri	M. Prasanth
16	17517	Ram PrasadKorlakunta	K. Ram Prasad
17	17518	RavaliBaduru	B. Ravali
18	17519	Ravi Durga Prasad Donti	D. Ravi Durga Prasad
19	17520	Ravi TejaMekala	M. Ravi Teja
20	17521	Sai PrakashPenchili	P. Sai Prakash
21	17522	SatishVeligiram	V. Satish
22	17523	Shaik Adam Basha	Shaik Adam Basha
23	17524	SravaniPidugu	P. Sravani
24	17525	SreekanthChejarla	C. Sreekanth
25	17526	SurekhaPulluru	P. Surekha
26	17527	Suri BabuMotikala	M. Suri Babu
27	17528	TejaswiniManyam	M. Tejaswini
28	17531	UrukunduMaliga	M. Urukundu
29	17532	VedhavathiPamula	P. Vedhavathi
30	17533	VeerannaBoya	B. Veeranna
31	17534	VeereshBoya	B. Veeresh
32	17535	VeereshNallagunti	N. Veeresh
33	17537	Vignesh KumarRaja	R. Vignesh Kumar
34	17538	YaminiThalithoti	T. Yamini
35	17540	YeswanthPangaluri	P. Yeswanth



 HEAD
 DEPARTMENT OF COMPUTER SCIENCE
 S.G.S. ARTS COLLEGE
 TIRUPATI

Attendance Sheet for II B.Sc. (MScS) EM

Sno	Roll No	Name of the Student	Signature
1	17541	AllipeeraDudekula	D. Allipeera
2	17542	AnjaliBodannagari	B. Anjali
3	17544	AnushaMinuku	M. Anusha
4	17545	BhuvaneshRagam	B. Bhuvanesh
5	17546	ChalapathiArava	A. Chalapathi
6	17547	Chandra SekharNama	N. Chandra Sekhar
7	17548	Deva.G	G. Deva
8	17550	Ganesh KumarGujjala	G. Ganesh Kumar
9	17551	HemanthMekala	M. Hemanth
10	17552	HimavanthSanambatla	S. Himavanth
11	17553	JagadeswariSakam	S. Jagadeswari
12	17554	Kiran KumarChakali	C. Kiran Kumar
13	17555	LokeshChitteni	C. Lokesh
14	17557	LoknathUpendram	V. Loknath
15	17558	MadhuPeta	B. Madhu
16	17559	ManikantaBalagani	B. Manikanta
17	17560	MeghanaMaddela	M. Meghana
18	17561	Muni Sai KumarPonna	P. Muni Sai
19	17562	NarsappaPandikona	P. Narsappa
20	17563	Raja Dhannasi	D. Raja
21	17564	RajuSingadi	S. Raju
22	17565	Rethish Sai.JM	M. Rethish Sai
23	17566	RupeshKanakappagari	K. Rupesh
24	17567	Sai KumarAnthati	A. Sai Kumar
25	17571	Santhosh KumarKantineni	K. Santhosh
26	17573	Sathish KumarGurrappagari	G. Sathish Kumar
27	17574	SrinivasuluBapana	B. Srinivasulu

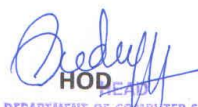

 DEPARTMENT OF COMPUTER SCIENCE
 S.G.S. ARTS COLLEGE
 TIRUPATI

Sno	Roll No	Name of the Student	Signature
28	17575	Sumalatha.B	B. Sumalatha
29	17576	SunilPasupula	P. Sunil
30	17577	SupriyaDidi	D. Supriya
31	17578	SureshGonimadathala	G. Suresh
32	17579	SuryaGuggilla	G. Surya
33	17584	VenugopalPutakala	P. Venug
34	17585	VinayJadapalli	J. Vinay
35	17586	Vishnu VardhanMamuduru	M. Vishnu Vardhan
36	17587	YeswanthGownipalli	G. Yeswanth


 HEAD
 DEPARTMENT OF COMPUTER SCIENCE
 S.G.S. ARTS COLLEGE
 TIRUPATI

REPORT:

The department of Computer Science has magnificently organised a seminar for II BSc MSCs-EM and MPCs-EM students on the above listed topics. Students of II BSc MSCs-EM and MPCs-EM have actively participated in this event and given away seminars on their respective topics. This Seminar was able to Encourage, motivate and Improve Communication skills for the better performance of students. Also this seminar enriched the knowledge in Fundamentals of Java Programming related topics. The Students and Lecturers of the Department of computer science participated the event and made this programme grand success.


HOD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

Student supporting programs / Events conducted

Name of the department: Statistics

Academic year:2019-20


S. No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	SEMINAR	25/02/2020	B.Sc: (MSCs MPS) BA(MES)	50	-

S.G.S. ARTS COLLEGE, TTD, TIRUPATI

Department of STATISTICS

EVENT ORGANISED REPORT: 2019-20

1	Name of the Department	Statistics
2	Name of Event Organized	Seminar
3	Title of Event Organized	Role of Statisticians – Census survey and Analysis
4	Name of Co-ordinator of Event	Dr.D.Chandra Kesavulu Naidu
5	Resource Person	Prof.B.Sarojamma,Board of Studies Chair person, Dept of Statistics, S.V.University.
6	Class of the Participants	B.Sc:MSCs MPS BA: MES
7	No. of Participants	50
8	Objective of the Event	➤ To create awareness in students about Role of Statisticians – Census survey and Analysis
9	Outcome of the Event	➤ Became aware of process of Census survey in India ➤ Understood the role of Statisticians in the process


HOD
D. CHANDRA KESAVULU NAIDU,
M.Sc., M.Phil., Ph.D
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
TIRUPATI

A seminar was organised and conducted by the Head of Dept, Statistics Dr. D. Chandra Kesavulu Naidu, for the students of Statistics (BA, BSc).

Topic: "Role of Statisticians - Census Survey and Analysis".

Resource Person: Prof. B. Sarojamma
Board of Studies - Chairperson
Dept of Statistics
S.V. University, Tirupati.

All the staff members and students of the dept attended the seminar. Students were made aware of the process of Census Survey in India, its importance and role of Statisticians.

Staff: 1) Dr. K. Visayakumar

2) Dr. K. Venkatesh

3) Dr. M. Vasudhara Reddy

D. CHANDRA KESAVULU NAIDU,

M.Sc., M.Phil., Ph.D.

LECTURER & HEAD

DEPT. OF STATISTICS

S.S.S. ARTS COLLEGE

TIRUPATI

① B. Mahesh

② D. Mounika

③ B. Lavanya

④ K. Akhil

⑤ M. Ashok

⑥ P. Bhadri

⑦ M. Varma Teja

⑧ U. Naveen

⑨ V. Venu Gopal

⑩ P. Sambayya

⑪ K. Anil Kumar

⑫ D. Lakshmini

⑬ M. Vignesh

B. Mahesh

D. Mounika

B. Lavanya

K. Akhil

M. Ashok

P. Bhadri

M. Varma Teja

U. Naveen

V. Venu Gopal

P. Sambayya

K. Anil Kumar

D. Lakshmini

M. Vignesh

- | | | | |
|-------------------------|--------------------|-----------------------|------------------|
| (14) T. Uma Maheswari | T. Uma Maheswari | (48) M. Dinesh | M. Dinesh |
| (15) B. Siva Jyothi | B. Sai Jyothi | (49) G. Mohan Krishna | G. Mohan Krishna |
| (16) N. Keerthi | N. Keerthi | (50) G.D. Gopal | G.D. Gopal |
| (17) N. Sireesha | N. Sireesha | | |
| (18) T. Raju | T. Raju | | |
| (19) P. Vamsi | P. Vamsi | | |
| (20) S. Manoj Kumar | S. Manoj Kumar | | |
| (21) B. Anand | B. Anand | | |
| (22) R. Chandu | R. Chandu | | |
| (23) P. Devi | P. Devi | | |
| (24) C. Iatha | C. Iatha | | |
| (25) G. Bhumiika | G. Bhumiika | | |
| (26) G. Lokeshwari | G. Lokeshwari | | |
| (27) E. Aruna | E. Aruna | | |
| (28) K. Pooja | K. Pooja | | |
| (29) S. Sunil Kumar | S. Sunil Kumar | | |
| (30) P. Suresh | P. Suresh | | |
| (31) N. Suresh Babu | N. Suresh Babu | | |
| (32) P. Pavan Kalyan | P. Pavan Kalyan | | |
| (33) T. Suresh | T. Suresh | | |
| (34) M. Sandhya | M. Sandhya | | |
| (35) U. Vamsi | U. Vamsi | | |
| (36) V. Mohan | V. Mohan | | |
| (37) Y. Govardhana | Y. Govardhana | | |
| (38) Y. Venkata Subbiah | Y. Venkata Subbiah | | |
| (39) P. Sishayya | P. Sishayya | | |
| (40) G. Sai Teja | G. Sai Teja | | |
| (41) M. Narendra | M. Narendra | | |
| (42) V. Leela prasadh | V. Leela prasadh | | |
| (43) T.G.S. Kishore | T.G.S. Kishore | | |
| (44) P. Jashwanth | P. Jashwanth | | |
| (45) S. Abdul Samadh | S. Abdul Samadh | | |
| (46) P. Bhanu prakash | P. Bhanu prakash | | |
| (47) N. Muni Chaitanya | N. Muni Chaitanya | | |

Seminar on "Role of Statisticians - Census Survey and Analysis" by Prof. B. Sarojamma.



Staff and Students participated in the Seminar.



Seminar By Prof. B. Chaitanya.



3. Elocution & Essay writing competition

Elocution competition helps the students to develop their self expression and confidence. Elocution competition enable the students to improve their articulation and pronunciation of speech and oratory skills.

Writing encourages and enhances critical thinking and helps in organizing thoughts and improve communication skills. It provides an opportunity to the students to express their thoughts ,feelings,ideas and beliefs. In this direction the college organized elocution and essay competitions.




Students participating in Elocution competitions organized by the department of sanskrit for the academic year 2020-21 on the topic “Importance of Upanishad”.

4. Coaching for competitive examinations

The college has conducted coaching for competitive examinations to promote the employability of the students.

Phone: (0877) 2264599 E-mail: sgsartscollegettd@gmail.com Web: sgsac.edu.in


TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501



Prof N. VENUGOPAL REDDY, M.Sc., M.Phil., Ph.D
PRINCIPAL (FAC)

5.1.4.1: Number of students benefitted by guidance for competitive examinations and career counseling offered by the institution year wise during last five years

Sno	Academic Year	No of Students benefitted by Guidance for Competitive Examination and career counseling
1	2020-2021	NIL
2	2019-2020	363
3	2018-2019	296
4	2017-2018	290
5	2016-2017	282



P. Venkatesh
PRINCIPAL
S.G.S. ARTS COLLEGE
T.T.D., TIRUPATI.

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5.1.4.1: Number of students benefitted by guidance for competitive examinations and career counseling offered by the institution year wise during last five years

ACADEMIC YEAR 2019-2020

Sno	competitive examinations and career counseling	Name of the organizer	No of Students Participated
1	Coaching for Competitive Examinations	Career Guidance and Placement Cell, S.G.S. Arts College	104
2	PG CET Coaching	Department of Statistics	30
3	PG CET Coaching	Department of Mathematics	23
4	PG CET Coaching	Department of Chemistry	20
5	PG CET Coaching	Department of Physics	44
6	PG CET Coaching	Department of Commerce	88
7	PG CET Coaching	Department of Microbiology	7
8	PG CET Coaching	Department of Biotechnology	27
9	PG CET Coaching	Department of Economics	20
Total Number of Students Participated in competitive examinations and career counseling			363

N. Venugopal Reddy
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ACADEMIC YEAR 2018-2019

Sno	competitive examinations and career counseling	Name of the organizer	No of Students Participated
1	Coaching for Competitive Examinations	Career Guidance and Placement Cell, S.G.S. Arts College	100
2	PG CET Coaching	Department of Statistics	33
3	PG CET Coaching	Department of Mathematics	20
4	PG CET Coaching	Department of Chemistry	20
5	PG CET Coaching	Department of Physics	22
6	PG CET Coaching	Department of Commerce	46
7	PG CET Coaching	Department of Microbiology	10
8	PG CET Coaching	Department of Biotechnology	25
9	PG CET Coaching	Department of Economics	20
Total Number of Students Participated in competitive examinations and career counseling			296


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ACADEMIC YEAR 2017-2018

Sno	competitive examinations and career counseling	Name of the organizer	No of Students Participated
1	Coaching for Competitive Examinations	Career Guidance and Placement Cell, S.G.S. Arts College	95
2	PG CET Coaching	Department of Statistics	29
3	PG CET Coaching	Department of Mathematics	26
4	PG CET Coaching	Department of Chemistry	20
5	PG CET Coaching	Department of Physics	23
6	PG CET Coaching	Department of Commerce	41
7	PG CET Coaching	Department of Microbiology	14
8	PG CET Coaching	Department of Biotechnology	22
9	PG CET Coaching	Department of Economics	20
Total Number of Students Participated in competitive examinations and career counseling			290


PRINCIPAL
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Prof N. VENUGOPAL REDDY, M.Sc., M.Phil., Ph.D
PRINCIPAL (FAC)

5.1.4.1: Number of students benefitted by guidance for competitive examinations and career counseling offered by the institution year wise during last five years

ACADEMIC YEAR 2016-2017

Sno	competitive examinations and career counseling	Name of the organizer	No of Students Participated
1	Coaching for Competitive Examinations	Career Guidance and Placement Cell, S.G.S. Arts College	105
2	PG CET Coaching	Department of Statistics	28
3	PG CET Coaching	Department of Mathematics	19
4	PG CET Coaching	Department of Chemistry	20
5	PG CET Coaching	Department of Physics	23
6	PG CET Coaching	Department of Commerce	48
7	PG CET Coaching	Department of Microbiology	4
8	PG CET Coaching	Department of Biotechnology	15
9	PG CET Coaching	Department of Economics	20
Total Number of Students Participated in competitive examinations and career counseling			282


PRINCIPAL
S.G.S. ARTS COLLEGE
T.T.D., TIRUPATI.

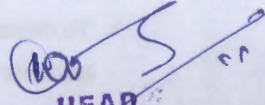
S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI
Student supporting programs / Events conducted
Name of the Department: Chemistry
Academic year:2019-20

S. No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	PGCET	13/12/2019	B.Sc : MPC,BBC,MZC&CBZ	20	-

HOD
HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

S.G.S. ARTS COLLEGE, TTD, TIRUPATI
Department of CHEMISTRY
EVENT ORGANISED REPORT: 2019 – 20

1	Name of the Department	Chemistry
2	Name of Event Organized	PGCET Coaching
3	Title of Event Organized	<u>PGCET Coaching – Chemistry</u>
4	Name of Co-coordinator of Event	Sri.D.Parameswara
5	Class of the Participants	B.Sc,MPC,BBC,MZC&CBZ
6	No. of Participants	20
7	Objective of the Event	An ability to develop critical thinking and efficient in problem solving skills
8	Outcome of the Event	To provide maximum scope for succeeding higher education


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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

TIRUMALA TIRUPATI DEVASTHANAM

S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501Prof N. VENUGOPAL REDDY,
M.Sc., M.Phil., Ph.D.
PRINCIPAL (FAC)**DEPARTMENT OF CHEMISTRY****CIRCULAR**

31-10-2019

S.G.S A.C / Chem / 2019 / DM / 04

The faculty members of department of chemistry are informed that there will be a staff meeting at 9.00 A.M in the HOD chamber on 01-11-2019 to discuss about P.G CET coaching to be conducted for the final year students for the academic year 2019-20.

S.No.	Name of the Faculty	Mobile No.	Signature
1	Sri D. Parameswari	9490728655	
2	Dr. V. Venkata Lakshmi	9492075161	
3	Dr. P. Suguna	7989925972	
4	Sri. K. Revathi	9177721917	

AGENDA

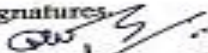
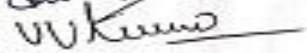
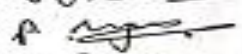
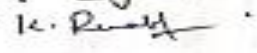
- To discuss about of P.G CET coaching to be conducted for the final year students for the academic year 2019-20.
- To finalize schedule for the P.G CET coaching
- To finalize timetable for the P.G CET coaching
- To discuss workload distribution for the P.G CET coaching


MINUTES

- ❖ The faculty members discussed about the schedule and time table for the P.G CET coaching and made the following resolutions unanimously.
- ❖ It was planned to conduct P.G CET coaching for the final year students in the second week of December 2019.
- ❖ Sri.D.Parameswara was made Coordinator for the P.G CET coaching.
- ❖ The workload distribution for the P.G CET was as shown below.
- ❖ It was planned to conduct P.G CET coaching classes at the rate of 3hrs per week on every Friday from 8.30 to 10 AM & 4.00 to 5.30 PM.

S.No	Name of the Faculty	Subject	No of Hours
1	Sri D. Parameswara	Physical chemistry	21
2	Dr. V. Venkata Lakshmi	Inorganic chemistry	20
3	Dr. P. Suguna	Organic chemistry	9
4	Smt.K.Revathi	Organic chemistry	10

Signatures

1. 
2. 
3. 
4. 


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Permission letter

02-11-2019

From

D. Parameswara,

HOD, Dept of chemistry,

S.G.S Arts College,

Tirupati.

To

The Principal,

S.G.S Arts College,

Tirupati.

Respected Sir,

Sub: Request for permission to conduct coaching classes for P.G
CET –Reg.

We have planned to conduct coaching classes for P.G CET in chemistry for the final year students of B.Sc with chemistry combination from the second week of December 2019. The coaching classes enable the students to perform better in P.G CET and other entrance examinations conducted for admission into M.Sc. So, please permit us to conduct the side coaching classes.

Thanking you

Yours faithfully


hel
HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

INTIMATION

SGSAC/Chem dept/PGCET 2019/HOD/01

04-11-2019

All the final year students of B.Sc with chemistry combination are informed that coaching for P.G CET has been scheduled to commence from 13-12-2019. The classes are conducted at the rate of 3hrs per week and every Friday from 8.30 to 10 AM & 4.00 to 6.00 PM, in room number 129. So, the interested students are instructed to enroll their name on or before 07-11-2019.


H.O.D of the Department
HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ANYS COLLEGE
TIRUPATHI

To

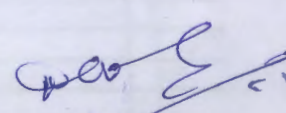
1. All teaching staff of chemistry
Department for circulation.
2. Department notice board.

DEPARTMENT OF CHEMISTRY
PGCET COACHING PLAN

PROGRAMME SCHEDULE : (WEEKLY 3HRS) VENUE : Room No.129

week	Date	Time	Topic	Faculty incharge
week-1	13/12/2019	8.30-10.30	S-block elements	Dr.V.Venkata lakshmi
		4.00-5.30	Structural theory in organic chemistry	Dr.B.Sathynarayana
	14/12/2019	8.30-10.00	P-Block elements	Dr.V.Venkata lakshmi
		4.00-5.30	Hydro carbons	Dr.B.Sathynarayana
week-2	20/12/2019	8.30-10.00	Organo metallic compounds	Dr.V.Venkata lakshmi
		4.00-5.30	Hydro carbons	Dr.B.Sathynarayana
	21/12/2019	8.30-10.00	d-block elements	Dr.V.Venkata lakshmi
		4.00-5.30	Hydro carbons	Dr.B.Sathynarayana
week-3	27/12/2019	8.30-10.00	f-block elements	Dr.V.Venkata lakshmi
		4.00-5.30	Halogen compounds ,Hydroxy compounds	Dr.B.Sathynarayana
	28/12/2019	8.30-10.00	Bonding in metals	Dr.V.Venkata lakshmi
		4.00-5.30	Carbonyl compounds	Dr.B.Sathynarayana
week-4	03/01/2020	8.30-10.00	Metal carbonyls	Dr.V.Venkata lakshmi
		4.00-5.00	Carboxylic acids	Dr.B.Sathynarayana
	04/01/2020	8.30-10.00	Co-ordination chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Carbohydrates	Dr.P.Suguna
week-5	10/01/2020	8.30-10.00	Co-ordination chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Amino acids & Nitro compounds	Dr.P.Suguna
	11/01/2020	8.30-10.00	Co-ordination chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Amino acids & Nitro compounds	Dr.P.Suguna
week-6	24/01/2020	8.30-10.00	Analytical chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Hetero cyclic compounds	Dr.P.Suguna
	25/01/2020	8.30-10.00	Metallurgy	Dr.V.Venkata lakshmi
		4.00-5.30	Stereo chemistry	Dr.P.Suguna

week-7	07/02/2020	8.30-10.00	Bioinorganic chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Stereo chemistry	Dr.P.Suguna
	08/02/2020	8.30-10.00	Gaseous state	Sri.D.Parameswara
		4.00-5.30	Solid state	Sri.D.Parameswara
week-8	14/02/2020	8.30-10.00	Solutions	Sri.D.Parameswara
		4.00-5.30	Surface chemistry & colloids	Sri.D.Parameswara
	15/02/2020	8.30-10.00	Electro chemistry	Sri.D.Parameswara
		4.00-5.30	Electro chemistry	Sri.D.Parameswara
week-9	21/02/2020	8.30-10.00	Phase rule	Sri.D.Parameswara
		4.00-5.30	Colligative properties of dilute solutions	Sri.D.Parameswara
	22/02/2020	8.30-10.00	Chemical kinetics	Sri.D.Parameswara
		4.00-6.00	Chemical kinetics	Sri.D.Parameswara
week-10	27/02/2020	8.30-10.00	Thermo dynamics	Sri.D.Parameswara
		4.00-5.30	Thermo dynamics	Sri.D.Parameswara
	29/02/2020	8.30-10.00	Photo chemistry	Sri.D.Parameswara
		4.00-5.30	Photo chemistry	Sri.D.Parameswara


 HEAD
 DEPARTMENT OF CHEMISTRY
 S. G. S. COLLEGE
 TIRUPATHI

HEAD
 DEPARTMENT OF CHEMISTRY
 S. G. S. COLLEGE
 TIRUPATHI

Year 2019-2020 PG CET → ADVANCED LEARNER

Pupils Attendance Register		పిల్లల హాజరుపట్టి	
TIME: 8.30 to 10 AM — 4 to 5.30 PM = 3 hours Per day			
1	2	3	4
Sl. No.	Name of the Student	Sl. No.	Sl. No.
1		1	2
2		3	4
3		5	6
4		7	8
5		9	10
6		11	12
7		13	14
8		15	16
9		17	18
10		19	20
11		21	22
12		23	24
13		25	26
14		27	28
15		29	30
16		31	
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18			
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28			
29			
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31			

DEPARTMENT OF CHEMISTRY
PGCET COACHING , STUDENT'S MARKS LIST

S.NO	Name of the student	Test-1	Test-2	Grand Tot
		08/01/2020 (Maximum marks)	09/01/2020 (Maximum marks)	08/01/2020 (Maximum marks)
1	MANDADHI GANESH	46	45	69
2	V PAVAN KALYAN NAIK	45	38	71
3	P PAVAN KALYAN	45	42	68
4	VAKITE NAIHESH BABU	39	41	AB
5	Y HEMANTHKUMAR	45	45	80
6	NEALLELA MALIKAIJUNA	42	49	73
7	BYRICKETTY HARI PRASAD	48	47	82
8	A K JAGADEESAN	43	45	81
9	Pathina nagaraju	49	43	79
10	YERUNALA HANUMANTHU	Ab	43	75
11	PEDDAULITHI RAVI SHANKAR	41	45	68
12	POOJARI HEMANATH	45	46	67
13	VELPULA NAGESWARA	47	49	83
14	Arigela Venkatesh	48	AB	85
15	ODUGURU SREERKANTH	43	45	71
16	TURRA SUNANTH NAIDU	41	43	85
17	P.VAMSIDHAR REDDY	42	45	66
18	B.VENKATUSHI	47	47	65
19	K.UMA	42	43	80
20	C.RENUKA	41	40	81

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S.G.S ARTS COLLEGE

STUDENT FEED BACK REPORT

STUDENT FEED BACK FORM:

Date:

Academic year: 2019-20

**Semester: VI
for PG CET**

Name of the programme : Coaching

Department :Chemistry

If you have attended our PG CET coaching classes we encourage you to submit feedback for us. This will help us to improve our future sessions and better meet the needs of our faculty.

S.NO	CRITERIA	RATING				
1	Content of the programme	Very poor	Poor	Average	Good	Excellent
2	Presentation style & delivery	Very poor	Poor	Average	Good	Excellent
3	Was it motivating	Very little	Less	Average	Lots	Too much
4	was the time given for questions	Very little	Less	Average	Lots	Too much
5	Was it useful for future	Very useless	Useless	Average	Useful	Very much
6	To What extent you have learnt from this programme	Very little	Less	Average	Lots	Too much
7	What is your all rating to this programme	Very poor	Poor	Average	Good	Excellent

General Comments:

It has motivated me to study well.

Students Signature:

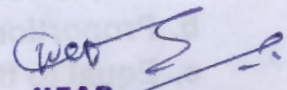
S.G.S ARTS COLLEGE :TIRUPATI
DEPARTMENT OF CHEMISTRY
PGCET COACHING FOR THE ACADEMIC YEAR 2019-20

NOTICE

14-12-2019

Students of PGCET coaching are informed that the test schedule for PGCET coaching is as shown bellow.

Test-1 : 05-01-2020
Test-2 : 09-02-2020
Grand test : 08-03-2020


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- c. Photosensitizer d, Promotor
11. In *Joule – Thomson* expansion ()
 a. $dS = 0$ b. $dG = 0$ c. $dE = 0$ D. $dH = 0$
12. The instrument used to measure optical rotation is ()
 a. Polari meter b. Spectrometer c. Calorimeter d. Voltmeter
13. Al_4C_3 , belongs to the category of ()
 a. Acetyl ides b. Allylides c. Mathanides d. Interstitial
14. The order of reactivity towards electrophilic substitution among benzene, pyrrole, furan and thiophene is ()
 a. Pyrrol>benzene>furan>thiophene
 b. Thiophene>benzene>pyrrol>furan
 c. Pyrrol>furan>thiophene>benzene
 d. Pyrrol>thiophene>furan>benzene
15. Which of the following cycle alkanes is the most stable? ()
 a. Cyclo propane b. Cyclo butane
 c. Cycle pentane d. Cycle hexane
16. Which of the following is the strongest acid? ()
 a. Phenol b. 2-Nitro phenol
 c. 2, 4 Dinitro phenol d. Picric acid
17. The trade name of trichloroethylene is ()
 a. Teron b. Westron c. Westrosol d. DDT
18. What metal ion is the oxidizing agent in tollens' reagent? ()
 a. Cu^+ b. Ag^+ c. Fe^{2+} d. Fe^{3+}
19. Which of the following compounds most acidic? ()
 a. CH_3Cl b. CH_3OH c. $H_2C = CH_2$ d. $H-C = C-H$
20. Which of the following cycloalkanes is most stable? ()
 a. Cyclopropane b. Cyclobutene
 c. Cyclopentane d. Cyclohexane
21. Which of the following compounds form only a mono substituted? ()
 Derivative on bromination ?
 a. N-butane b. Propane
 c. 2-methyl butane d. 2, 2-demithyl propane
22. The number of possible stereoisomers for 2, 3 dibromo butane are ()
 a. Two b. Four c. Three d. Six
23. Which of the following do not respond to 'Iodoform' test? ()
 a. Acetaldehyde b. Benzaldehyde
 c. Acetophenone d. Methylethyl ketone
24. Which of the following is a rearrangement reaction? ()
 a. Wurtz reaction b. Claisen condensation
 c. Pinacol-pinacoline reaction d. Sulphonation
25. Which of the following prove the wave nature of electron? ()
 a. Uncertainty principle b. Photoelectric effect

- c. Electron diffraction d. Zeeman Effect
26. When sodium chloride is added to a mixture of oil and caustic soda, the soap gets precipitated because ()
- NaCl is ionic compound
 - Of an increase in concentration of Na^+ ion
 - Of the insolubility of the soap in presence of Cl^- ions
 - Of decrease in the solubility product of NaCl in presence of the soap
27. For a gas the deviation from ideal behaviour is maximum at ()
- 0°C , 1 atmos
 - -10°C , 1 atmos
 - 0°C , 5 atmos
 - -10°C , 5 atmos
28. The degrees of freedom at eutectic point ()
- 0
 - 1
 - 2
 - 3
29. In which of the following processes, the entropy will decrease? ()
- Boiling of water
 - Freezing of a liquid
 - Expansion of a gas
 - Melting of ice
30. The equivalent conductance at infinite dilution is equal to ()
- The sum of the ionic conductances of cation and anion
 - The sum of the transport numbers of the cation and anion
 - The sum of the specific conductivity and cell constant
 - The sum of equivalent and specific conductivities
31. Inversion temperature is ()
- The temperature at which a gas liquefies
 - The temperature to which a gas should be cooled before applying pressure
 - The temperature above which the liquid cannot exist
 - The temperature at which both gas and liquid can coexist
32. Gold number means ()
- The number of gold atoms present in a standard gold sol
 - The minimum amount of an electrolyte which should be added to a Standard gold sol which just prevents its coagulation
 - The minimum amount of gold chloride which should be dispersed in water to give a stable gold sol
 - The number of moles of a protective colloid which should be added to a standard gold sol to prevent its coagulation
33. The two effects associated with the Debye-Huckel theory are: ()
- Relaxation effect and asymmetry effect
 - Relaxation effect and mesomeric effect
 - Relaxation effect and electrophoretic effect
 - Only mesomeric effect
34. The light emitted by a glowworm is an example of ()
- Fluorescence
 - Phosphorescence
 - Photosynthesis
 - Chemiluminescence
35. Which of the following ion has a magnetic moment of 3.87 B.M.? ()

- a. Mn^{2+} b. Fe^{3+} c. Cr^{3+} d. Sc^{3+}
36. The molecular shape of XeF_4 is ()
 a. Linear b. Square
 c. Square planar d. Pentagonal bipyramid
37. The strong field ligand is ()
 a. en b. NH_3 c. NO_2^0 d. CN^0
38. The most important air pollutants responsible for depletion of the Ozone layer are ()
 a. Oxides of nitrogen b. Oxides of sulphur
 c. Oxides of halogens d. Chlorofluoro hydrocarbons
39. Which of the following ions is colourless? ()
 a. Sc^{3+} b. Ti^{3+} c. Fe^{3+} d. V^{3+}
40. The molarity of pure water ()
 a. 55.6 b. 5 c. 100 d. 18
41. The half life of radon is 3.8 days. The time (in days) by which its 1/20 of the Amount will be left behind is ()
 a. 9 approx. b. 16.45 c. 7.6 d. 1.9
42. The function of haemoglobin in red blood cells is ()
 a. To transport dioxygen from lungs to tissues
 b. To convert CO_2 of the tissues (absorbed from environment) to HCO_2 ions
 c. The O_2 transforms haemoglobin to oxyhaemoglobin which is facilitated PH Fe^{2+} ions for storage of energy in lungs, but not in muscles
 e. The affinity for haemoglobin towards O_2 decreases with pH and ligands like CO, CN etc. reduce it to deoxyhaemoglobin, which is not good O^2 carrier
43. The most stable oxidation state of lanthanides is : ()
 a. +4 b. +2 c. +3 d. +1
44. The indicator employed in the complexometric titration of magnesium using E.D.T.A. is : ()
 a. Phenolphthalein b. Methyl Orange
 c. Erichrome Black – T d. Catechol Violet
45. According to Fajan's rules, the bond will acquire covalent character, if the cation is ()
 a. Large Size b. Small Size
 c. Largest Size d. Independent of its size
46. Which of the following compounds undergo cannizaro's reaction when treated with an alkali? ()
 (a) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CHO}$ (b) $(\text{CH}_3)_3\text{CHCHCH}_2\text{CHO}$
 (c) $\text{CH}_3\text{CH}_2\text{CH}(\text{CH}_3)\text{CHO}$ (d) $(\text{CH}_3)_3\text{C-CHO}$
47. Which of the following is six membered ring ()
 (a) Pyrrole (b) Pyridine
 (c) Galactose (d) Maltose
48. The order of molecularity of reaction of inversion of cane sugar respectively

is ()

- (a) 1,2 (b) 2,1 (c) 1,1 (d) 2,2

49. The quantum yield of H_2 and Br_2 reaction in the presence of light is ()

- (a) 1.0 (b) 0.1 (c) 0.01 (d) 10

50. During fermentation, the mixture is stirred well because ()

- (a) Of heating (b) CO_2 is liberated
(c) Yeast is present (d) Alcohol is obtained

ANSWERS									
1	d	11	d	21	d	31	b	41	b
2	b	12	a	22	C	32	d	42	c
3	d	13	c	23	D	33	d	43	c
4	d	14	c	24	C	34	d	44	c
5	b	15	d	25	C	35	c	45	b
6	a	16	d	26	B	36	c	46	b
7	d	17	c	27	a	37	d	47	b
8	c	18	b	28	a	38	d	48	a
9	c	19	d	29	b	39	a	49	c
10	a	20	d	30	a	40	a	50	b

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TEST-II

1. Nitrobenzene combines with hydrogen in the presence of platinum to yield (c)
(a) Aniline (b) Benzene
(c) Toluene (d) Azobenzene
2. Aromatic compounds undergo (d)
(a) Nucleophilic addition reactions
(b) Nucleophilic substitution reactions
(c) Electrophilic addition reactions
(d) Electrophilic substitution reactions
3. Which of the following bases is most likely to favour SN^1 reaction rather than SN^2 reaction in the hydrolysis of an alkyl halide? (c)
(a) OH^- (b) OEt^- (c) H_2O (d) NH_2^-
4. The natural amino acids are (a)
(a) D – rotatory (b) L - rotatory
(c) L-Configuration (d) D- Configuration
5. Camphor can be easily purified by the process of (c)
(a) Sublimation (b) Distillation
(c) Crystallisation (d) Sedimentation
6. The number of acidic hydrogen atoms in 1-butyne and 2-butyne respectively are (b)
(a) 1,0 (b) 0,1 (c) 1,1 (d) 1,2
7. HNC and HCN are (a)
(a) Functional isomers (b) Tautomers
(c) Chain isomers (d) Position isomers
8. HNC and HCN are (b)
(a) Functional isomers (b) Tautomers
(c) Chain isomers (d) Position isomers
9. Grignard's reagent reacts with ketones by addition or hydrolysis forming (c)
(a) Primary alcohol (b) Secondary alcohol
(c) Tertiary alcohol (d) Aldehyde
10. The reagent used for distinguishing alcohols is (b)
(a) Br_2/CCl_4 (b) Lucas reagent
(c) Schiff's reagent (d) $FeCl_3$
11. The product of destructive distillation of wood is (b)
(a) Acetic acid (b) Ethyl alcohol
(c) Soap (d) Kerosene
12. Ammonolysis of an ester leads to the formation of (b)

- (a)Uride (b)Amide
(c)Amine (d)Nitrite
13. Phenyl isocyanide test is used to identify (c)
(a) Aromatic secondary amines
(b) Aromatic tertiary amines
(c) Aromatic primary amines
(d) Quarternary ammonium compounds
14. The vanthoff factor i for a dilute solution of NaHSO_4 IS (d)
(a) $1/2$ (b) 1 (c) 2 (d) 3
15. Real gas exist at (b)
(a) high pressure and high temperature
(b) high pressure and low temperature
(c) low pressure and high temperature
(d) low pressure and low temperature
16. The electrolyte which is effective in coagulating gold sol is (c)
(a) Sodium chloride (b) Barium chloride
(c) Aluminium chloride (d) Ferric chloride
17. EDTA is a (d)
(a) bidentate ligand (b) tridentate ligand
(c) pentadentate ligand (d) hexadentate ligand
18. Chalcogens can exhibit oxidation states of (c)
(a) odd number (b) even number
(c) odd and even numbers (d) fractional numbers
19. The oxidation state of oxygen in O_2F_2 is (a)
(a) +1 (b) +2 (c) +4 (d) -2
20. An ester on hydrolysis gives a carboxylic acid which on Kolbe's electrolysis gives ethane (d)
(a) Ethyl methanoate (b) Ethyl propanoate
(c) Ethyl ethanoate (d) Methyl methanoate
21. Identify the alkane which gives only one product on chlorination (b)
(a) n - Pentane (b) Neopentane
(c) 2-Methyl butane (d) n - Butane
22. Reaction of benzene with n-propyl chloride in presence of AlCl_3 gives the following compound as the major product (b)
(a) n - Propyl benzene (b) Isopropyl benzene
(c) Chlorobenzene (d) Benzyl chloride
23. The carbon content present in Sesquiterpenes is (d)
(A) C_{10} (B) C_{20}
(C) C_{30} (D) C_{15}
24. The addition of H_2O to acetylene in presence of $\text{H}_2\text{SO}_4 + \text{HgSO}_4$ is an example for (A)
Electrophilic addition (B) Nucleophilic addition

- (C) Free radical addition (D) Electrophilic substitution
25. Aniline on reaction with bromine water gives (d)
 (A) o-Bromoaniline (B) p-Bromoaniline
 (C) 2,4-Dibromoaniline (D) 2,4,6- Tribromoaniline
26. The number of asymmetric carbons present in glucose is (d)
 (A) 1 (B) 2 (C) 3 (D) 4
27. Which of the following statements is correct in case of diastereomers? (c)
 (A) Identical physical and chemical properties
 (B) Mirror image relationship
 (C) No mirror image relationship
 (D) Optically inactive
28. The compound that reacts fastest with Leucas reagent is (d)
 (A) I-Butanol (B) 2- Butanol
 (C) 3- Methyl-2-butanol (D) 2- Methyl-2- butanol
29. Isocyanide test is used for the identification of (d)
 (A) Alcohols (B) Nitroalkanes
 (C) Aldehydes (D) Amines
30. Sugar gets charred with con.H₂SO₄. It is because of (b)
 (A) Dehydrogenation (B) Dehydration
 (C) Hydration (D) Hydrolysis
31. An aqueous solution containing Ni²⁺ and Cd²⁺ has been acidified and then H₂S gas passed into it. In this process. (b)
 (A) Both Ni²⁺ and Cd²⁺ get precipitated
 (B) Only Cd²⁺ gets precipitated
 (C) Only Ni²⁺ gets precipitated
 (D) Neither Ni²⁺ nor Cd²⁺ gets precipitated
32. The mean free path of the molecule of a gas is (b)
 (A) Directly proportional to pressure
 (B) Inversely proportional to pressure
 (C) Independent of pressure
 (D) Inversely proportional to square root of pressure
33. Which of the solution has the highest vapour pressure at a specified temperature? (b)
 (A) 0.1 m glucose (B) 0.1m NaCl
 (C) 0.1m CaCl₂ (D) 0.1m FeCl₃
34. The process in which an excited molecule in its triplet state returns to its ground state in its single state by emitting the radiation is known as (b)
 (A) Fluorescence (B) Phosphorescence
 (C) Photosensitization (D) Chemiluminescence
35. Which of the following is a double salt? (d)

- (A) $\text{CuSO}_4 \cdot 4\text{NH}_3 \cdot \text{H}_2\text{O}$ (B) $\text{Fe}(\text{CN})_2 \cdot 4\text{KCN}$
(C) $\text{CoCl}_3 \cdot 6\text{NH}_3$ (D) $\text{KCl} \cdot \text{MgCl}_2 \cdot 6\text{H}_2\text{O}$
36. $\text{K}_4[\text{Fe}(\text{CN})_6]$ is named according to IUPAC system as (c)
(A) Potassium hexacyanoferrum (II) (B) Potassium ferrocyanide
(C) Potassium hexacyanoferrate (II) (D) Potassium hexacyanoferrate
37. Which among the following is not coloured? (a)
(A) TiO_2 (B) VOSO_4 (C) TiCl_3 (D) CrCl_3
38. The oxidation state of S in peroxomonosulphuric acid (H_2SO_5) is (c)
(A) +2 (B) +4 (C) +6 (D) +8
39. The shape of $\text{Ni}(\text{CO})_4$ is (a)
(A) Tetrahedral (B) Trigonal
(C) Square planar (D) Trigonalbipyramidal
40. Cu^{2+} and Cd^{2+} ions can be separated from each other using (b)
(A) H_2S (B) KCN (C) HCl (D) None of these
41. The reaction of formaldehyde with organolithium compound followed by hydrolysis gives (a)
(A) Primary alcohol (B) Secondary alcohol
(C) Tertiary alcohol (D) None of these
42. A diatomic molecule has eight electrons in bonding orbitals, four electrons in non-bonding orbitals and two electrons in antibonding orbitals. The bond order present in it is (b)
(A) 1 (B) 2 (C) 3 (D) 4
43. Detergent have common group (d)
a. RNA b. RONa c. RCOONa d. $\text{RC}_6\text{H}_4\text{SO}_3\text{Na}$
44. Ethyl alcohol is denatured by (d)
a. Methanol and Formic acid
b. KCN
c. CH_3OH and C_6H_6
d. CH_3OH and Pyridine
45. Which is strongest acid? (d)
a. CCl_3COOH b. CH_3COOH
c. CBr_3COOH d. CF_3COOH
46. Product formed when benzaldehyde is refluxed with alcoholic KCN Solution (a)
a. Benzoin b. Phenyl cyanide
c. Phenyl isocyanate d. Phenyl isocyanide
47. The correct structure of H_2O_2 is (d)
a. H_2O_2 b. $\text{H}-\text{O}-\text{O}-\text{H}$
c. $\text{H}-\text{O}$ d. $\text{H}-\text{O}$

O-H

(b)

b. Cu, Zn, Ni

d. Zn, Ge, Ag

(c)

b. Thiocarbide

d. Thiocyanate

(c)

b. Ca

c. Mg

d. all

ANSWERS									
1	C	11	D	21	D	31	B	41	A
2	D	12	B	22	B	32	B	42	B
3	C	13	C	23	D	33	B	43	D
4	A	14	D	24	B	34	B	44	D
5	C	15	B	25	D	35	D	45	D
6	B	16	C	26	D	36	C	46	A
7	A	17	D	27	C	37	A	47	D
8	B	18	C	28	D	38	C	48	B
9	C	19	A	29	D	39	A	49	C
10	D	20	D	30	B	40	B	50	C

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GRAND TEST

SECTION – A

(100M)

1. Which of the following is not a chiral compound? [b]
(O) a. 2-bromopentane b. 3-chloropentane
c. 2-chlorohexane d. 3-bromohexane
2. A tri-substituted benzene derivative will have [a]
a. 2 isomers b. 4 isomers
c. 3 isomers d. 5 isomer
3. The most stable carbocation is [c]
a. CH_3^+ b. CH_3CH_2^+ c. $\text{C}_6\text{H}_5 - \text{CH}_2^+$ d. $\text{C}_3\text{H}_7\text{CH}_2^+$
4. Select the antiaromatic species from the following: cyclopropenyl cation, cyclopentadiene, naphthalene, benzene, pyridine, cyclobutadiene, cyclopentadienyl cation [d]
a. cyclopentadiene
b. cyclopentadiene & pyridine
c. cyclobutadiene & pyridine
d. cyclopentadienyl cation & cyclobutadiene
5. Identify the chiral α – amino acid from the following: [b]
a. glycine b. valine c. naproxen d. quinidine
6. Identify the most appropriate reaction that proceeds through carbene intermediate (from the following) [a]
a. Reimer – Tiemann reaction b. Diels – Alder reaction
c. Wittig reaction d. Grignard reaction
7. How many stereoisomers are possible for butane-2,3-dicarboxylic acid? [b]
a. 1 b. 2 c. 3 d. 4
8. Cannizzaro reaction requires [b]
a. formaldehyde and HCl b. benzaldehyde and NaOH
c. benzaldehyde and aniline d. benzaldehyde and Na_2SO_3
9. Trimethylbenzene will have the following number of isomers. [b]
a. Aldol condensation b. Cannizzaro reaction
c. Claisen condensation d. Beckmann rearrangement
10. Michael reaction is [c]
a. 1,2-addition of nucleophiles to ketones
b. 1,2-addition of nucleophiles to olefins
c. 1,4-addition of nucleophiles to α, β -unsaturated carbonyl compounds

- d. 1,4-electrophilic addition to α , β -unsaturated carbonyl compounds
11. Identify the most appropriate reagent for conversion of RCOOH to RCH_2OH from the following [a]
 a. NaBH_4 b. NaH c. BH_3 THF d. H_2 [Pd] C
12. Identify the reaction that most appropriately proceeds through a carbocation intermediate [b]
 a. Diels-Alder reaction b. Friedel-Crafts reaction
 c. Wittig reaction d. Hydroboration
13. Conversion of benzaldehyde to cinnamic acid is an example of [b]
 a. Grignard reaction b. Knoevenagel condensation
 c. Wolff rearrangement d. Mannich reaction
14. Identify the most appropriate reagent for conversion of cyclopentanone to cyclohexanone [c]
 a. Methyl iodide b. methane sulfonic acid
 c. diazomethane d. dimethyl sulphate
15. The structure of well-known pesticide DDT is [c]
 a. 1,1,1-trichloro-2,2-bis(*p*-chlorophenyl) ethane
 b. trimethylaluminium
 c. 1,1-dichloro-2,2-difluoroethane
 d. tetraethyl lead
16. The intermediate in the anti-Markonikov addition of HBr to an alkene is a [d]
 a. carbanion b. carbocation
 c. carbene d. free radical
17. Identify the reaction in which two C-C bonds are formed in a single step [b]
 a. Grignard reaction b. Diels-Alder reaction
 c. Michael reaction d. Friedel-Crafts reaction
18. Which of the following is a chiral molecule? [a]
 a. 3-hexanol b. 1-hexanal
 c. 1,6 hexanediol d. hex-3-en-1-ol
19. Identify the molecule that has the highest boiling point [d]
 a. n-decane b. 1-decane c. 1-decyne d. 1-decanol
20. Beckmann rearrangement is the conversion of [c]
 a. oxime to nitrile b. amine to acid
 c. nitro group to amine d. oxime to amide
21. The compound that exhibits geometrical isomerism is : [b]
 a. 1-butene b. 2-butene c. 1-butanol d. 2-butanol
22. A terpene among the following is : [b]
 a. D-xylose b. geraniol c. nicotine d. L-proline
23. At pH 7.0, amino acids exist as : [d]
 a. negatively charged species b. positively charged species

- c. uncharged species d. zwitterions
24. Molecular formula of an olefinic ketone containing a ring will be [c]
 a. $C_nH_{2n}O$ b. $C_nH_{2n+2}O$ c. $C_nH_{2n-4}O$ d. $C_nH_{2n-6}O$
25. Which of the following molecules has the lowest pK_s ? [c]
 a. phenol b. *p*-nitro phenol
 c. *p*-methyl phenol d. *p*-amino phenol
26. Which of the following statements is NOT true about S_N2 reactions: [b]
 a. The transition state is linear
 b. The configuration is inverted
 c. The intermediate is planar
 d. The reaction is second order
27. The maximum number of isomers that a molecule with 5 asymmetric centres can have is : [b]
 a. 5 b. 32 c. 64 d. 16
28. Which of the following properties is characteristic of only chiral molecules : [c]
 a. a sharp melting point b. strong UV-VIS absorption
 c. non-zero optical rotation d. high dipole moment
29. 19. Identify the terpene from the following [b]
 a. quinine b. menthol c. glucose d. cholesterol
30. The compound which would exhibit intramolecular hydrogen bonding is [a]
 a. *o*-hydroxybenzaldehyde b. *p*-hydroxybenzaldehyde
 c. *m*-hydroxybenzaldehyde d. *o*-hydroxytoluene
31. R-OH have higher boiling points than R-Cl because of [b]
 a. molecular shape b. hydrogen bonding
 c. molecular size d. dipole moment
32. Which is the correct order of basicity? [c]
 a. pyrrole > pyridine > Et_3N
 b. Et_3N > pyrrole > pyridine
 c. pyridine > Et_3N > pyrrole
 d. Et_3N > pyridine > pyrrole
33. Which of the following forms intramolecular hydrogen bond? [d]
 a. benzoic acid b. salicylic acid
 c. phenol d. *p*-nitrobenzoic acid
34. α -D-glucose and β -D-glucose are best described as [a]
 a. anomers b. epimers
 c. enantiomers d. conformers
35. Which of the following molecules does not have $4n+2$ π -electrons [c]
 a. benzene b. phenanthrene
 c. cyclobutadiene d. chlorobenzene
36. The most appropriate reagent for conversion of $RCOOH$ to RCH_2OH is [d]
 a. $NaBH_4$ b. $Na/EtOH$ c. $LiBH_4$ d. $LiAlH_4$
37. Which of the isomers of pentane has the highest boiling point? [a]
 a. $CH_3(CH_2)_3CH_3$ b. $(CH_3)_2CHCH_2CH_3$
 c. $(CH_3)_2$ d. All isomers have the same b.p.

38. A mixture of benzoic acid and naphthalene can be separated most easily by [d]
- a. sublimation b. distillation
c. fractional distillation d. crystallization
39. An ester RCOOR' is hydrolyzed. The products are RCOOH and R'OH. Which of the following is true? [b]
- a. Both O-atoms in the acid were present in the ester
b. Only the carbonyl O-atom of the ester is retained in the acid
c. Only the alkoxy O-atom of the ester is retained in the acid
d. O-atoms in the ester are scrambled in the reaction
40. When phenyl magnesium bromide is reacted with methanol, the product of the reaction is [c]
- a. phenol b. anisole c. benzene d. bromobenzene

SECTION – B

41. The hydrogen halide with the maximum dipole moment is
a. HF b. HCl c. HBr d. HI [a]
42. Which property among the following indicates weak intermolecular forces?
a. small molar volume b. low boiling point
c. large dipole moment d. high electrical conductivity [b]
43. An octahedral complex MA_4B_2
a. will have two constitutional isomers
b. will have two stereoisomers
c. can not show isomerism
d. will be optically active [b]
44. Which of the following is true about the S_N2 displacement reaction?
a. always proceeds with racemisation
b. always proceeds with retention of configuration
c. sometimes inversion, sometimes retention
d. always proceeds with inversion of ion [d]
45. The structure of X_4F_4 is
a. square pyramid b. tetrahedral
c. square planar d. octahedral [c]
46. Among the ions, Sc^{3+} , Ti^{3+} , Cr^{3+} and Fe^{3+} , which one is expected to give a colourless aqueous solution?
a. Sc^{3+} b. Ti^{3+} c. Cr^{3+} d. Fe^{3+} [a]
47. Which of the following molecules is linear?
a. CO_2 b. NO_2 c. SO_2 d. OF_2 [a]
48. The hydrogen atom transitions involving absorption from $n = 0$ to all other n correspond to
a. Lyman series b. Balmer series
c. Brackett series d. Paschen-Beck series [a]

49. The oxidation state of oxygen in H_2O_2 is [c]
 a. 0 b. +1 c. -1 d. -2
50. Water gas is an equimolar mixture of [c]
 a. CO_2 and N_2 b. CO and N_2
 c. CO and H_2 d. CO and H_2O
51. The IUPAC name for $[\text{Fe}(\text{CN})_6]^{3-}$ is [a]
 a. hexacyano ferrate (II) ion b. hexacyano ferrate (III) ion
 c. hexacyano iron (II) ion d. iron (III) hexacyanide ion
52. The catalyst used in Haber's process for the manufacture of ammonia contains [d]
 a. Pt b. Pd c. Ni d. Fe
53. Which of the following is a by-product of soap manufacture? [d]
 a. Phenol b. glucose c. glycol d. glycerol
54. Identify the most appropriate reagent for the reduction of benzoic acid to benzyl alcohol [a]
 a. NaBH_4 b. LiBH_4 c. LiAlH_4 d. DiBAL-H
55. Identify the most polar molecule from the following [d]
 a. dec-1-ene b. decanal
 c. decan-2-one d. decan-1-ol
56. Among the following, which is the most appropriate description for the characteristic of an electron? [c]
 a. particle b. wave
 c. both particle and wave d. neither particle nor wave
57. An example of benzene mechanism is : [a]
 a. $\text{PhCl} + \text{NH}_3 / \text{NH}_2^- \longrightarrow \text{PhNH}_2$
 b. $\text{PhH} + \text{HNO}_3 + \text{H}_2\text{SO}_4 \longrightarrow \text{PhNO}_2$
 c. $\text{PhNH}_2 + \text{HONO} \longrightarrow \text{PhN}_2$
 d. $\text{PhI} + \text{CuCN} \longrightarrow \text{PhCN}$
58. The splitting of nuclear energy levels in a magnetic field is known as [a]
 a. Zeeman effect b. Stark effect
 c. Mossbauer effect d. Cotton effect
59. Which of the following pairs of chemicals cannot be used to make hydrogen gas? [c]
 a. HPO_3 b. H_3PO_3 c. H_3PO_4 d. $\text{H}_4\text{P}_2\text{O}_7$
60. The transition state in $\text{S}_{\text{N}}2$ displacement reaction has the geometry [d]
 a. trigonal bipyramid b. tetrahedral
 c. octahedral d. planar
61. Which of the following chemicals are responsible for depletion of ozone layer [b]
 a. hydrocarbons b. chlorofluorocarbons

73. In diborane, there are [c]
 a. two 3c-4e BHB bonds b. four 3c-4e BHB bonds
 c. two 3c-2e BHB bonds d. four 3c-2e BHB bonds
74. The number of unpaired electrons in gaseous species of Mn^{3+} , Cr^{3+} and V^{3+} respectively are: [c]
 a. 4, 4 and 2
 b. 3, 3 and 2
 c. 4, 3 and 2
 d. 3, 3 and 3
75. In the standard Daniel cell, the flow is from [a]
 a. Zn electrode to the Cu electrode
 b. Cu electrode to the Zn electrode
 c. Zn electrode to the H_2 electrode
 d. Pt electrode to the Cu electrode
76. Which one of the following compounds shows an electron spin resonance (ESR) spectrum? [d]
 a. benzene b. zinc oxide
 c. methyl blue d. copper sulphate
77. The number of degrees of freedom of water at its triple point is [c]
 a. 3 b. 1 c. 0 d. 2
78. In a spontaneous process [a]
 a. both entropy and free energy increase
 b. both entropy and free energy decrease
 c. entropy decreases and free energy increases
79. Work done in an isothermal reversible expansion of a perfect gas from a volume V_i to volume V_f is [a]
 a. greater than that in an adiabatic reversible expansion from the same initial to final state
 b. equal to that in an adiabatic reversible expansion from the same initial to final state
 c. less than that in an adiabatic reversible expansion from the same initial to final state
 d. such data cannot be compared
80. The lock-and-key hypothesis was proposed by Emil Fischer to explain [a]
 a. enzyme catalysis b. photosynthesis
 c. respiration d. heredity
81. Entropy of activation of the Diels-Alder reaction will be [c]
 a. positive b. negative c. zero d. positive or negative
82. The total entropy change in a Carnot cycle is [a]
 a. zero b. positive c. negative d. infinite
83. When a non-volatile solute is added to a solvent what is the difference in vapor pressure expressed as a fraction of original vapor pressure equal to? [c]
 a. Mole fraction of solute in vapor phase
 b. Mole fraction of solvent in vapor phase
 c. Mole fraction of solute in liquid phase
 d. Mole fraction of solvent in liquid phase

84. The relation $dG = Vdp - SdT$ leads to which of the following Maxwell's equation?

[d]

a. $\left(\frac{\delta T}{\delta V}\right)_S = - \left(\frac{\delta p}{\delta S}\right)_V$

b. $\left(\frac{\delta T}{\delta p}\right)_S = - \left(\frac{\delta p}{\delta S}\right)_P$

c. $\left(\frac{\delta p}{\delta T}\right)_V = - \left(\frac{\delta S}{\delta V}\right)_T$

d. $\left(\frac{\delta V}{\delta T}\right)_P = - \left(\frac{\delta S}{\delta p}\right)_T$

85. In an exothermic chemical reaction as the temperature rises, the equilibrium

[a]

- a. shifts towards the reactants b. shifts towards the products
c. remain unchanged d. exhibits an anomalous behaviour

86. An emulsion is

[b]

- a. liquid dispersed in gas b. liquid dispersed in liquid
c. solid dispersed in liquid d. solid dispersed in gas

87. The van der Waals equation for 'n' mole of a gas is given by:

[c]

- a. $(P + a/V^2)(V-b) = nRT$
b. $(P + na/V^2)(V-nb) = nRT$
c. $(P + n^2a/V^2)(V-nb) = nRT$
d. $(P + na/V^2)(V-nb) = nRT$

88. If a solid sphere is reshaped into a solid cube, the surface area

[a]

- a. increases b. decreases
c. does not change d. becomes half

89. Viscosity of liquid increases with rise in temperature

[c]

- a. Decreases b. Constant
c. Increases d. Independent

90. A water drop is spherical in shape due to

[a]

- a. Surface tension b. poise
c. Viscosity d. Reflection

91. At critical temperature, the surface tension will be

[d]

- a. 1 b. 2 c. 3 d. 0

92. The unit of rate constant of Zero-order reaction is

[b]

- a. 5^{-5} b. mol L^{-1} c. $\text{L}^2 \text{mol}^{-2} \text{s}^{-1}$ d. $\text{L mol}^{-1} \text{s}^{-1}$

93. Chemical dissociation is

[b]

- a. exothermic b. reversible by endothermic
c. reversible d. endothermic

94. Spin multiplicity can be calculated by using

[b]

- a. $2S$ b. $2S+1$ c. $2S-1$ d. $2J+1$

95. The properties of a real gas coincide with those of an ideal gas at

[c]

- a. $p \rightarrow 0$ at any T b. $p \rightarrow 0$ at any Boyle temperature T_B
c. T_B and at any pressure d. $T \rightarrow \infty$ and $p \rightarrow 0$

96. Chemiluminescence involves the emission of [c]
a. UV radiation b. Infra-red radiation
c. Visible radiation c. All the above
97. The main function of the salt bridge is [d]
a. to allow ions to go from one cell to another
b. to provide link between the half cells
c. to keep the emf of the cell positive
d. to maintain electric neutrality of the solution in two half cells
98. Transport number of CH_3COO^- ions is greater than that of $\text{C}_6\text{H}_5\text{COO}^-$ because [a]
 CH_3COO^- ion
a. has lower ionic mass b. is smaller in size
c. has CH_3 in place of C_6H_5 d. all the above
99. An acidic buffer solution can be prepared by mixing solutions by [b]
a. ammonium acetate and acetic acid
b. ammonium chloride and ammonium hydroxide
c. sulphuric acid and sodium sulphate
d. sodium chloride and sodium hydroxide
100. If the relative lowering of pressure of o-xylene is 0.005 due to addition of 0.5 grams of non-volatile solute in 500 grams of solvent, what is the molecular weight of the solute? [a]
a. 21.3 g/mole b. 23.1 g/mole c. 32.1 g/mole d. 1.23 g/mole

5. Group discussions

The institute encourage the advance learners to participate in group discussions which enhance their critical thinking, Understanding ability, Communication skills management skills through the use of roles within the groups and assessment skills involved in assessing options to make decisions about their group's final answers.



Students engaging in group discussion on TMV (Tobacco Mosaic virus) conducted by Department of Microbiology on 09-12-2021.

Department of Zoology

S G S ARTS COLLEGE, TIRUPATI

DEPARTMENT OF ZOOLOGY

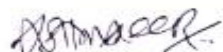
2018-19

GROUP DISCUSSION CONDUCTED

ON

EVOLUTIONARY CHANGES IN HEART IN CHORDATES

Group Discussion		
Academic Year 2018-2019		
S No	Names	Details
1	Name of the Department	Zoology
2	Date of the Group Discussion conducted	Sep 29, 2018
3	Name of Organizers	S N Shameer, B Ravi
4	Class of the Students	B.Sc., (CBZ & MZC)
6	Topic	Evolutionary changes in heart in chordates
7	No of Students attended	30
8	No of Faculty attended	2
9	Feedback	To know how the heart has evolved in the chordates.



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TIRUPATI-517 501

S G S ARTS COLLEGE, TIRUPATI
DEPARTMENT OF ZOOLOGY
GROUP DISCUSSIONS

Date: 14-9-18

CIRCULAR ISSUED TO THE STUDENTS

This is to inform that the department is intended to conduct group discussion on faculty preferred topic. In this regard, I hereby request all the B.Sc. CBZ/MZC Zoology students that those who are interested to participate in Discussion can enrol their names in the Zoology staff room.



HEAD OF THE DEPARTMENT,
ZOOLOGY.

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TIRUPATI-517 502

GROUP DISCUSSION CONDUCTED ON
EVOLUTIONARY CHANGES IN HEART IN CHORDATES



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S.G.S. Arts College
TIRUPATI-517 501

S.G.S ARTS COLLEGE ,T.T.D,TIRUPATI.

DEPARTMENT OF ZOOLOGY 2018-19

The students who have participate in the Group Discussion

S.NO	ROLL NO	NAME OF THE STUDENT	SIGNATURE
1	18701	A.BALAJI	Balaji
2	18702	C.BHANU VARA PRASAD	Bhanu vara prasad
3	18703	G.BHEEMUDU	Bheemudu
4	18704	T.CHAKRAPANI	chakrapani
5	187-5	G.CHANDR SEKHAR	chandr sekhar
6	18706	Y.DINESH BABU	Dinesh babu
7	18707	B.GANESH	Ganesh
8	187-8	R.GAYATHRI	Gayathri
9	18709	K.GOPAL	Gopal
10	18710	T.GOWTHAM	Gowtham
11	18711	B.GOWTHAM GOWDA	Gowtham gowda
12	18712	M.GURRAMMA	Gurramma
13	18713	K.KALYANI	kalyani
14	18714	K.LKITHA	Likitha
15	18715	T.MOUNIKA	mounika

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16	19673	M.NAVEEN NAIK	Naveen Naik
17	19674	G.RAJASEKHAR	Raja sekhar
18	19675	A.M.SAMEERA	Sameera
19	19676	P.SANDEEP KUMAR	Sandeep kumar.
20	19677	P.SANTOSH KUMAR	Santosh kumar.
21	19678	K.SASI KUMAR	Sasi kumar.
22	19679	E.SREEKANTH	Sreekanth.
23	19680	D.SURYA	Surya.
24	19681	K.SWATHI	Swathi
25	19682	K.USHA	Usha.
26	19683	G.VEERANJINEYULU	veeranjineyulu.



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35	19773	C.RAMA KRISHNA	Rama krishna
36	19774	A.RAMESH	Ramesh
37	19775	P.REDDY PRAKASH	Reddy prakash
38	19776	G.REDDI SEKHAR NAIDU	Reddishakar naidu
39	19777	REENA SHAIK	Reena shalk
40	19778	V.SAI CHAITANYA	sai chaitanya
41	19779	G.SIVA KUMAR	Siva kumar
42	19780	K.SOHITH KUMAR	sohith kumar
43	19781	C.SURESH BABU	Suresh babu
44	19782	N.USMAN	Usman
45	19783	P.VAISHNAVI	Vaishnavi
46	19784	C.VENKATA RAMANA	Venkata ramana
47	19785	K.VINOD	Vinod

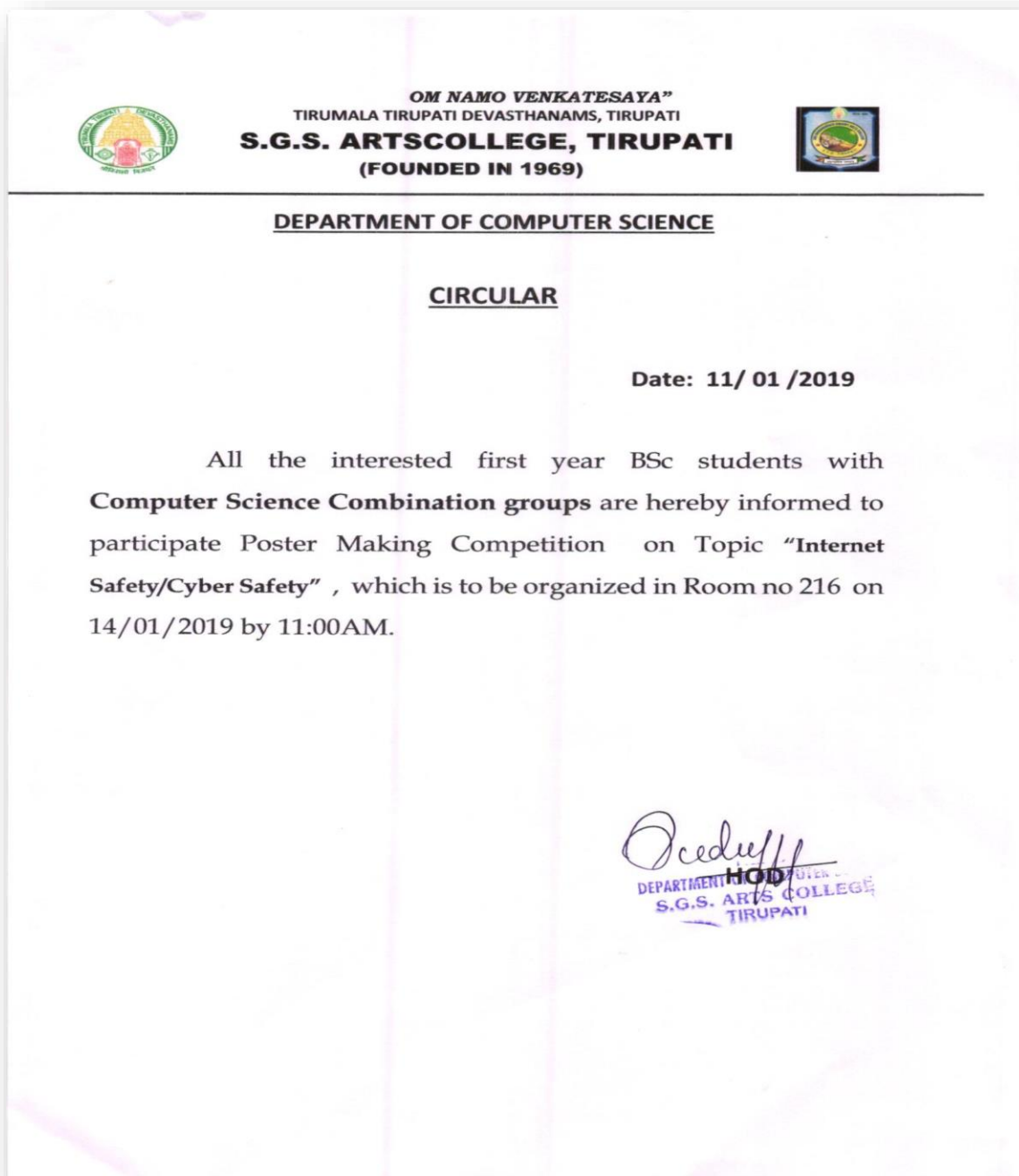

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Department of Computer Science

Academic Year : 2019 -2020

Activity : Poster Making Competition organized by Department of Computer Science

Date:14-01-2019





OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

Sri Govindaraja Swamy Arts College,TTD,Tirupati
Department of Computer Science

POSTER MAKING COMPETITION

Date:14-01-2019

OBJECTIVES:

1. To bring out the hidden talents in Poster designing .
2. To explore creative expression of the students and also to gauge their awareness quotient.
3. To improve overall performance of students by using different pedagogical techniques.
4. Ability to write simple tips for staying safe online in the form of poster thereby creating awareness to all on specified topic .

CONVENOR

- Sri P.Udaya, HOD, Department of Computer Science

ORGANIZING MEMBERS

- Smt B.Triveni
- Smt N.Jayalakshmi
- Smt D.Chaitanya Kumar

TOPIC: Internet Safety / Cyber Safety

ABSTRACT:

The department of computer science thought to organize Poster Making Competition programme for I BSc. MSCs-EM and MPCs-EM on the Topic **"Internet Safety/Cyber Safety"** to create awareness of personal safety and security risks to private information and property associated with using the internet, and the self-protection from computer crime.

PHOTO GALLERY

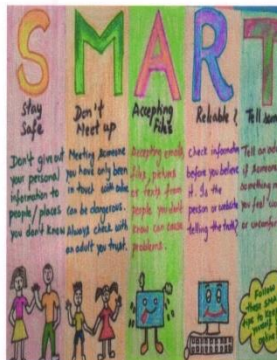
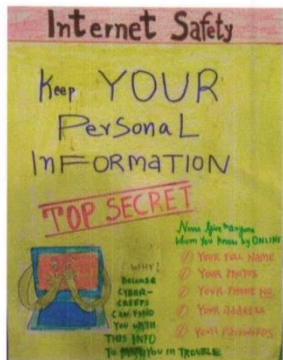
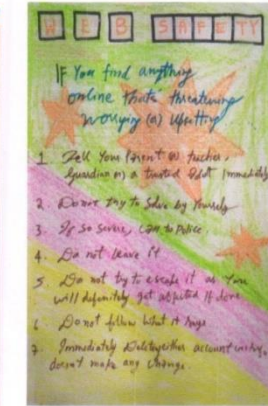
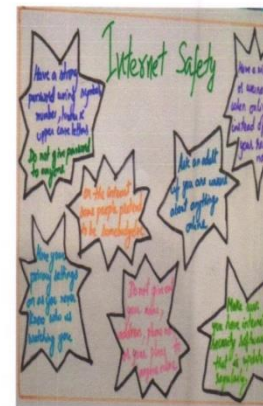
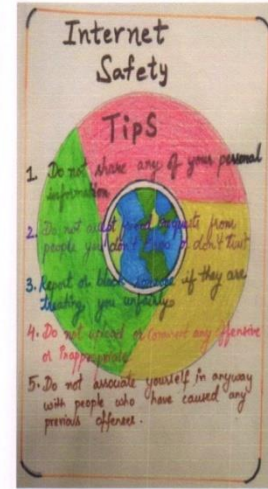
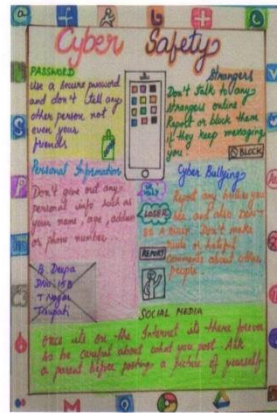






Posters Gallery

POSTERS GALLERY



Sri Govindaraja Swamy Arts College, TTD, Tirupati

Department of Computer Science

POSTER MAKING COMPETITION

ATTENDANCE SHEET

Sno	Roll No	Group & Year	Name of the Student	Signature
1	18520	MPC6 I st year	B. Harshavardhan Reddy	B. Harshavardhan Reddy
2	18565	MSC6 I st year	B. Ajay Kumar	B. Ajay Kumar
3	18568	MSC6 I st year	T. Anitha	T. Anitha
4	18513	MPC6 I st year	E. Balaji	E. Balaji
5	18569	MSC6 I st year	B. Aravind Naidu	B. Aravind Naidu
6	18575	MSC6 I st year	V. Chanidini	V. Chanidini
7	18515	MPC6 I st year	A. Balarama Krishna	A. Balarama Krishna
8	18514	MPC6 I st year	M. Balanijayulu	M. Balanijayulu
9	18576	MSC6 I st year	M. Charan Teja	M. Charan Teja
10	18581	MPC6 I st year	V. Govardhan Krishna	V. Govardhan Krishna
11	18577	MSC6 I st year	K. Deepa	K. Deepa
12	18516	MPC6 I st year	M. Charan Kumar	M. Charan Kumar
13	18519	MPC6 I st year	P. Varma guru sai Kumar	P. Varma guru sai Kumar
14	18521	MPC6 I st year	P. Hemant Kumar	P. Hemant Kumar
15	18584	MSC6 I st year	P. Karitha	P. Karitha
16	18586	MSC6 I st year	L. Indhumathi	L. Indhumathi
17	18588	MSC6 I st year	K. Kinshora	K. Kinshora
18	18589	MSC6 I st year	M. Lakshmi	M. Lakshmi
19	18523	MPC6 I st year	N. Jagannoud	N. Jagannoud
20	18524	MPC6 I st year	K. Jayalakshmi	K. Jayalakshmi
21	18595	MSC6 I st year	A. Panchala Naranimhola	A. Panchala Naranimhola
22	18596	MSC6 I st year	G. Ramesh	G. Ramesh
23	18598	MSC6 I st year	K. M. Rithik	K. M. Rithik
24	18602	MSC6 I st year	C. Sreedevi	C. Sreedevi
25	18525	MPC6 I st year	G. Habibu Shaif	G. Habibu Shaif

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Sno	Roll No	Group & Year	Name of the Student	Signature
26	18605	MSC6 I st year	S. Himmaiah	S. Himmaiah
27	18606	MSC6 I st year	K. Vamsi Krishna	K. Vamsi Krishna
28	18607	MSC6 I st year	M. Venkata Anand	M. Venkata Anand
29	18529	MPC6 I st year	G. MOUNIKA	G. MOUNIKA
30	18533	MPC6 I st year	K. Naren Kumar	K. Naren Kumar
31	18591	MPC6 I st year	Saiteja	Saiteja
32	18546	MPC6 I st year	Shaik Yamin	Shaik Yamin
33	18563	MPC6 I st year	G. Surenndra	G. Surenndra
34	18542	MPC6 I st year	D. Sandhya	D. Sandhya
35	18613	MSC6 I st year	E. VIJAY GOUD	E. VIJAY GOUD
36	18553	MPC6 I st year	Mangali Swathi	Mangali Swathi
37	18615	MPC6 I st year	P. Geethanjali	P. Geethanjali
38	18554	MPC6 I st year	V. SWAPNA	V. SWAPNA
39	18557	MPC6 I st year	B. Vamsi	B. Vamsi
40	18598	MPC6 I st year	K. Phanakumar	K. Phanakumar
41	18558	MPC6 I st year	P. Parshini	P. Parshini
42	18559	MPC6 I st year	K. Venkata Jagath	K. Venkata Jagath
43	18547	MPC6 I st year	G. Shireesha	G. Shireesha
44	18560	MPC6 I st year	M. Venkata Anand	M. Venkata Anand
45	18561	MPC6 I st year	M. VIGAY KUMAR	M. VIGAY KUMAR
46	18562	MPC6 I st year	A. Venkata Chaitanya	A. Venkata Chaitanya

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DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Prize winners :

- 1) K. Jayalakshmi - Ist Prize
- 2) B. Harshavardhan Reddy - 2nd Prize
- 3) G. Mounika - 3rd Prize

6. Expert Talks /Guest lecturers/Seminars.

Dept. of Physics organized a seminar on date 21.11.2019 on Analytical Techniques for material characterization, delivered by Prof. P.Sreedhar Reddy, Registrar, S.V.University, Tirupati.





Dept. of Physics organized a seminar on date 15.06.2016 on Nano Science and Nano Technology, delivered by Prof. OMD Hussain, Dept. of Physics, S.V.University, Tirupati.





Dept. of Physics organized a seminar on date 13.10.2016 on Atmospheric observations on Troposphere using Radar system, delivered by Prof. D. Narayana Rao, Former Director, GST Radar Centre, Gadanki.



7. STUDENTS' OFFICE BEARERS (STUDENTS' COUNCIL)

The student office bearers work under the guidance of the college. The main function of the students office bearers is to organize various activities to provide opportunities to the students of the college to exhibit their talents.

Objectives:

True education aims at holistic and integrated development of an individuals personality including intellectual , physical, emotional, moral and ethical aspects in search of truth.

Since the students are the central force of any educational institution, their participation in its activities undoubtedly essential within the overall perspective stated above. Student office bearers for various association are nominated on merit basis as per S.V.University guidelines,

Chairman, Secretary and Joint secretary are selected on rotation basis from B.A., B.Com and B.Sc courses . Secretary and Joint secretary for Maths and Science, Arts, Commerce, Language Biological Sciences, Fine arts and Sports are nominated on merit basis.

Institution facilitates students' representation and engagement in various administrative, co-curricular and extracurricular activities following duly established processes and norms (student council, students representation on various bodies)

STUDENT UNION OFFICE BEARERS: 2020-2021

2020 - 2021 – STUDENTS UNION

Union Chairman : **M. Gurramma**, III B.Sc., M.Z.C., (E.M.)

Union Secretary: **A. Anusha** III B.A., M.E.S., (E.M.)

MATHS & SCIENCE ASSOCIATION

Secretary : **B. Harshavardhan Reddy**, III B.Sc., M.P.Cs. (E.M.)

Joint Secretary : **P. Sakeel**, II B.Sc., M.P.Cs., (E.M.)

LIFE SCIENCE ASSOCIATION

Secretary : **R. Gayatri**, III B.Sc., M.Z.C. (E.M.)

Joint Secretary : **A. Chandini**, II B.Sc., B.B.C. (E.M.)

COMMERCE ASSOCIATION

Secretary : **D. Srikanth**, II B.Com., CA (E.M.)

Joint Secretary : **V. Yugandhar**, II B.Com., CA (E.M.)

ARTS ASSOCIATION

Secretary : **P. Ganesh Achary**, III B.A., A.S.C.A., (E.M.)

Joint Secretary : **V. Pavan Kumar**, II B.A., H.P.T. (T.M.)

LANGUAGE ASSOCIATION

Secretary : **E. Vijay Goud**, III B.Sc., M.S.Cs. (E.M.)

Joint Secretary : **D. Divyasri**, II B.Com., CA (E.M.)

FINE ARTS ASSOCIATION

Secretary : **M. Prabhu Kumar**, III B.A., H.E.P. (E.M.)

Joint Secretary : **A. Karunya**, II B.A., H.E.P. (E.M.)

SPORTS ASSOCIATION

Sports & Games_Secretary: **P. Veera Guru Saikumar**,
III B.Sc., M.P.Cs (E.M.)

*** * ***

2019-2020

Sno	Class	Name of the Representative	Roll No
1	III Bcom-CA-E.M	Kum. K.Deepa(Union Chairman)	17167
2	III B.Sc-MSCs-EM	Mr. A.Venkatesh(Union Secretary)	17582
3	III B.Sc-MPCs-EM	Mr. B.Mohana Krishna(Maths & Science Association Secretary)	17508
4	II B.Sc-MSCs-EM	Kum. K.Sireesha(Maths & Science Association Joint Secretary)	18599
5	III BSc-BBC-E.M	Kum. Jahnvi Takkolu (Life Sciences Association Secretary)	17672
6	II BSc-MZC-EM	Kum. M.Guramma(Life Sciences Association Joint Secretary)	18712
7	III BCom-CA-E.M	Mr. V.Sravan Kumar(Commerce Association Secretary)	17208
8	II BCom-CA-EM	Mr. D.Sreekanth(Commerce Association Joint Secretary)	18178
9	III BA-HEP-TM	Mr. K.Rajendra Prasad(Arts Association Secretary)	17307
10	II BA-MES-EM	Kum. A.Anusha(Arts Association Joint Secretary)	18373
11	III BSc-MPCs-EM	Kum. K.JyotsnaLanguage Association Secretary)	17501
12	II BSc-MPCs-EM	Kum. Shaik Yashmin(Language Association Joint Secretary)	18546
13	III BA-MES-EM	Mr. Y.Anil(Fine Arts Association Secretary)	17391
14	II BCom-GEN-EM	Kum. P.Yamini(Fine Arts Association Joint Secretary)	18101
15	I BA-HEP-EM	Kum. K.Karunya(Fine Arts Association Joint Secretary)	HT: 0120008001
16	III BCom-GEN-E.M	Mr. Y.Prudviraj(Sports&Games Association Secretary)	17112
17	III B.Sc-MZC-EM	Mr. D.Sandeep(Sports&Games Association Secretary)	17741

BSc Class Representatives			
Sno	Class	Name of the Representative	Roll No
1	III B.Sc-MSCs-EM	Mr. A.Venkatesu	17582
2	III B.Sc-MPCs-EM	Mr. M.Prasanth	17515
3	III B.Sc-MPC-EM	Mr. P.Pavan Kalyan	17473
4	III B.Sc-MPC-TM	Mr. Y. Hemanth Kumar	17428
5	III B.Sc-MPS-TM	Mr. O.Jayakumar	17598
6	III B.Sc-BBC-EM	Mr. B.Mahesh Babu	17681
7	III B.Sc-MZC-EM	Mr. P.Devanadh Goud	17719
8	III B.Sc-CT&HM-EM	Kum.Pavani Krishna Benguluru	17782
9	III B.Sc-CBZ-TM	Mr. V.Venkatesh	17644
10	II B.Sc-MSCs-EM	Mr. Bharath kumar	18571
11	II B.Sc-MPCs-EM	Mr. B.Harshavardhan Reddy	18520
12	II B.Sc-MPC-EM	Mr. Obul Reddy	18453
13	II B.Sc-MPC-TM	Mr. P.Chandra Sekhar	18405
14	II B.Sc-MPS-TM	Mr. O.Jayakumar	HT: 0318008134
15	II B.Sc-BBC-EM	Mr. M.Mareepa	18670
16	II B.Sc-MZC-EM	Mr. V.Pavan Kumar Goud	18718
17	II B.Sc-CT&HM-EM	Mr. Chandu Chainuru	18756
18	II B.Sc-CBZ-TM	Mr. A.Ashok	18616
19	I B.Sc-MSCs-EM	Mr. G. Dasari Gopal	HT: 0320008161
20	I B.Sc-MPCs-EM	Mr. K.Adimuni Peddaiah	HT: 0320008099
21	I B.Sc-MPC-EM	Mr. A.M.Raghuram Kumar	HT: 0320008026

22	I B.Sc-MPC-TM	Mr. A.M. RAGHURAM KUMAR	HT: 0320008040
23	I B.Sc-MPS-TM	Mr. G.Reddi Prasad	HT: 0320008057
24	IB.Sc-BBC-EM	Mr. M.Vivek Amruth Narayana	HT: 0320008282
25	I B.Sc-MZC-EM	Mr. A..Karthik	HT: 0320008214
26	I B.Sc-CT&HM-EM	Mr. V. Durga Nagendra	HT: 0320008346
27	I B.Sc-CBZ-TM	Mr. G. Rajasekhar	HT: 0320008012
BCom Class Representatives			
28	III B.Com-CA-EM	Mr. M.Anji Naik	17162
29	III B.Com-GEN-EM	Mr. C.Ramesh	17041
30	III B.Com-GEN-TM	Mr. Murali Chinthala	17032
31	II B.Com-CA-EM	Mr. M. Venkata Suresh	18187
32	II B.Com-GEN-EM	Miss . P.Yamini	18101
33	II B.Com-GEN-TM	Mr. Y.Sai Kumar	18042
34	I B.Com-CA-EM	Mr. G.Tarun	HT: 0220008143
35	IB.Com-GEN-EM	Mr. E.Mohan	HT: 0220008031
36	I B.Com-GEN-TM	Mr. M.Ravi	HT:0220008070
BA Class Representatives			
37	III B.A-HEP-EM	Mr. Hemanth Vuranduru	17345
38	III B.A-HEP-TM	Mr. Venkatanarayana Reddy Devireddy	17315
39	III B.A-HPT-TM	Mr. T.Eswaraiah	17367
40	III B.A-ASCA-EM	Mr. Vamsi Gundlapalli	17253
41	III B.A-MES-EM	Pradeep Madaka	17399
42	II B.A-HEP-EM	Mr. Y.Surya Mahesh	18336

43	II B.A-HEP-TM	Mr. A.Kiran Kumar	18273
44	II B.A-HPT-TM	Mr. M.Eswar Naik	18346
45	II B.A-ASCA-EM	Mr. A.Adisankar	18201
46	II B.A-MES-EM	Dilhi Babu.G	18374
47	I B.A-HEP-EM	Mr. A.Hari Prasad Reddy	0120008002
48	I B.A-HEP-TM	Mr. N.Siddeswaraiyah	0120008041
49	I B.A-HPT-TM	Mr. G.Nagaraju	0120008069
50	I B.A-ASCA-EM	Mr. J.Vinay Kumar	0120008086

HOSTEL BEARERS: 2019-2020

Sno	Committee Members	Designation	Role
1	Dr. C.Ramesh	Principal	Chief Warden
2	Dr.Y.Mallikarjun Rao	HOD,Commerce	Warden
3	Sri K. Rajesh	Lecturer,Economics	Deputy Warden
4	Mr. Prasanth	Student,III BSc-MPCs-EM	Mess Secretary
5	H.Mahadeva	Student,III BA-HEP	Store Secretary

MESS COMMITTEE

Sno	Name	Group	Role
1	P.Pavan Kalyan	III BSc-MPC-EM	Associate Members
2	D.Naveen	III Bcom-General	
3	G.Vamshi	III BA-ASCA	
4	G.Varaprasad	III BSc-CT&HM	
5	B.Harsha Vardhan	II BSc-MPCs	
6	B.Srikanth	II Bcom -General	
7	Ganesh Achari	II BA-ASCA	



[Signature]
PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

1. Student Orientation Programme

Objectives:

To help new students to adjust and feel comfortable in new environment.

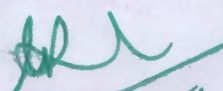
- To inculcate the culture of the college
- To make them aware of the curriculum
- To explain discipline regulations of the college
- To make the students to get familiarization with the college, departments
- To make the students to know extra-curricular activities in the college

2019-2020

Brief Report of Programme

Name of the Programme	Student Induction Programme for B.Sc. First Year
Date of the Programme	19-06-2019
Name of the Organizing Committee	Admission Committee
Objectives for organising the Programme	To create awareness among the students about academic curriculum, College discipline and activities
Nature of Programme	Lectures and Q & A
Brief Report	Principal Dr. C. Ramesh delivered a speech on College activities and discipline
No. of participants	102




PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

INDUCTION PROGRAMME -2019-20

SL NO	Name of the Student	Group	Med	SIGNATURE
1	Balasubramanyam Palli	B.Sc. (MPC)	EM	P. Balasubramanyam
2	Balasundaram Parampogu	B.Sc. (MPC)	EM	P. Balasundaram
3	Chandana Priya Esukapalli	B.Sc. (MPC)	EM	E. Chandana Priya
4	Chandu Prakash Avulapalli	B.Sc. (MPC)	EM	A. Chandu Prakash
5	Chinthalapattadai Chandran Uday Ku	B.Sc. (MPC)	EM	U. Chinthalapattadai
6	Tharun Kumar Mailay	B.Sc. (MPC)	EM	Tharun Kumar
7	Varaprasad Boya	B.Sc. (MPC)	EM	Varaprasad
8	Varaprasad Cherukuri	B.Sc. (MPC)	EM	Varaprasad
9	Venkata Arjun Amirisetty	B.Sc. (MPC)	EM	A. Venkata Arjun
10	Venkatesh Yanati	B.Sc. (MPC)	EM	Y. Venkatesh
11	Vishnu Prasad Kammapatti	B.Sc. (MPC)	EM	K. Vishnu Prasad
12	Vishnu Vardhan Talari	B.Sc. (MPC)	EM	Vishnu Vardhan
13	Akhil Kothapalli	B.Sc. (MPS)	EM	K. Akhil
14	Anil Kumar Kedri	B.Sc. (MPS)	EM	K. Anil Kumar
15	Ashok Mundla	B.Sc. (MPS)	EM	Ashok
16	Badri Penchala	B.Sc. (MPS)	EM	P. Badri
17	Bharath Kumar Amilineni	B.Sc. (MPS)	EM	Bharath Kumar
18	Chethan S	B.Sc. (MPS)	EM	S. Chethan
19	Teja Muneswar Vulasala	B.Sc. (MPS)	EM	Teja Muneswar Vulasala
20	Varma Teja Mondla	B.Sc. (MPS)	EM	M. Varma Teja
21	Venu Gopal Varikunta	B.Sc. (MPS)	EM	V. Venu Gopal
22	Vignesh Obalavaram	B.Sc. (MPS)	EM	Vignesh
23	Adimuni Peddalaiah Kattubadi	B.Sc. (MPCS)	EM	K. Adimuni Peddalaiah
24	Govardhan Manchala	B.Sc. (MPCS)	EM	M. Govardhan
25	Harathi Kalyani	B.Sc. (MPCS)	EM	Harathi Kalyani
26	Hari Krishna Singamala	B.Sc. (MPCS)	EM	S. Hari Krishna
27	Hema Neerugatti	B.Sc. (MPCS)	EM	N. Hema
28	Hemalatha Aruru	B.Sc. (MPCS)	EM	Hemalatha Aruru
29	Pavan Kalyan Junuguru	B.Sc. (MPCS)	EM	J. Pavan Kalyan
30	Pavan Kumar Pulimadugu	B.Sc. (MPCS)	EM	P. Pavan Kumar
31	Pavan Kumar Reddy Katam	B.Sc. (MPCS)	EM	P. Pavan Kumar
32	Pedda Swamy Dudekula	B.Sc. (MPCS)	EM	Pedda Swamy
33	Praveen Kolle	B.Sc. (MPCS)	EM	P. Praveen
34	Veerendrachari Kamhari	B.Sc. (MPCS)	EM	K. Veerendrachari
35	Vemana Mangali	B.Sc. (MPCS)	EM	M. Vemana
36	Venkateswara Rao Dumpala	B.Sc. (MPCS)	EM	Venkateswara Rao
37	Vidhyadhar Madapuri	B.Sc. (MPCS)	EM	Vidhyadhar

INDUCTION PROGRAMME -2019-20

SL NO	Name of the Student	Group	Med	SIGNATURE
38	Vijayarathnam Basavamgari	B.Sc. (MPCS)	EM	Vijayarathnam
39	Mahesh Naik Bukke	B.Sc. (MSCS)	EM	Mahesh Naik
40	Mallikarjuna Ganta	B.Sc. (MSCS)	EM	Mahesh Naik
41	Manoj Kumar Sreeram	B.Sc. (MSCS)	EM	Manojkumar
42	Mohan Vatambeti	B.Sc. (MSCS)	EM	Mohan
43	Mohan Krishna Galam	B.Sc. (MSCS)	EM	G. Mohan Krishna
44	Sireesha Narukuru	B.Sc. (MSCS)	EM	N. Sireesha
45	Siva Jyothi Bodicherla	B.Sc. (MSCS)	EM	Siva Jyothi B.
46	Siva Shankar Reddy Gundre	B.Sc. (MSCS)	EM	S. Siva Shankar Reddy
47	Somasekhar Kukati	B.Sc. (MSCS)	EM	Somasekhar
48	Srinivasulu Punamalli	B.Sc. (MSCS)	EM	S. Srinivasulu
49	Adi Lakshmi Bukya	B.Sc. (CBZ)	TM	Adi Lakshmi
50	Balachandra Devula	B.Sc. (CBZ)	TM	D. Balachandra
51	Chandra Sekhar Chowdam	B.Sc. (CBZ)	TM	Balachandra
52	Deepa Marimani	B.Sc. (CBZ)	TM	Deepa M
53	Dinesh Naik Bukke	B.Sc. (CBZ)	TM	Dinesh Naik Bukke
54	Sandeep Kumar Pachigalla	B.Sc. (CBZ)	TM	P. Sandeep Kumar
55	Santhosh Kumar Pasala	B.Sc. (CBZ)	TM	D. Santhosh Kumar
56	Sasi Kumar K	B.Sc. (CBZ)	TM	K. Sasi Kumar
57	Bhavani Jeepalem	B.Sc. (BBC)	EM	Bhavani
58	Chandini Adimulam	B.Sc. (BBC)	EM	Chandini
59	Devendra Mala	B.Sc. (BBC)	EM	M. Devendra
60	Devi Chitta	B.Sc. (BBC)	EM	C. Devi
61	Goutham Krishna Bandila	B.Sc. (BBC)	EM	Goutham Krishna
62	Mahesh Reddy Thippathigari	B.Sc. (BBC)	EM	Mahesh Reddy T
63	Mallikarjuna Karena	B.Sc. (BBC)	EM	K. Mallikarjuna
64	Manjunath Naik Bhukya	B.Sc. (BBC)	EM	Manjunath Naik
65	Poojitha Anthati	B.Sc. (BBC)	EM	A. Poojitha
66	Poomima Nelavoy	B.Sc. (BBC)	EM	N. Poomima
67	Pradeep Kumar Pemara	B.Sc. (BBC)	EM	Pradeep Kumar
68	Prakash Naik Mude	B.Sc. (BBC)	EM	Prakash Naik
69	Siva Kumar Muga	B.Sc. (BBC)	EM	Sivakumar
70	Sowmya Gangalapudi	B.Sc. (BBC)	EM	G. Sowmya
71	Sreelatha Thamarla	B.Sc. (BBC)	EM	T. Sreelatha
72	Brahma Reddy Marupolu	B.Sc. (MZO)	EM	Brahma Reddy
73	Chandu Chinthagunta	B.Sc. (MZO)	EM	C. Chandu
74	Charan Teja Bathalapalli	B.Sc. (MZO)	EM	B. Charan Teja

INDUCTION PROGRAMME				
SL NO	Name of the Student	Group	Med	SIGNATURE
75	Dhamodar Naik Bhukya	B.Sc. (MZO)	EM	Dhamodar Naik
76	Marudhapandyan.M	B.Sc. (MZO)	EM	M. Marudhapandyan
77	Mary Savarambaka	B.Sc. (MZO)	EM	M. Savarambaka
78	Mayuri Nallapareddy	B.Sc. (MZO)	EM	Mayuri
79	Moneesh Kumar Yadav Pamula	B.Sc. (MZO)	EM	Moneesh Kumar
80	Nadiya Bogyari	B.Sc. (MZO)	EM	Nadiya
81	Narendra Reddy Sura	B.Sc. (MZO)	EM	S. Narendra Reddy
82	Naveen Harijana	B.Sc. (MZO)	EM	H. Naveen
83	Nikhitha Revuri	B.Sc. (MZO)	EM	R. Nikhitha
84	Ramakrishna Chakali	B.Sc. (MZO)	EM	Ramakrishna
85	Ramesh Akkagari	B.Sc. (MZO)	EM	Ramesh
86	Reddi Prakash Petaraju	B.Sc. (MZO)	EM	P. Prakash Petaraju
87	Reddisekhar Naidu Gurramkonda	B.Sc. (MZO)	EM	Reddisekhar
88	Reena Shaik	B.Sc. (MZO)	EM	S. Reena
89	Sai Chaithanya Vallivedu	B.Sc. (MZO)	EM	V. Sai Chaithanya
90	Bhargav Sriakula	B.Sc. (CT&HM)	EM	S. Bhargav
91	Dayakar Pasala	B.Sc. (CT&HM)	EM	P. Dayakar
92	Devendranath Parsha	B.Sc. (CT&HM)	EM	Devendranath
93	Dileep Kumar Pamula	B.Sc. (CT&HM)	EM	Dileep Kumar
94	Dileep Kumar Sanambatia	B.Sc. (CT&HM)	EM	S. Dileep Kumar
95	Durga Nagendra Veeramallu	B.Sc. (CT&HM)	EM	Nagendra
96	Gayathri Mypati	B.Sc. (CT&HM)	EM	Gayathri Mypati
97	Madhuri Kamale	B.Sc. (CT&HM)	EM	Madhuri
98	Mahesh Babu Badakuri	B.Sc. (CT&HM)	EM	Mahesh Babu
99	Manikanta Patnam	B.Sc. (CT&HM)	EM	P. Manikanta
100	Manoj Vamsi Krishna kotapati	B.Sc. (CT&HM)	EM	Manoj Vamsi
101	Sreenu Soka	B.Sc. (CT&HM)	EM	S. Sreenu
102	Sri Hari Lokesh Velagacherla	B.Sc. (CT&HM)	EM	S. Sri Hari Lokesh
103	Srinivasulu Yalakacharla	B.Sc. (CT&HM)	EM	Srinivasulu
104	Sudhakar Muthamsetty	B.Sc. (CT&HM)	EM	Sudhakar
105	Vamsi Gunipati	B.Sc. (CT&HM)	EM	G. Vamsi
106	Vineeth Modiboyina	B.Sc. (CT&HM)	EM	M. Vineeth
107	Vishal Yadav.Y	B.Sc. (CT&HM)	EM	Vishal Yadav
108	Yaswanth Thalari	B.Sc. (CT&HM)	EM	Yaswanth





Dr.C.Ramesh, Principal delivering key- note address during the orientation programme conducted for first year on 09-08-2019.

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

Student supporting programs / Events conducted

Name of the department: Statistics

Academic year: 2019-20


S. No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	ORIENTATION PROGRAMME	06/07/2019	B.Sc:MSCs MPS BA: MES,ASCA	43	-

S.G.S. ARTS COLLEGE, TTD, TIRUPATI

Department of STATISTICS

EVENT ORGANISED REPORT: 2019-20

1	Name of the Department	Statistics
2	Name of Event Organized	Orientation programme for I year students
3	Title of Event Organized	Awareness program on Statistics in real life.
4	Name of Co-ordinator of Event	Dr.D.Chandra kesavulu Naidu
5	Resource Person	Dr P.Maheswari, Assistant Professor in Statistics , SVA Degree College for Women, Srikalahasthi.
6	Class of the Participants	B.Sc:MSCs MPS BA: MES,ASCA
7	No. of Participants	43
8	Objective of the Event	➤ To create awareness on Statistics in real life. Career opportunities and higher education.
9	Outcome of the Event	➤ Appreciated use of Statistics in real life ➤ Got to know Career Opportunities for Statistics graduates and Higher education Opportunities.


D. CHANDRASEKHAR NAIDU,
M.Sc., M.Phil., Ph.D.
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
TIRUPATI

Orientation Programme

06/07/2019

The department of Statistics has conducted Orientation Programme for the 1 year students of BA, B.Sc (- MSc), MPS, MES & ASCA).

Objective: To bring awareness among the students about Statistics in real life applications, career opportunities and higher studies.

Resource Person: Dr. P. Maheswari,

Asst. Professor in Statistics

S.V.A. Govt. Degree College for women,
Sri Kalahasti.

Keywords: 1. Statistics - in real life
2. Career opportunities after UG program
3. Institution offering Ph program in Statistics

The Head of dept Dr. D. Chandra Kesavulu Naidu and faculty members addressed students to enlighten about rules & regulation, discipline to be followed in the college for better future.

Faculty: 1) Dr. K. Visalakshmi - Head

D. CHANDRA KESAVULU NAIDU,
M.Sc., M.Phil., Ph.D.
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
(TIRUPATI)

2) Dr. K. Venkatesh

3) Dr. M. Vasudevarao

M. V.G.

① P. Chakri

P. Chakri

② K. Bhaskar

K. Bhaskar

③ U. Charan

U. Charan

④ K. Deepa

K. Deepa

⑤ B. Anandh Raj

B. Anandh Raj

⑥ M. Ajay Naidu

M. Ajay Naidu

⑦ M. Suresh

M. Suresh

⑧ K. Venkata Chankya

K. Venkata Chankya

⑨ S. Vinay

S. Vinay

(10) C. Vijay Goud	(44) M. Meghana
(11) P. Haritha	P. Haritha
(12) R. Hemanth	R. Hemanth
(13) K. Jyothi	Jyothi K
(14) M. Lakshmi	M. Lakshmi
(15) N. Manoj	N. Manoj
(16) G. Ramesh	G. Ramesh
(17) K. Sireesha	K. Sireesha
(18) C. Sridevi	C. Sridevi
(19) B. Vijay Vishva	B. Vijay
(20) B. Viswanath	B. Viswanath
(21) J. Chandrudu	J. Chandrudu
(22) S. Gurunath	S. Gurunath
(23) B. Vasu	B. Vasu
(24) P. Rakesh	P. Rakesh
(25) G. Surya	P. Rakesh
(26) M. Hemanth	M. Hemanth
(27) P. Lokesh	P. Lokesh
(28) D. Raja	D. Raja
(29) S. Jagadeeswari	S. Jagadeeswari
(30) T. Vinay	T. Vinay
(31) R. Bhuvanesh	R. Bhuvanesh
(32) D. Supriya	D. Supriya
(33) M. Anusha	M. Anusha
(34) T. Sunil	T. Sunil
(35) C. Lokesh	C. Lokesh
(36) G. Suresh	G. Suresh
(37) B. Sumalatha	B. Sumalatha
(38) M. Anusha	M. Anusha
(39) K. Rupesh	K. Rupesh
(40) D. Allipeera	D. Allipeera
(41) C. Sandeep	C. Sandeep
(42) U. Sai Kumar	U. Sai Kumar
(43) U. Lokanath	U. Lokanath

Orientation Programme on to bring awareness among the students about statistics in real life applications, career opportunities and higher studies by Dr P/Lakshmi.



All three year statistics students participated in Orientation Programme.



2. Bridge Course

Abstract of the Bridge Course conducted is shown here complete details are available in college Website

Phone: (0877) 2264599

E-mail: sgsartscollegettd@gmail.com

Web: sgsac.edu.in



TIRUMALA TIRUPATI DEVASTHANAMS S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Academic year 2016-17

Summary of Bridge courses

The Institution has conducted Bridge courses with an objective to Bridge the gap between subjects studied at their higher secondary level and the subjects they would be studying in their Graduation level. The syllabus for the course is framed in such way that they get basic knowledge and the subjects they would be learning through Graduation.

S.NO	Department	Date of conduction	Total Number of hours taught	Students to whom Bridge course was given	Number of students benefited
1	Chemistry	18-07-2016 to 23-07-2016	30	Students who have no chemistry back ground in their 10+2 level(Vocational study)	3
2	Mathematics	18-07-2016 to 23-07-2016	30	Students with biological background in their 10+2 level	96
3	Statistics	11-07-20-16 to 16-07-2016	30	Students with biological background in their 10+2 level	96




PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

Department of Statistics

Tirupathi
06.03.2021

From
Sri. D. Chandrakesavulu Naidu,
HOD of Statistics,
SGS Arts College,
Tirupathi.

To
The Principal,
SGS Arts College,
Tirupathi.


Sir,

Sub:- Permission for conducting Statistics Bridge course to students of
first year Bi.P.C Degree – request – Regarding.

&&&&

Kindly permit the Department of Statistics to conduct Bridge
Course to the students of first year, who have not studied statistics in
intermediate level, for one week i.e. from 15-03-2021 to 16- 03-2021. Time
table and syllabus are enclosed.

Thanking you sir


Yours faithfully
D. CHANDRA KESAVULU NAIDU,
M.Sc., M.Phil., Ph.D
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
TIRUPATI

2020-2021
SGS Arts College, Tirupati
Department of Statistics

Circular

First-year degree students who studied Bi.P.C Course at intermediate level are informed to attend the Statistics bridge course which will be held at Room No. 222, Department of Statistics, Tirupati from 15-03-2021 to 16-03-2021. All Biology Student should attend the with out fail


D. CHANDRA KESAVULU NAIDU,
Head, M.Phil., Ph.D
LECTURER & HEAD
Department of Statistics
S.G.S. ARTS COLLEGE
TIRUPATI

2020-2021
Bridge Course-Summary

Name of the programme	Bridge course in Statistics
Duration of Course	One Week
Date	15-03-2021 to 16-03-2021
Venue	Room No. 222

S.No	Subject name	Faculty Name	No. of Students
1	Meaning Scope and Limitations of Statistics	Dr. K Vijaya Kumar	116
2	Measures of Central Tendency	Dr. K Venkatesh	116
3	Measures of Dispersion	Dr. M Vasudeva Reddy	116


D. CHANDRA KESAVULU NAIDU,
 M.Sc., M.Phil., Ph.D
 LECTURER & HEAD
 DEPT. OF STATISTICS
 S.G.S. ARTS COLLEGE
 TIRUPATI

S.G.S. ARTS COLLEGE
TIRUPATI


2020-2021
Department of Statistics
Bridge Course Timetable

Day	10AM-11:30 AM	11:30AM-1PM	2PM-3PM	3PM-4PM
Wednesday	Meaning Scope and Limitations of Statistics	Measures of Dispersion	Measures of Central Tendency	Q&A Session
Thursday	Measures of Central Tendency	Meaning Scope and Limitations of Statistics	Measures of Dispersion	
Friday	Measures of Dispersion	Measures of Central Tendency	Meaning Scope and Limitations of Statistics	
Saturday	Meaning Scope and Limitations of Statistics	Measures of Dispersion	Measures of Central Tendency	
Monday	Measures of Central Tendency	Meaning Scope and Limitations of Statistics	Measures of Dispersion	
Tuesday	Measures of Dispersion	Measures of Central Tendency	Meaning Scope and Limitations of Statistics	


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2020-2021
Bridge Course in Statistics
Syllabus

S.No	Subject	Syllabus
1	Meaning Scope and Limitations of Statistics	Primary and Secondary, Classification and Tabulation, Construction of frequency distribution. Graphical Representation : Histogram, Bar, Pie and Frequency polygon.
2	Measures of Central Tendency	Features of good average, Arithmetic Mean, Median, Mode. Empirical relationship between Mean, central Median and Mode and skewness based on values.
3	Measures of Dispersion	Range, Quartile Deviation (QD), Mean Deviation (MD), Variance, Standard Deviation (SD), relationship between QD, MD and SD. Familiarization of the concepts relating to Correlation and Linear Regression


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DEPARTMENT OF STATISTICS
BBC :: 2020-21 Bridge Course : Basic Statistics

S.NO	Roll number	Name of the student	Marks
1	20919	Chakradhar Goud E	22
2	20920	Chan Basha P	19
3	20921	Dheeraj M	21
4	20922	Divya Bharathi Reddy G	19
5	20923	Ganesh B	23
6	20924	Harshavardhan K.G	25
7	20925	Hemalatha M	20 A
8	20926	Hemanth K	23
9	20927	Indira Mythi D	17
10	20928	Janaki Raman RS	20
11	20929	Kalyan M	22
12	20930	Kiran N	21
13	20931	Krishna D	19
14	20932	Kushwanth T	23
15	20933	Lakshmana T	19
16	20934	Leena P	25
17	20935	Mamatha A	25
18	20936	Manjusha K	24
19	20937	Mohana A	23
20	20938	Naga Roshini N	19
21	20939	Nataraj D	17
22	20940	Naveen Kumar K	21
23	20941	Pavan Kumar Raju O	22
24	20942	Pavan Kumar K	19
25	20943	Phanendra Kumar B	23
26	20944	Praveen S	24
27	20945	Ravi Kumar T	19
28	20946	Sai Praathap K	21
29	20947	Sai Sree Harsha P V	17
30	20948	Sai Sreenivas K	19
31	20949	Subba Reddy B	22
32	20950	Surendra Babu S	24
33	20951	Tejasree R	25
34	20952	Usha Sree P	25
35	20953	Vara Prasad M	19
36	20954	Vidhyavathi B	20
37	20955	Vignesh Yadav T	17
38	20956	Vijaya Brahma C	18
39	20957	Yamini B	23

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SGS ARTS COLLEGE:TTD:TPT
DEPARTMENT OF STATISTICS
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S.NO	Roll number	Name of the student	15.03.2021	16.03.2021
1	20919	Chakradhar Goud.E	!	!
2	20920	Chan Basha.P	!	!
3	20921	Dheeraj.M	!	!
4	20922	Divya Bharathi Reddy.G	!	!
5	20923	Ganesh.B	!	!
6	20924	Harshavardhan.K.G	!	!
7	20925	Hemalatha.M	!	!
8	20926	Hemanth.K	!	!
9	20927	Indira Mythri.D	!	!
10	20928	Janaki Raman.RS	!	!
11	20929	Kalyan.M	!	!
12	20930	Kiran.N	!	!
13	20931	Krishna.D	!	!
14	20932	Kushwanth.T	!	!
15	20933	Lakshmana.T	!	!
16	20934	Leena.P	!	!
17	20935	Mamatha.A	!	!
18	20936	Manjusha.K	!	!
19	20937	Mohana.A	!	!
20	20938	Naga Roshini.N	!	!
21	20939	Nataraj.D	!	!
22	20940	Naveen Kumar.K	!	!
23	20941	Pavan Kumar Raju.O	!	!
24	20942	Pavan Kumar.K	!	!
25	20943	Phanendra Kumar.B	!	!
26	20944	Praveen.S	!	!
27	20945	Ravi Kumar.T	!	!
28	20946	Sai Praathap.K	!	!
29	20947	Sai Sree Harsha.P.V	!	!
30	20948	Sai Sreenivas.K	!	!
31	20949	Subba Reddy.B	!	!
32	20950	Surendra Babu.S	!	!
33	20951	Tejasree.R	!	!
34	20952	Usha Sree.P	!	!
35	20953	Vara Prasad.M	!	!
36	20954	Vidhyavathi.B	!	!
37	20955	Vignesh Yadav.T	!	!
38	20956	Vijaya Brahma.C	!	!
39	20957	Yamini.B	!	!


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S.NO	Roll number	Name of the student	Signature
1	20919	Chakradhar Goud E	E. Chakradhar Goud
2	20920	Chan Basha P	Chan Basha
3	20921	Dheeraj M	M. Dheeraj
4	20922	Divya Bharathi Reddy G	Divya Bharathi Reddy
5	20923	Ganesh B	Ganesh B
6	20924	Harshavardhan K G	K. Harshavardhan
7	20925	Heemalatha M- T C	M. Heemalatha
8	20926	Hemanth K	K. Hemanth
9	20927	Indira Mythri D	D. Indira Mythri
10	20928	Janaki Raman RS	R. S. Janaki
11	20929	Kalyan M	M. Kalyan
12	20930	Kiran N	N. Kiran
13	20931	Krishna D	D. Krishna
14	20932	Kushwanth T	T. Kushwanth
15	20933	Lakshmana T	T. Lakshmana
16	20934	Leena P	P. Leena
17	20935	Mamatha A	A. Mamatha
18	20936	Manjusha K	K. Manjusha
19	20937	Mohana A	A. Mohana
20	20938	Naga Roshini N	N. Naga Roshini
21	20939	Nataraj D	D. Nataraj
22	20940	Naveen Kumar K	K. Naveen Kumar
23	20941	Pavan Kumar Raju O	O. P. K. Raju
24	20942	Pavan Kumar K	K. Pavan Kumar
25	20943	Phanendra Kumar B	B. Phanendra Kumar
26	20944	Praveen S	S. Praveen
27	20945	Ravi Kumar T	T. Ravi Kumar
28	20946	Sai Praathap K	K. Sai Praathap
29	20947	Sai Sree Harsha P V	P. V. Sai Sree Harsha
30	20948	Sai Sreenivas K	K. Sai Sreenivas
31	20949	Subba Reddy B	B. Subba Reddy
32	20950	Surendra Babu S	S. Surendra Babu
33	20951	Tejasree R	R. Tejasree
34	20952	Usha Sree P	P. Usha Sree
35	20953	Vara Prasad M	M. Vara Prasad
36	20954	Vidhyavathi B	B. Vidhyavathi
37	20955	Vignesh Yadav T	T. Vignesh Yadav
38	20956	Vijaya Brahma C	C. Vijaya Brahma
39	20957	Yamini B	B. Yamini

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S.NO	Roll number	Name of the student	Marks
1	20969	ARAVIND K	23
2	20970	ASHOK DUPPANAGURTHI	24
3	20971	AVINASH SEDAM	19
4	20972	AYESHA SHAIK	21
5	20973	BALAKULLEYAPPA AAKULEDU	22
6	20974	CHAITANYA BALINENI	19
7	20975	DEEPAK ROYAL BANDLA	17
8	20976	Dora Babu. Kaputhota	16
9	20977	DHANUSH CHINTHALA	22
10	20978	DHEERAJ KUMAR SODEM	24
11	20979	GOWTHAM NAIK MUDE	25
12	20980	GURU RAJU KURUVA	24
13	20981	HANISH BABU BODANAPU	25
14	20982	HARSHITA DAKSHINADI	25
15	20983	HEMANTH KARNAM	17
16	20984	JEEVAN KUMAR KEELAPATLA	19
17	20985	KISHORE NAIK MUDE	18
18	20986	LOKANADHAM GUTHI	20
19	20987	KUSHWANTH KUMAR M	21
20	20988	LAKSHMIKANTH NAIK B	24
21	20989	LOKESH ANKE	22
22	20990	LOKESH PUTTURU	19
23	20991	MADHU BANDILA	17
24	20992	MANESH KUMAR JAGGALI	22
25	20993	MANOJ KUMAR REDDY V	19
26	20994	PAVAN KUMAR JANGAM	19
27	20995	PRBHAKARA CHINTAKAYALA	20
28	20996	RAMAKRISHNA NAIK BUKKE	21
29	20997	RISHITHA POKALA	22
30	20998	SAI PRAKASH VEMULA	24
31	20999	SANDHYA YERRABALLI	21
32	21000	SATHEESH GOLLAPALLI	23
33	21001	Seetharamanjaneyulu K	19
34	21002	SHARATH KUMAR Y	17
35	21003	SHIK TASNEEM MALLELA	16
36	21004	SIVANI BOMMI	19
37	21005	THULASI NALLIAHGARI	21
38	21006	VENKATESH AADHULA	24
39	21007	VENKATESH T	25


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S.NO	Roll number	Name of the student	15.03.2021	16.03.2021
1	20969	ARAVIND .K	1	1
2	20970	ASHOK DUPPANAGURTHI	1	1
3	20971	AVINASH SEDAM	1	1
4	20972	AYESHA SHAIK	1	1
5	20973	BALAKULLEYAPPA AAKULEDU	1	1
6	20974	CHAITANYA BALINENI	1	1
7	20975	DEEPAK ROYAL BANDLA	1	1
8	20976	Dora Babu. Kapulhota	A	1
9	20977	DHANUSH CHINTHALA	1	1
10	20978	DHEERAJ KUMAR SODEM	1	1
11	20979	GOWTHAM NAIK MUDE	1	1
12	20980	GURU RAJU KURUVA	1	1
13	20981	HANISH BABU BODANAPU	1	1
14	20982	HARSHITA DAKSHINADI	1	1
15	20983	HEMANTH KARNAM	1	1
16	20984	JEEVAN KUMAR KEELAPATLA	1	1
17	20985	KISHORE NAIK MUDE	1	1
18	20986	LOKANADHAM GUTHI	1	1
19	20987	KUSHVANTH KUMAR.M	1	1
20	20988	LAKSHMIKANTH NAIK B	1	1
21	20989	LOKESH ANKE	1	1
22	20990	LOKESH PUTTURU	1	1
23	20991	MADHU BANDILA	1	1
24	20992	MANESH KUMAR JAGGALI	1	1
25	20993	MANOJ KUMAR REDDY .V	1	1
26	20994	PAVAN KUMAR JANGAM	A	1
27	20995	PRBHAKARA CHINTAKAYALA	1	1
28	20996	RAMAKRISHNA NAIK BUKKE	1	1
29	20997	RISHITHA POKALA	1	1
30	20998	SAI PRAKASH VEMULA	1	1
31	20999	SANDHYA YERRABALLI	1	1
32	21000	SATHEESH GOLLAPALLI	1	A
33	21001	Seetharamanjanyulu.K	1	1
34	21002	SHARATH KUMAR.Y	1	1
35	21003	SHIK TASNEEM MALLELA	1	1
36	21004	SIVANI BOMMI	1	1
37	21005	THULASI NALLIAHGARI	1	1
38	21006	VENKATESH AADHULA	1	1
39	21007	VENKATESH.T	1	1

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S.NO	Roll number	Name of the student	Signature
1	20969	ARAVIND K	K. Aravind
2	20970	ASHOK DUPPANAGURTHI	D. Ashok
3	20971	AVINASH SEDAM	S. Avinash
4	20972	AYESHA SHAIK	af. Ayesha
5	20973	BALAKULLEYAPPA AAKULEDU	A. Balakulleyappa
6	20974	CHAITANYA BALINENI	B. Chaitanya
7	20975	DEEPAK ROYAL BANDLA	B. Deepak
8	20976	Dora Babu. Kaputhota	K. Dora Babu
9	20977	DHANUSH CHINTHALA	C. Dhaniush
10	20978	DHEERAJ KUMAR SODEM	A. Dheeraj Kumar
11	20979	GOWTHAM NAIK MUDE	M. Gowtham Naik
12	20980	GURU RAJU KURUVA	K. Gururaju
13	20981	HANISH BABU BODANAPU	B. Hanish Babu
14	20982	HARSHITA DAKSHINADI	D. Harshitha
15	20983	HEMANTH KARNAM	K. Hemanth
16	20984	JEEVAN KUMAR KEELAPATLA	K. Jeevan Kumar
17	20985	KISHORE NAIK MUDE	M. Kishore Naik
18	20986	LOKANADHAM GUTHI	G. Lokanadham
19	20987	KUSHWANTH KUMAR M	M. Kushwanth Kumar
20	20988	LAKSHMIKANTH NAIK B	B. Lakshmi Kanth
21	20989	LOKESH ANKE	A. Lokesh
22	20990	LOKESH PUTTURU	P. Lokesh
23	20991	MADHU BANDILA	B. Madhu
24	20992	MANESH KUMAR JAGGALI	J. Manesh Kumar
25	20993	MANOJ KUMAR REDDY V	V. Manoj Kumar
26	20994	PAVAN KUMAR JANGAM	A. Pavan Kumar
27	20995	PRBHAKARA CHINTAKAYALA	C. Prabhakar
28	20996	RAMAKRISHNA NAIK BUKKE	N. Ramakrishna Naik
29	20997	RISHITHA POKALA	P. Rishitha
30	20998	SAI PRAKASH VEMULA	V. Sai Prakash
31	20999	SANDHYA YERRABALLI	S. Sandhya
32	21000	SATHEESH GOLLAPALLI	G. Satheesh
33	21001	Seetharamanjanyulu K	K. Seetharamanyulu
34	21002	SHARATH KUMAR Y	Y. Sharath Kumar
35	21003	SHIK TASNEEM MALLELA	M. Tasneem
36	21004	SIVANI BOMMI	B. Sivani
37	21005	THULASI NALLIAHGARI	N. Tulasi
38	21006	VENKATESH AADHULA	A. Venkatesh
39	21007	VENKATESH T	T. Venkatesh

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SGS ARTS COLLEGE-TTD-TPT
DEPARTMENT OF STATISTICS
CBZ :: 2020-21 Bridge Course : Basic Statistics

S.NO	Roll number	Name of the student	Signature
1	20869	ANUPRIYA KUPPALA	K. Anupriya
2	20870	BHANU KIRAN YADALA	Y. Bhanukirana
3	20871	BHANU PRASAD NALAVALI	N. Bhanu Prasad
4	20872	BHARATH KUMAR .K	K. Bharath Kumar
5	20873	BHARGAVA DAKKA	D. Bhargava
6	20874	BHUVANESH CHERUKURI	C. Bhuvanesh
7	20875	CHAITANYA MADDURI	M. Chaitanya
8	20876	CHANDRA VAMSI AREKANTI	A. Chandra Vamsi
9	20877	CHANDU KOPPALA	K. Chandu
10	20878	GANESH ERRABALLI	E. Ganesh
11	20879	GOWTHAMI KATINENI	K. Gowthami
12	20880	GUNATEJA RAYAPATI	R. Gunateja
13	20881	HARI OM MADURI	M. Hari Om
14	20882	JANARDHAN KOTHAKOTA	K. Janardhan
15	20883	JAVEED SHAIK	J. Javeed Shaik
16	20884	JHANSI BAI KELUTH	K. Jhansi Bai
17	20885	KAMAL KATUPAKAM	K. Kamal
18	20886	KIRAN KUMAR VADIGILA	V. Kiran Kumar
19	20887	MAHESH ABBIGANDLA	A. Mahesh
20	20888	MEGHANA ALWAR	M. Meghana
21	20889	MOHAN CHEEMALA	C. MOHAN
22	20890	NARASIMHA VRADAN DAMAA	N. Narasimha
23	20891	NAVEEN KUMAR MEKALA	M. Naveen
24	20892	PAVAN KALYAN BURRI	B. Pavan Kalyan
25	20893	PRAVEEN LOKASANI	L. Praveen
26	20894	PRAVEEN VALLURU	V. Praveen Kumar
27	20895	PREMSAGAR NAKKA	N. Premagar
28	20896	RADHIKA JANAM	R. Radhika Janam
29	20897	RAGHU BHEEMANNAGARI	R. Raghu Bheemanna
30	20898	Sai Yaswanth Kumar.C	C. Sai Yaswanth
31	20899	SARAT CHANDRA .S	S. Sarath Chandra
32	20900	SIVA KUMAR BOLLA	B. Siva Kumar



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33	20901	SIVANATH REDDY MUDAMALA	Siva nambreddy
34	20902	SWATHI GUNJI	G. Swathi
35	20903	UMESH CHANDRA MANGALI	U. Umesh chandra
36	20904	VAMSI BANDILA	Vamsi
37	20905	VINEELA PARICHERLA	P. Vineela
38	20906	VINOD KUMAR MALLEPOGU	H. Vinod Kumar


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2	20870	BHANU KIRAN YADALA	1	1
3	20871	BHANU PRASAD NALAVALI	1	1
4	20872	BHARATH KUMAR .K	1	1
5	20873	BHARGAVA DAKKA	1	1
6	20874	BHUVANESH CHERUKURI	1	1
7	20875	CHAITANYA MADDURI	1	1
8	20876	CHANDRA VAMSI AREKANTI	1	1
9	20877	CHANDU KOPPALA	1	1
10	20878	GANESH ERRABALLI	1	1
11	20879	GOWTHAMI KATINENI	1	1
12	20880	GUNATEJA RAYAPATI	1	1
13	20881	HARI OM MADURI	1	1
14	20882	JANARDHAN KOTHAKOTA	1	1
15	20883	JAVEED SHAIK	1	1
16	20884	JHANSI BAI KELUTH	1	1
17	20885	KAMAL KATUPAKAM	1	1
18	20886	KIRAN KUMAR VADIGILA	1	1
19	20887	MAHESH ABBIGANDLA	1	A
20	20888	MEGHANA ALWAR	1	1
21	20889	MOHAN CHEEMALA	1	1
22	20890	NARASIMHA VRADAN DAMAA	1	1
23	20891	NAVEEN KUMAR MEKALA	1	1
24	20892	PAVAN KALYAN BURRI	1	1
25	20893	PRABEEN LOKASANI	1	1
26	20894	PRABEEN VALLURU	1	1
27	20895	PREMSAGAR NAKKA	1	1
28	20896	RADHIKA JANAM	1	1
29	20897	RAGHU BHEEMANNAGARI	1	1
30	20898	Sai Yaswanth Kumar.C	1	1
31	20899	SARAT CHANDRA .S	1	1
32	20900	SIVA KUMAR BOLLA	1	1
33	20901	SIVANATH REDDY MUDAMALA	1	1
34	20902	SWATHI GUNJI	1	1
35	20903	UMESH CHANDRA MANGALI	1	1
36	20904	VAMSI BANDILA	1	1
37	20905	VINEELA PARICHERLA	1	1
38	20906	VINOD KUMAR MALLEPOGU	1	1


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2	20870	BHANU KIRAN YADALA	22
3	20871	BHANU PRASAD NALAVALI	16
4	20872	BHARATH KUMAR K	19
5	20873	BHARGAVA DAKKA	20
6	20874	BHUVANESH CHERUKURI	12
7	20875	CHAITANYA MADDURI	23
8	20876	CHANDRA VAMSI AREKANTI	24
9	20877	CHANDU KOPPALA	22
10	20878	GANESH ERRABALLI	20
11	20879	GOWTHAMI KATINENI	19
12	20880	GUNATEJA RAYAPATI	19
13	20881	HARI OM MADURI	12
14	20882	JANARDHAN KOTHAKOTA	20
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16	20884	JHANSI BAI KELUTH	23
17	20885	KAMAL KATUPAKAM	12
18	20886	KIRAN KUMAR VADIGILA	20
19	20887	MAHESH ABBIGANDLA	23
20	20888	MEGHANA ALWAR	24
21	20889	MOHAN CHEEMALA	25
22	20890	NARASIMHA VRADAN DAMAA	25
23	20891	NAVEEN KUMAR MEKALA	20
24	20892	PAVAN KALYAN BURRI	19
25	20893	PRAVEEN LOKASANI	12
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36	20904	VAMSI BANDILA	12
37	20905	VINEELA PARICHERLA	20
38	20906	VINOD KUMAR MALLEPOGU	19

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- SGS ARTS COLLEGE

MODEL EXAM PAPER - 2020-21

BRIDGE COURSE - BIOLOGY STUDENTS

21
25

NAME:

Group:

Total marks: 25

Time: 1hr

SECTION - A

All questions carry equal marks

Fill in the blanks

(10 x 1 = 10)

- ① Data obtained by conducting a survey is called Primary data ✓
- ② The data collected from published reports is known as Secondary data ✓
- ③ statistics deals with only Quantitative data ✓
- ④ The difference between the upper and lower limit of a class is called class interval ✓
- ⑤ an arrangement of data in rows and columns is known as tabulation ✓
- ⑥ measures of central tendency are also known as typical value or _____.
- ⑦ median is same as fiftieth percentile. ✓
- ⑧ Best measure of dispersion is Mean deviation ✓
- ⑨ The term regression was introduced by Sir Francis Galton ✓
- ⑩ The significance of regression parameters can be tested by S-test ✓

SECTION-B

Multiple choice questions

(15 x 1 = 15)

- ⑪ The word 'statistics' is used as: [c] ✓
a) singular b) plural
c) singular and plural both d) none of the above
- ⑫ Which of the following represents data? [c] ✓
a) a single value b) Only two values in a set
c) a group of values in a set d) none of the above
- ⑬ In a grouped data, the number of classes preferred are [a] ✓
a) minimum possible b) adequate
c) maximum possible d) any arbitrarily chosen number
- ⑭ Trilinear chart is used to portray simultaneously: [b] ✓
a) two variables b) three variables
c) four variables d) any number of variables
- ⑮ Non-dimensional diagrams are also known as: [a] ✗
a) cubes b) spheres
c) pictograms d) All the above
- ⑯ Ogive curve occur for, [c] ✓
a) more than type distribution
b) less than type distribution
c) both (a) and (b)
d) none of (a) and (b)
- ⑰ Which of the following is not a measure of dispersion? [d] ✓
a) mean deviation b) quartile deviation
c) standard deviation d) average deviation from mean.

which one of the given measures of dispersion is considered best?

- a) standard deviation
- b) quartile deviation
- c) range
- d) variation

[a]

19) which of the following is a unitless measure of dispersion [c]

- a) standard deviation
- b) mean deviation
- c) coefficient of variation
- d) range

20) The term regression was introduced by:

- a) R.A. Fisher
- b) Sir Francis Galton
- c) Karl Pearson
- d) none of the above

[b]

21) Regression equation is also named as:

- a) prediction equation
- b) estimating equation
- c) line of average relationship
- d) all the above

[d]

22) In the regression line $Y = a + bX$, b is called the

- a) slope of the line
- b) intercept of the line
- c) neither (a) nor (b)
- d) both (a) and (b)

[a]

23) If $\beta_{YX} > 1$, then β_{XY} is

- a) less than 1
- b) greater than 1
- c) equal to 1
- d) equal to 0

[a]

24) The shape of trilinear charts is that of a:

- a) cone
- b) cube
- c) equilateral triangle
- d) pyramid

[c]

25) Ogives for more than type and less than type distributions intersect at:

- a) mean
- b) median
- c) mode
- d) origin

[b]

Key - 2020-2021 model exam paper.

SECTION - A

- ① primary, ② secondary, ③ quantitative data,
④ class interval, ⑤ tabulation, ⑥ typical value, ⑦ Fiftieth,
⑧ standard deviation, ⑨ Sir Francis Galton,
⑩ t-test

SECTION - B

- ⑪ c, ⑫ c, ⑬ b, ⑭ b, ⑮ c, ⑯ c, ⑰ d, ⑱ a, ⑲ c,
⑳ b, ㉑ d, ㉒ a, ㉓ a, ㉔ c, ㉕ b

Phone: (0877) 2264599

E-mail: sgsartscollegettd@gmail.com
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

Web: sgsac.edu.in

Prof N. VENUGOPAL REDDY,
M.Sc., M.Phil., Ph.D
PRINCIPAL (FAC)



Tirupathi

Date: 22.07.2019

From
Sri. D. Parameswara,
HOD of Chemistry,
SGS Arts College,
Tirupathi.

To
The Principal,
SGS Arts College,
Tirupathi.

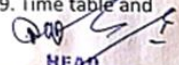
Sir,

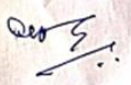
Sub:- Permission for conducting Chemistry Bridge course to students of
first year degree – request – Regarding.

&&&&

Kindly permit the Department of Chemistry to conduct Bridge
Course to the students of first year degree who studied Vocational Course in
intermediate for one week i.e. from 24-07-2019 to 30-07-2019. Time table and
syllabus are enclosed.

Thanking you sir

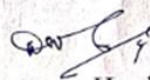

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
Yours faithfully



S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI
Department of Chemistry

Circular

First-year Life Sciences students who studied vocational courses at the intermediate level are informed to attend the bridge course which will be held at seminar hall, SGS Degree College, Tirupati from 24-07-2019 to 30-07-2019.



Head,
Department of Chemistry
HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Bridge Course-Summary

Name of the programme	Bridge course in chemistry for vocational students
Duration of Course	One Week
Date	24-07-2019 to 30-07-2019
Venue	Seminar Hall

S.No	Subject name	Faculty Name	No. of Students
1	Inorganic Chemistry	P Suguna	4
2	Physical Chemistry	V Venkata Lakshmi	4
3	Organic Chemistry	D Parameswara	4
4	General Chemistry	K. Revathi	4


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

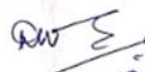
2019-2020
Bridge Course in Chemistry
Students List

S.No	Name of the student	Signature
1	P.Munendra	P. Munendra
2	M.Naveen Naik	M. Naveen Naik
3	N.Lokesh	N. Lokesh
4	Shaik Reena	Shaik Reena
5	S.Mary	S. Mary
6	M.Pandiyam	M. Pandiyam


SIGNATURE
HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Department of Chemistry Bridge Course Time Table

Day	10AM-11AM	11AM-12NOON	12NOON-1PM	2PM-3PM	3PM-4PM
Wednesday	Inorganic	Physical	Organic	General	Q&A Session
Thursday	Physical	Organic	General	Inorganic	
Friday	Organic	General	Inorganic	Physical	
Saturday	General	Inorganic	Physical	Organic	
Monday	Inorganic	Physical	Organic	General	
Tuesday	Physical	Organic	General	Inorganic	


HEAD
 DEPARTMENT OF CHEMISTRY
 S. G. S. ARTS COLLEGE
 TIRUPATHI

Bridge Course in Chemistry Syllabus

S.No	Subject	Syllabus
1	Inorganic Chemistry	Electronic Configuration of elements Classification of elements in to s,p,d,f blocks. Periodic properties
2	Physical Chemistry	States of matter- Gaseous state,solid state Solutions Ionic Equilibria
3	Organic Chemistry	Nomenclature of Organic Compounds Types of Reactions Types of Reagents Bond polarising Effects Reaction Intermediates
4	General Chemistry	Stereo Chemistry Colloids Adsorption Analytical Chemistry


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Department of Chemistry
Bridge Course Student Attendance List

S.No	Name of the student	Course Studied in Intermediate	Programme Joined in UG	24/7/19	25/7/19	26/7/19	27/7/19	29/7/19	30/7/19
1	P.Munendra	Dairy Science	CBZ	X	X	X	a	X	X
2	M.Naveen Na	Dairy Science	CBZ	X	X	X	X	X	X
3	N.Lokesh	Agl.Polytechnic	CBZ	X	a	X	X	X	X
4	Shaik Reena	MLT (Medical Lab Technician Course)	MZC	X	X	X	X	X	X
5	S.Mary	MLT (Medical Lab Technician Course)	MZC	X	X	X	X	X	X
6	M.Pandiyan	MLT (Medical Lab Technician Course)	MZC	X	X	X	X	X	X

1.
Signature of the faculty

2. *V. V. Kumar*

3. *P. S. S. S.*

Dr. S.
HEAD.
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COL.
TIRUPATHI

Bridge Course in Chemistry
Students Marks List

S.No	Name of the student	Marks
1	P.Munendra	13 - Thirteen only
2	M.Naveen Naik	12 - Twelve only
3	N.Lokesh	15 - Fifteen only
4	Shaik Reena	13 - Thirteen only
5	S.Mary	14 - Fourteen only
6	M.Pandiyan	13 - Thirteen only


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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

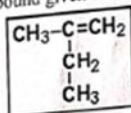
S.G.S ARTS COLLEGE: TIRUCHANOOR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY
OBJECTIVE TEST - BRIDGE COURSE 2019-2020

Answer all questions

Maximum Marks: 20

1. The physical and chemical properties of elements are the periodic () function of their atomic weights is
 - a) Mendeleev's periodic law
 - b) Lothar Meyer's periodic law
 - c) Moseley's periodic law
 - d) Bohr's periodic law
2. Calcium has an atomic number of 20. A stable calcium atom () has an electronic configuration of:
 - a) $1s^2 2s^2 2p^6 3s^2 3p^4 s^2$
 - b) $1s^2 1p^6 1d^{10} 1f^2$
 - c) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2$
 - d) $1s^2 2s^2 2p^6 3s^2 3p^6$
3. The correct order of second ionization potential of C, N, O and F is ()
 - a) $C > N > O > F$
 - b) $O > N > F > C$
 - c) $O > F > N > C$
 - d) $F > O > N > C$
4. Of the different allotropic forms of phosphorus, most reactive is ()
 - a) Violet phosphorus
 - b) White phosphorus
 - c) Scarlet phosphorus
 - d) Red Phosphorus
5. Which of the following is not an element of first transition series? ()
 - a) Iron
 - b) Chromium
 - c) Magnesium
 - d) Nickel
6. Which group contains coloured ions out of the following? ()
 1. Cu^+ 2. Ti^{4+} 3. Co^{2+} 4. Fe^{2+}
 - a) 1, 2, 3, 4
 - b) 3, 4
 - c) 2, 3
 - d) 1, 2
7. In the equation of state of an ideal gas $PV = nRT$, the value of the () universal gas constant would depend only on
 - a) The pressure of the gas
 - b) The units of the measurement

- c) The nature of the gas
 d) None of these
8. At the same temperature and pressure, which of the following gases () will have the highest kinetic energy per mole
- a) Oxygen
 b) Methane
 c) Hydrogen
 d) All the same
9. On dissolving sugar in water at room temperature solution feels cool () to touch. Under which of the following cases dissolution of sugar will be most rapid?
- a) Sugar crystals in cold water.
 b) Sugar crystals in hot water.
 c) Powdered sugar in cold water.
 d) Powdered sugar in hot water.
10. Which one of the following is non-crystalline or amorphous? ()
- a) Diamond
 b) Graphite
 c) Glass
 d) Common Salt
11. Which of the following is a characteristic property of equilibrium? ()
- a) Number of moles of reactants and products is always equal.
 b) Catalyst affects the equilibrium state.
 c) It never proceeds to completion.
 d) Rate of forward and backward reactions are not equal. ()
12. Find the incorrect statement for a nucleophile
- a) A nucleophile is a Lewis acid
 b) Nucleophiles do not seek electron
 c) Ammonia is a nucleophile
 d) Nucleophiles attack low electron density sites ()
13. The correct name for the compound given below is:



- a) 2-methyl-1-butene
 b) 2-ethyl-1-propene
 c) 2-ethyl-1-pentane
 d) 3-methyl-2-butene

14. Alkene under high temperature and high-pressure forms ____ ()
- a) Alcohol
 - b) Polyalkyne
 - c) Polyalkane
 - d) Polyalkene
15. The phenomenon in which 2 or more structures, involving identical position of atoms can be written for a particular molecule, is called ____ ()
- a) Conjugation
 - b) Resonance
 - c) Hyper conjugation
 - d) Vibration
16. Which of the following is a permanent electron displacement effect? ()
- a) Inductomeric
 - b) Electromeric
 - c) Inductive
 - d) All of the mentioned
17. Tyndall effect confirms the ____ ()
- (a) gravity effect on the sol. particles
 - (b) light scattering by the sol. particles
 - (c) heterogeneous nature of sols.
 - (d) Brownian motion of the sol. particles
18. Which of the following can adsorb larger volume of hydrogen gas? ()
- (a) Finely divided platinum
 - (b) Colloidal solution of palladium
 - (c) Small pieces of palladium
 - (d) A single metal surface of platinum
19. Which of the following aqua regia is. ____ ()
- a) 1 part conc. HCl and 3 parts conc. HNO₃
 - b) 3 part conc. HCl and 1 part conc. HNO₃
 - c) 2 parts conc. HCl and 1 part conc. HNO₃
 - d) 3 parts conc. HCl and 2 parts conc. HNO₂
20. Which of the following can make difference in optical isomers? ()
- a) heat
 - b) temperature
 - c) polarized light
 - d) pressure

Answers:

1	A	11	C
2	A	12	A
3	C	13	A
4	D	14	C
5	C	15	B
6	B	16	C
7	B	17	C
8	D	18	B
9	D	19	B
10	C	20	C


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Phone: (0877) 2264599

E-mail:sgsartscollegettd@gmail.com

Web: sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Prof N. VENUGOPAL REDDY,
M.Sc., M.Phil., Ph.D
PRINCIPAL (FAC)



Tiruprti

Date:05.03.2021

From
Sri. D. Parameswara,
HOD of Chemistry,
SGS Arts College,
Tirupathi.

To
The Principal,
SGS Arts College,
Tirupathi.

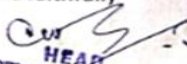
Sir,

Sub:- Permission for conducting Chemistry Bridge course to students of
first year degree – request – Regarding.
&&&&&

Kindly permit the Department of Chemistry to conduct Bridge
Course to the students of first year degree who studied Vocational Course in
intermediate for two days i.e. 08-03-2021 and 09-03-2021. Time table and
syllabus are enclosed.

Thanking you sir


Yours faithfully


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

**S.G.S Arts College, Tirupati
Department of Chemistry**

Circular

First-year Life Sciences students who studied vocational courses at the intermediate level are informed to attend the bridge course which will be held at seminar hall, SGS Degree College, Tirupati from 08-03-2021 to 09-03-2021.


Head,
Department of Chemistry
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI


Bridge Course in Chemistry Students List

S.No	Name of the student	Signature
1	J.Radhika	J. Radhika
2	N.Bhanu prasad	N. Bhanu Prasad
3	K.Kamal	K. Kamal
4	K.Chandu	K. Chandu
5	T.Vignesh yadav	T. Vignesh Yadav
6	P.Chand basha	P. Chand basha
7	M.Kishor naik	M. Kishor Naik
8	G. Satheesh	G. Satheesh

SIGNATURE
HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Department of Chemistry Bridge Course Time Table

Day	10AM-11AM	11AM-12NOON	12NOON-1PM	2PM-3PM	3PM-4PM
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Tuesday	Physical	Organic	General	Inorganic	


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

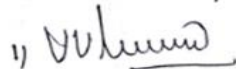
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4	General Chemistry	Stereo Chemistry Colloids Adsorption Analytical Chemistry

Department of Chemistry
Bridge Course Student Attendance List


S.No	Name of the student	Course Studied in Intermediate	Programme Joined in UG	8/3/21	9/3/21
1	J. Radhika	Diploma in agriculture	CBZ	X	X
2	N. Bhanu prasad	Diploma in agriculture	CBZ	X	X
3	K. Kamal	(Medical Lab Technician Diploma in seed	CBZ	X	X
4	K. Chandu		CBZ	X	X
5	T. Vignesh yadav	Diploma in agriculture	MZC	X	X
6	P. Chand basha	Diploma in agriculture	BBC	X	a
7	M. Kishor naik	Diploa in organic farming	MZC	X	X
8	G. sathesh	MLT (Medical Lab Technician Course)	MZC	X	X

Signature of the faculty

1) 

2) K. R. Srinivas

3) P. Srinivas


Head, Department of Chemistry
HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Bridge Course in Chemistry
Students Marks List

S.No	Name of the student	Marks
1	j.Radhika	13 - Thirteen only
2	N.Bhanu prasad	16 - Sixteen only
3	K.Kamal	14 - fourteen only
4	K.Chandu	15 - fifteen only
5	T.Vignesh yadav	14 - fourteen only
6	P.Chand basha	17 - seventeen only
7	M.Kishor naik	16 - sixteen only
8	G.Satheesh	13 - Thirteen only

SIGNATURE

HEAD

DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Bridge Course-Summary

Name of the programme	Bridge course in chemistry for vocational students
Duration of Course	Two days
Date	08-03-2021 to 09-03-2021.
Venue	Seminar Hall

S.No	Subject name	Faculty Name	No. of Students
1	Inorganic Chemistry	P Suguna	4
2	Physical Chemistry	V Venkata Lakshmi	4
3	Organic Chemistry	D Parameswara	4
4	General Chemistry	K. Revathi	4


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY
OBJECTIVE TEST - BRIDGE COURSE 2020-2021

Answer all questions

Maximum Marks: 20

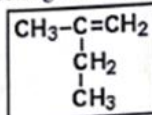
1. Which of the following is a characteristic property of equilibrium? ()

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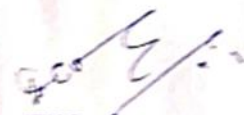
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c) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2$
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c) 2, 3
d) 1, 2
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c) Powdered sugar in cold water.
d) Powdered sugar in hot water.
20. Which one of the following is non-crystalline or amorphous? ()
a) Diamond
b) Graphite
c) Glass
d) Common Salt

ANSWERS:

1	C	11	A
2	A	12	A
3	A	13	C
4	C	14	D
5	B	15	C
6	C	16	B
7	C	17	B
8	B	18	D
9	B	19	D
10	C	20	C


HEAD
DEPARTMENT OF CHEMISTRY
S. S. JNTU ANANDRAO
TIRUPATI

1	C	11	A
2	A	12	A
3	A	13	C
4	C	14	D
5	B	15	C
6	C	16	B
7	C	17	B
8	B	18	D
9	B	19	D
10	C	20	C

HEAD
DEPARTMENT OF COMMERCE
U. S. S. AND T. BUREAU
TREASURY

3. ENDOWMENT PRIZES

With an objective to identify the best performing students, to motivate a students for achieving excellence in both curricular and extra-curricular activities our college has an institutionalized endowment prizes.



TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Students benefited by institutional level Prizes/Awards/ Medals for Academic achievements

Fund received from the donors for the ward of endowment cash prizes, awards, medals etc:

Existing total capital amount Rs 1,63,928 ,amount has been deposited as Fixed Deposit in Union Bank of India, the amount received as interest and the above deposit will be for the award of cash prize or award or medal who stood as Toppers.

1. **Ramesh Memorial** cash award: This awards instituted by his father Sri.G.Ramachandra Reddy Retired Chief Engineer .A.P to be awarded to the 3rd year BCom student who has secures highest aggregate marks in all semesters in Ist and 2nd years.
2. Prof. **G. Leela Krishna** Endowment Cash award: This award is instituted by Prof. **K.V.Naidu**, Former, Director of SVIMS, Tirupati as a token of Respect on his teacher, Prof. **G. Leela Krishna** to be awarded to the student of 3rdB.Scwho has secures the highest marks in Botany in all semesters of Ist year and 2nd year.
- 3 . **SmtPaluvai. Vijayalakshmi** Memorial cash award: This cash award of was instituted by her husband **Sri P. Subbarao** , Retd. HOD. Of Chemistry, S.V.Arts College, and former HOD. Of Chemistry of our College to be

awarded to (the student of III BSc with Chemistry combination of either medium scoring the highest total marks in all the semesters of Ist and 2nd year) to top rankers in Chemistry , one each at 2nd year and 3rd year who have secured the highest aggregate marks .

3. **Sri T. Surendranath Reddy** cash award: This prize is awarded to the topper in commerce who secures the highest marks in I & II year B.Com
4. **Dr. Dasa Raju, Rama Raju** endowment cash award: This prize is awarded to the student of 3rd year who got highest marks in Statistics I & II .
5. **Sri K.V.Gopalakrishna Murthy and Smt. K.Saraswathy** endowment cash award: This prize is awarded to the topper in English who has secured the highest aggregate marks in I and II Year.
6. **Mr. A. Goutham, Endowment** cash award: This prize is awarded to (The Topper in Mathematics who secures the highest marks in 2nd year BSc) to Top rankers in Mathematics, one each at Ist year and 2nd year, who have secured highest marks.
7. **Late Sri.P.Purushotham Naidu** Memorial endowment cash awards: The awards are instituted by his affectionate friends BA(HEP) 1992-1995 batch .This prize is awarded to EIGHT Top rankers one each at first year BA(HEP) and 2nd year BA(HEP).
First year BA (HPT, MES, ASA)
Second year BA (HEP, HPT, MES).

8. **Prof K.Sri Ramarao** Rtd HOD of Botany endowment prize: This prize is Top rankers one in Botany and the other in Microbiology who secures highest marks in all the semesters of First year & Second Year in subjects of Botany and Microbiology.

9. **S.G.S. Arts College** Science cash award: This prize is awarded to the student who secures the highest aggregate marks in all the subjects in all the semesters of BSc.

10. **Sri S.RangaRajan** cash award: This prize is awarded to the student of Maths combinations who secures the highest marks in all the group subjects of First and second year students.

Winners of the above prizes, Awards and Medals for their Academic Achievements.

2019-2020

Sno	Name of the Student	Group & Year	Marks/ Percentage	Hall Ticket No	Winner of the prize
1	K.Deepa	III Bcom-CA	87.3%	0218008014	Ramesh Memorial cash award, T. Surendranath Reddy cash award
2	S.Munikumar	III Bsc-BBC	551/600	0318008166	G. Leela Krishna Endowment Cash award
3	T.Avinash	III BSc-BBC	87.0%	0318008143	Smt Paluvai. Vijayalakshmi Memorial cash award
4	T.Avinash	III BSc-BBC	279/300	0318008143	Smt Paluvai. Vijayalakshmi Memorial cash award
5	A.Venkatesh	III BSc-MSCs	590/600	0318008121	Dasa Raju, Rama Raju endowment cash award
6	A.Venkatesh	III BSc-MSCs	89.3%	0318008121	S.RangaRajan cash award
7	B.Mohana Krishna	III BSc-MSCs	239/300	0318008311	K.V.Gopalakrishna Murthy and Smt. K.Saraswathy endowment cash ward
8	B.Mohana Krishna	III BSc-MSCs	197/200	0318008311	A. Goutham, Endowment cash ward
9	B.Srinivasulu	II BSC-MPCs	200/200	0319008675	A. Goutham, Endowment cash ward
10	K.Sireesha	II BSc-MPC	200/200	0319008172	A. Goutham, Endowment cash ward
11	G.Surendra	II BSc-MPC	200/200	0319008120	A. Goutham, Endowment cash ward

2018-2019

Sno	Name of the Student	Group & Year	Marks/ Percentage	Hall Ticket No	Winner of the prize
1	N.Purushotham	III Bcom-CA	86%	0217008579	Ramesh Memorial cash award, T. Surendranath Reddy cash award
2	S.Snehalatha	III Bsc-BBC	559/600	0317008681	G. Leela Krishna Endowment Cash award
3	P.Bharath Naveen	III BSc-BBC	550/600	0317008652	Smt Paluvai. Vijayalakshmi Memorial cash award
4	P.Guru sai	III BSc- MSCs	560/600	0317008675	Dasa Raju, Rama Raju endowment cash award
5	P.Gurusai	III BSc- MSCs	88%	0317008675	S.RangaRajan cash award
6	P.Shilpa	III BSc-MZC	245/300	0317008668	K.V.Gopalakrishna Murthy and Smt. K.Saraswathy endowment cash Award
7	S.Yerramala	II BA-HPT TM	84%	0117008019	Late.Purushottam Naidu Endowment cash ward
8	B.Pavan Kumar	II BA- HEP (EM)	84%	0117008505	Late.Purushottam Naidu Endowment cash ward
9	G.C.Gunasekhar	II BA-ASCA (EM)	83%	0117008513	Late.Purushottam Naidu Endowment cash ward
10	C.Sreenu	II BA-MES (EM)	87%	0117008510	Late.Purushottam Naidu Endowment cash ward
11	N.Purushotham	III Bcom-CA	86%	0217008579	T. Surendranath Reddy cash award

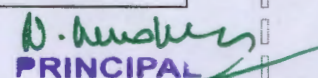
2017-2018

S. No	Name of the Student	Group & Year	Marks/ Percentage	Hall Ticket No	Winner of the prize
1	P.Gowtham Kumar	III BCom-CA	87.84%	0216008592	Ramesh Memorial cash award
2	N.Asthvadath	III BSc-BBC	475/550	0316008653	G. Leela Krishna Endowment Cash award
3	G.Venkateswarulu	III BSc-MPC(EM)	86.72%	0316008515	A. Goutham, Endowment cash ward
4	M.Sunil Kumar	III BSc-MPC(EM)	255/300	0316008523	A. Goutham, Endowment cash ward
5	N.Giridhar	III Bom	75.3%	0216008584	T. Surendranath Reddy cash award
6	B.Yuva Kishore	III BSc-MSCs	375/400	0316008709	Dasa Raju, Rama Raju endowment cash award
7	B.Muthyalu	II Bsc-MScs	81.0%	0316008572	K.V.Gopalakrishna Murthy and Smt. K.Saraswathy endowment cash ward
8	G.Venkateswarulu	III BSc-MPC(EM)	200/200	0316008515	A. Goutham, Endowment cash ward
9	P.Gurusai	II BSc-MSCs	195/200	0317008675	A. Goutham, Endowment cash ward
10	R.Kalairasi	II BSc-MSCs	195/200	0317008678	A. Goutham, Endowment cash ward
11	A.Ganesh	II BSc-MPCs	86.18%	0316008533	S.RangaRajan cash award

2016-2017

Sno	Name of the Student	Group & Year	Marks	Hall Ticket No	Winner of the prize
1	B.Bharath	III BCom-CA	1186/1300	215042814	Ramesh Memorial cash award
2	B.Bharath	III BCom-CA	355/500	215042814	T. Surendranath Reddy cash award
3	B.SuryaPrathap	III BSc- BBC	265/300	315042844	Ramesh Memorial cash award
4	B.SuryaPrathap	III BSc- BBC	118/150	315042844	SmtPaluvai. Vijayalakshmi Memorial cash award
5	A.Chandra Sekhar Naidu	III BSc-MZC	377/450	315042812	SmtPaluvai. Vijayalakshmi Memorial cash award
6	B.Masthan	III BSc-MSCs	187/300	315042771	Dasa Raju, Rama Raju endowment cash award
7	B.Masthan	III BSc--MSCs	831/900	315042771	S.G.S. Arts College Science cash award
8	G.Moulali	III BSc-MSCs	157/200	315042780	K.V.Gopalakrishna Murthy and Smt. K.Saraswathy endowment cash ward
9	R.Prathap	III BSc-MPC	300/300	315042726	A. Goutham, Endowment cash ward
10	B.Narendra Reddy	II BSc-MPCs	200/200	316008541	A. Goutham, Endowment cash ward
11	B.Masthan	II BSc-MScs	1099/1300	315042771	S.RangaRajan cash award




PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

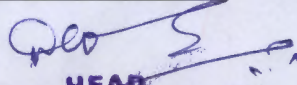
V-G 517

DEPARTMENT OF CHEMISTRY

Endowment Award Winners

The students of chemistry who excelled in their academics and award winners for their outstanding performance in S.V.U Examinations.

ACADEMIC YEAR 2016-17					
S.No.	Name of the Student	Group & Year	Marks	Hall Ticket No	Prize Winner
1	B.SuryaPrathap	III BSc- BBC	118/150	315042844	Smt. Paluvai. Vijayalakshmi Memorial cash award
2	A.Chandra Sekhar Naidu	III BSc-MZC	377/450	315042812	SmtPaluvai. Vijayalakshmi Memorial cash award
ACADEMIC YEAR 2018-19					
1	P.Bharath Naveen	III BSc-BBC	550/600	0317008652	Smt. Paluvai. Vijayalakshmi Memorial cash award
ACADEMIC YEAR 2019-20					
1	T.Avinash	III BSc-BBC	87.0%	0318008143	Smt Paluvai. Vijayalakshmi Memorial cash award
2	T.Avinash	III BSc-BBC	279/300	0318008143	Smt Paluvai. Vijayalakshmi Memorial cash award


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

4. Certificate Courses

No University curriculum can adequately cover all areas of importance and relevance. It is important for higher education institutions to supplement the curriculum to make students prepared better to meet the demands of industry and develop their own interest and attitudes. The institution offers a wide variety of short term certificate courses which are conducted out of college hours and semester breaks. The institution offers two types of courses 1. General interest courses which are open to all 2. Subject specific courses. The student can pursue these courses as per his/her required. The courses are conducted by many departments.

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e-mail: sgsacttd9@gmail.com
sgsartscollegeitd@gmail.com



TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

Prof N. VENUGOPAL REDDY M.Sc., M.Phil., Ph.D
PRINCIPAL (FAC)

1.2.2.1: How many Add on /Certificate programs are offered within the last 5 years.

S. No	Academic Year	Total Number of Certificate Courses
1	2020-21	Nil
2	2019-20	10
3	2018-19	10
4	2017-18	10
5	2016-17	10




PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

SGS ARTS COLLEGE, TIRUPATI

CERTIFICATE COURSES

Academic Year: 2019-20

BROCHURE

No University Curriculum can adequately cover all areas of importance and relevance. It is important for higher education institutions to supplement the curriculum to make students prepared better to meet the demands of industry and develop their own interests and aptitudes.

Our college offers a wide variety of short term certificate courses which are conducted out of college hours and semester breaks.

Our institution offers two types of courses.

1. General interest courses which are open to all.
2. Subject specific courses.

The student can pursue these courses as per his requirement. The courses are organized by many departments. Each course has a coordinator whose phone number is prominently displayed in the brochure.

Each course has its own assessment pattern, both practical and theoretical. Successful completion of the course earns you certification issued by the college.

S. No	Certificate Course	Organized by	Coordinator	Contact No
1	Photoshop	Computer Science	Sri P. Udaya	9440257300
2	Tally	Commerce	Dr. Y. Mallikarjuna Rao	9848533623
3	Archaeology	History	Sri A. Madhusudana Rao	9440506830
4	Floriculture	Botany	Smt T. Sakunthala	9052060327
5	Hotel & Hospitality Management	HMCT	Dr. N. Jagadeesh Kumar	8499840536
6	Soft skills	English	Dr. K. Madhu Kumar	9440257300
7	E-Commerce	Commerce	Dr. Y. Mallikarjuna Rao	9848533623
8	Maintenance of Household Electrical Appliance	Physics	Dr. K. Ravindranath Reddy	9440164872
9	MS-Office	Computer Science	Sri P. Udaya	9440257300
10	Vedic Mathematics	Mathematics	Sri P. Munibhaskar Rao	9989305142

Note: For admission contact coordinator for the respective course.

As the basic idea is to make ample opportunities available to students, there are no restrictions on how many courses a student can join and one student can acquire several certifications during the period of degree course.


PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

SGS ARTS COLLEGE, TIRUPATI
Certificate Courses
Academic Year: 2019-20

The following departments are conducting the certificate courses. The students those who are interested to study the certificate courses are requested to consult the HOD of the following department.

S. No	Department	Name of the Course	Eligibility of the Students
1	Computer Science	Photoshop	All Graduate students
2	Commerce	Tally	B.Com & B.A ASCA students
3	History	Archaeology	B.A H.E.P & H.P.T students
4	Botany	Floriculture	All Life Science Group students
5	HMCT	Hotel & Hospitality Management	HMCT students
6	English	Soft skills	All Graduate students
7	Commerce	E-Commerce	B.Com & B.A ASCA students
8	Physics	Maintenance of House hold Electrical Appliance	All B. Sc students
9	Computer Science	MS-Office	All B. Sc students
10	Mathematics	Vedic Mathematics	B. Sc M.S.Cs, M.P.Cs, M.P.S, M.P.C & M.E.S students


PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE : MS-OFFICE
ACADEMIC YEAR : 2019-2020

Permission Letter

SGSAC/Dept.of Computer Science/HOD/Certificate Course/2019-20

From

Date: 14-01-2020.

HOD,
Department of Computer Science,
S.G.S Arts College,
Tirupati.

To

The Principal,
S.G.S Arts College,
Tirupati.

Respected Sir,

Sub: Requesting for permission to conduct Certificate Course in **MS-Office** – Regarding.

We have planned to conduct a forty hours Certificate Course on **MS-Office** for two batches of Second Year students from **20-01-2020 to 13-02-2020**. Therefore kindly permit us to conduct this Certificate Course.

Thanking You,

Yours Faithfully,



CIRCULAR

It is to inform all the faculty of **Department of Computer Science** to attend staff meeting at **11:00 PM** in the HOD chamber to discuss the conduct of Certificate Courses for the academic year **2019-20.**

Faculty Members Attended:

HOD	Name of the Faculty	Signature
Sri P.Udaya	1. Smt B.Triveni 2. Smt N.Jayalakshmi 3. Sri D.Chaitanya Kumar	B. Triveni N. Jayalakshmi D. Chaitanya Kumar

Agenda: Certificate courses to be conducted for the Academic Year 2019-20.

All the staff members of the department discussed the agenda. The following resolutions were passed unanimously.

- It was decided to conduct a certificate course on **MS-Office** to the Second Year students in the first week of **September** month.
- It was decided to conduct a certificate course on **MS-Office** to the Second year students in the first week of **February** month.
- HOD offered vote of thanks.

Faculty:

1. Smt B.Triveni
2. Smt N.Jayalakshmi
1. Sri D.Chaitanya Kumar

Signature:

B. Triveni
N. Jayalakshmi
D. Chaitanya Kumar

Sri P. Udaya

NOTICE

SGSAC/Dept.of Computer Science/HOD/Certificate Course/2019-20

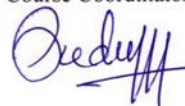
It is to inform to all the **II Year B.Sc MPCS, MSCS, MPC, MPS** students that we are going to conduct a Certificate Course. The details of the course offered are as follows.

Course name	<u>MS-Office</u>
Course date & duration	20/January/2020 : 40 Hours
Timings	<u>Batch 1:</u> Theory: 8:00 AM to 9:00 AM Practical: 4:00 PM to 5:00 PM <u>Batch 2:</u> Theory: 5:00 PM to 6:00 PM Practical: 8:00 AM to 9:00 AM
Course Mentors	<u>Batch 1:</u> B.Triveni, D.Chaitanya Kumar <u>Batch 2:</u> N.Jayalakshmi, D.Chaitanya Kumar

Students who are interested to join the course can contact the Coordinator of Certificate Course Sri P.Udaya for further details of course content, study material and prerequisites of the course.


HOD

Course Coordinator



Details of the Course

Name of the Certificate Course	:	MS-Office
Batch	:	1
Timing	:	Theory: 8:00 AM to 9:00 AM Practical: 4:00 PM to 5:00 PM
No. of students/ Dept/Year/ Course Duration:	:	60/ Comp.Science/ 2019-20/ 40 Hours
Venue	:	Theory: Room No 215 Practical: Comp.Science Lab
Eligible Groups	:	II Year B.Sc MPCS, MSCS, MPC, MPS
Trainers	:	B.Triveni, D.Chaithanya Kumar
Certificate	:	Yes
Course Coordinator	:	Sri P.Udaya



Details of the Course

Name of the Certificate Course	:	MS-Office
Batch	:	2
Timing	:	Theory: 4:00 PM to 5:00 PM Practical: 8:00AM to 9:00AM
No. of students/ Dept/Year/ Course Duration:	:	60/ Comp.Science/ 2019-20/ 40 Hours
Venue	:	Theory: Room No 215 Practical: Comp.Science Lab
Eligible Groups	:	II Year B.Sc MPCS, MSCS, MPC, MPS
Trainers	:	N.Jayalakshmi, D.Chaithanya Kumar
Certificate	:	Yes
Course Coordinator	:	Sri P.Udaya



**S.G.S.ARTS COLLEGE, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN MS-Office**

Course Contents

Theory (20 Hours)

MS WORD

Module 1: Text Basics, Text Formatting and saving file, Working with Objects

- Typing the text, Alignment of text
- Editing Text: Cut, Copy, Paste, Select All, Clear
- Find & Replace
- New, Open, Close, Save, Save As
- Formatting Text: Font Size, Font Style
- Font Colour, Use the Bold, Italic, and Underline
- Change the Text Case
- Line spacing, Paragraph spacing
- Shading text and paragraph
- Working with Tabs and Indents
- Shapes, Clipart and Picture, Word Art, Smart Art
- Columns and Orderings - To Add Columns to a Document
- Change the Order of Objects
- Page Number, Date & Time
- Inserting Text boxes
- Inserting Word art
- Inserting symbols
- Inserting Chart

Module 2: Header & Footers, Working with bullets and numbered lists

- Inserting custom Header and Footer
- Inserting objects in the header and footer
- Add section break to a document
- Multilevel numbering and Bulleting
- Creating List
- Customizing List style
- Page bordering
- Page background

Module 3: Tables, Styles and Content

- Working with Tables, Table Formatting
- Table Styles
- Alignment option
- Merge and split option
- Using Build- in Styles, Modifying Styles
- Creating Styles, Creating a list style
- Table of contents and references
- Adding internal references

- Adding a Footnote
- Adding Endnote

Module 4: Merging Documents

- Typing new address list
- Importing address list from Excel file
- Write and insert field
- Merging with outlook contact
- Preview Result
- Merging to envelopes
- Merging to label
- Setting rules for merges
- Finish & Merge options

Module 5: Proofing the document, Printing

- Check Spelling As You Type.
- Mark Grammar Errors As You Type.
- Setting AutoCorrect Options
- Page Setup, Setting margins
- Print Preview, Print

MS EXCEL

Module 6: Introduction to Excel, Formatting excel work book:

- Introduction to Excel interface
- Understanding rows and columns, Naming Cells
- Working with excel workbook and sheets
- New, Open, Close, Save, Save As
- Formatting Text: Font Size, Font Style
- Font Color, Use the Bold, Italic, and Underline
- Wrap text, Merge and Centre
- Currency, Accounting and other formats
- Modifying Columns, Rows & Cells

Module 7: Perform Calculations with Functions

- Creating Simple Formulas
- Setting up your own formula
- Date and Time Functions, Financial Functions
- Logical Functions, Lookup and Reference
- Functions Mathematical Functions
- Statistical Functions, Text Functions.

Module 8: Sort and Filter Data with Excel

- Sort and filtering data
- Using number filter, Text filter
- Custom filtering
- Removing filters from columns
- Conditional formatting

Module 9: Create Effective Charts to Present Data Visually

- Inserting Column, Pie chart etc.
- Create an effective chart with Chart Tool
- Design, Format, and Layout options
- Adding chart title
- Changing layouts
- Chart styles
- Editing chart data range
- Editing data series
- Changing chart

Module 10: Proofing and Printing

- Page setup, Setting print area, Print titles
- Inserting custom Header and Footer
- Inserting objects in the header and footer
- Page Setup, Setting margins, Print Preview, Print
- Enable back ground error checking
- Setting AutoCorrect Options

MS POWERPOINT

Module 11: Setting Up PowerPoint Environment:

- New, Open, Close, Save, Save As
- Typing the text, Alignment of text
- Formatting Text: Font Size, Font Style
- Font Color, Use the Bold, Italic, and Underline
- Cut, Copy, Paste, Select All, Clear text
- Find & Replace
- Working with Tabs and Indents

Module 12: Creating slides and applying themes

- Inserting new slide
- Changing layout of slides
- Duplicating slides
- Copying and pasting slide
- Applying themes to the slide layout
- Changing theme color
- Slide background
- Formatting slide background
- Using slide views

Module 13: Working with bullets and numbering

- Multilevel numbering and Bulleting
- Creating List
- Page bordering
- Page background
- Aligning text
- Text directions
- Columns option

Module 14: Working with Objects

- Shapes, Clipart and Picture, Word Art, Smart Art
- Change the Order of Objects
- Inserting slide header and footer
- Inserting Text boxes
- Inserting shapes, using quick styles
- Inserting Word art
- Inserting symbols
- Inserting Chart

Module 15: Hyperlinks and Action Buttons

- Inserting Hyperlinks and Action Buttons
- Edit Hyperlinks and Action Button
- Word Art and Shapes

Module 16: Working With Movies and Sounds

- Inserting Movie From a Computer File
- Inserting Audio file
- Audio Video playback and format options
- Video options, Adjust options
- Reshaping and bordering Video

Module 17: Using SmartArt and Tables

- Working with Tables, Table Formatting
- Table Styles
- Alignment option
- Merge and split option
- Converting text to smart art

Module 18: Animation and Slide Transition

- Default Animation, Custom Animation
- Modify a Default or Custom Animation
- Reorder Animation Using Transitions
- Apply a Slide Transition, Modifying a Transition, Advancing to the Next Slide

Module 19: Using slide Master

- Using slide master
- Inserting layout option
- Creating custom layout
- Inserting place holders
- Formatting place holders

Module 20: Slide show option

- Start slide show
- Start show from the current slide
- Rehearse timing
- Creating custom slide show
- Check Spelling As You Type
- Setting AutoCorrect Options
- Save as video
- Save as JPEG files
- Save as PowerPoint Show file
- Print Preview, Print

Odeyff

Practical (20 Hrs): Based upon above content

Duration of the Course: 40 Hours

Theory : 20 Hours

Practical : 20 Hours

Eligibility : II Year BSc MPCS, MSCS, MPC, MPS Groups

Intake : 60 Students /Batch

Head of the Department : Sri P.Udaya

Course Mentors :

1. Smt B.Triveni
2. Smt N.Jayalakshmi
3. Sri D.Chaithanya Kumar



S.G.S.ARTS COLLEGE, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN MS-Office
Academic Year: 2019 -2020

BATCH 1: TIME TABLE

20 Jan 2020 to 13 Feb 2020

DAY/TIME	THEORY	PRACTICAL
	8.00 – 9.00 AM	4.00 – 5.00 PM
MONDAY	BT	DCK
TUESDAY	DCK	BT
WEDNESDAY	BT	DCK
THURSDAY	DCK	BT
FRIDAY	BT	DCK
SATURDAY	DCK	BT


HOD

BT: B.Triveni
DCK: D.Chaitanya Kumar

S.G.S.ARTS COLLEGE, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN MS-Office
Academic Year: 2019 -2020

BATCH 1: TIME TABLE

20 Jan 2020 to 13 Feb 2020

DAY/TIME	THEORY	PRACTICAL
	4.00 – 5.00 PM	8.00 – 9.00 AM
MONDAY	NJ	DCK
TUESDAY	DCK	NJ
WEDNESDAY	NJ	DCK
THURSDAY	DCK	NJ
FRIDAY	NJ	DCK
SATURDAY	DCK	NJ


HOD

NJ : N.Jayalakshmi
DCK: D.Chaitanya Kumar

Department of Computer Science:

MS-Office Certificate Course Students Name List Batch 1

S.No	Name of the Student	Class	Roll No
1	Anil Kumar Mukku	II B.Sc(MPCS)-EM	18511
2	Anil Kumar Talari	II B.Sc(MPCS)-EM	18512
3	Balaji.E	II B.Sc(MPCS)-EM	18513
4	Balanjineyulu Mangala	II B.Sc(MPCS)-EM	18514
5	Balarama Krishna Yadav A	II B.Sc(MPCS)-EM	18515
6	Charan Kuamar Mandati	II B.Sc(MPCS)-EM	18516
7	Chenchaiiah Gunji	II B.Sc(MPCS)-EM	18517
8	Dinesh Kumar Mangali	II B.Sc(MPCS)-EM	18518
9	Guru Sai Kumar.P.V	II B.Sc(MPCS)-EM	18519
10	Harshavardhan Reddy B	II B.Sc(MPCS)-EM	18520
11	Hemanth Kumar Paipuru	II B.Sc(MPCS)-EM	18521
12	Hemasri Kuppa	II B.Sc(MPCS)-EM	18522
13	Jagan Goud Nandimalla	II B.Sc(MPCS)-EM	18523
14	Jayalakshmi Kailasam	II B.Sc(MPCS)-EM	18524
15	Mabbu Sharif Gajuguduru	II B.Sc(MPCS)-EM	18525
16	Mahendra Dasari	II B.Sc(MPCS)-EM	18526
17	Manendra Kumar R	II B.Sc(MPCS)-EM	18527
18	Mohan Krishna Kotekallu	II B.Sc(MPCS)-EM	18528
19	Mounika Gali	II B.Sc(MPCS)-EM	18529
20	Mounika Sanambatla	II B.Sc(MPCS)-EM	18530
21	Nagadinesh Nagisetty	II B.Sc(MPCS)-EM	18531
22	Narasimhudu Kukka	II B.Sc(MPCS)-EM	18532
23	Naveen Kumar Kanchi	II B.Sc(MPCS)-EM	18533
24	Purushotham Reddy B	II B.Sc(MPCS)-EM	18534
25	Raj Kumar Giddaiahgari	II B.Sc(MPCS)-EM	18535
26	Rajesh Golla	II B.Sc(MPCS)-EM	18536
27	Rajesh Pannamareddi	II B.Sc(MPCS)-EM	18537
28	Rakesh Gajjala	II B.Sc(MPCS)-EM	18538
29	Rasool Yendhoti	II B.Sc(MPCS)-EM	18539

30	Ravi Teja Kamineni	II B.Sc(MPCS)-EM	18540
31	Saiteja Padma	II B.Sc(MPCS)-EM	18541
32	Sandhya Devarakonda	II B.Sc(MPCS)-EM	18542
33	Sandhya Putturu	II B.Sc(MPCS)-EM	18543
34	Sasidhar Bottla	II B.Sc(MPCS)-EM	18544
35	Shaik Yasmin	II B.Sc(MPCS)-EM	18546
36	Shireesha Gounori	II B.Sc(MPCS)-EM	18547
37	Siva Sankar Thati	II B.Sc(MPCS)-EM	18548
38	Sravan Kumar Chagalamarri	II B.Sc(MPCS)-EM	18549
39	Sreenivasulu Boya	II B.Sc(MPCS)-EM	18550
40	Subbarayudu Bayanaboina	II B.Sc(MPCS)-EM	18551
41	Suresh Kota	II B.Sc(MPCS)-EM	18552
42	Suresh Mangali	II B.Sc(MPCS)-EM	18553
43	Aravind Obalavaram	II B.Sc(MPC)-TM	18404
44	Chandrasekhar.B	II B.Sc(MPC)-TM	18405
45	Dathathreyudu Bandarupalli	II B.Sc(MPC)-TM	18406
46	Hareesh Derangi	II B.Sc(MPC)-TM	18407
47	Munisekhar Pamanji	II B.Sc(MPC)-TM	18408
48	Sunil Chemuru	II B.Sc(MPC)-TM	18409
49	Suresh Jangamreddy	II B.Sc(MPC)-TM	18411
50	Vishnuvardhan Chalampalem	II B.Sc(MPC)-TM	18412
51	Bharadhwaz Yannam	II B.Sc(MPC)-EM	18441
52	Giri Kondari	II B.Sc(MPC)-EM	18442
53	Govardhan Dattam	II B.Sc(MPC)-EM	18443
54	Hemadri.P	II B.Sc(MPC)-EM	18444
55	Hemantha Kumar Aare	II B.Sc(MPC)-EM	18445
56	Kalathur Vetti Nithish	II B.Sc(MPC)-EM	18446
57	Karimulla Pinjari	II B.Sc(MPC)-EM	18447
58	Mallikarjuna Avula	II B.Sc(MPC)-EM	18448
59	Mahesh Vanka	II B.Sc(MPC)-EM	18449
60	Manjunath Bhagathagalla	II B.Sc(MPC)-EM	18450

Qudus

Department of Computer Science:

MS-Office Certificate Course Students Name List Batch 2

S.No	Name of the Student	Class	Roll No
1	Ajay Kumar Bandar	II B.Sc(MSCS)-EM	18565
2	Ajay Kumar Medikonda	II B.Sc(MSCS)-EM	18566
3	Adishesu Mabbu	II B.Sc(MSCS)-EM	18567
4	Anitha Thalari	II B.Sc(MSCS)-EM	18568
5	Aravind Naidu Buraga	II B.Sc(MSCS)-EM	18569
6	Balaji Ballikura	II B.Sc(MSCS)-EM	18570
7	Bharath Kumar Paidikalva	II B.Sc(MSCS)-EM	18571
8	Bhaskar Kamatham	II B.Sc(MSCS)-EM	18572
9	Chakri Paipuri	II B.Sc(MSCS)-EM	18573
10	Charan Ulasala	II B.Sc(MSCS)-EM	18574
11	Chandini Velagambeti	II B.Sc(MSCS)-EM	18575
12	Charan Tej Markapuram	II B.Sc(MSCS)-EM	18576
13	Deepa Kadirimangalam	II B.Sc(MSCS)-EM	18577
14	Dhana Kumar Kota	II B.Sc(MSCS)-EM	18578
15	Eswar Kanakappagari	II B.Sc(MSCS)-EM	18579
16	Eswara Narasimha Vara Prasad	II B.Sc(MSCS)-EM	18580
17	Govardhan Krishna Vallepu	II B.Sc(MSCS)-EM	18581
18	Guru Swamy Velampati	II B.Sc(MSCS)-EM	18582
19	Hari Prasad Mala	II B.Sc(MSCS)-EM	18583
20	Haritha Pujari	II B.Sc(MSCS)-EM	18584
21	Hemanth Rajolu	II B.Sc(MSCS)-EM	18585
22	Indhumathi Lalapeta	II B.Sc(MSCS)-EM	18586
23	Jyothi Kuruva	II B.Sc(MSCS)-EM	18587
24	Kishore Kovuru	II B.Sc(MSCS)-EM	18588
25	Lakshmi Manne	II B.Sc(MSCS)-EM	18589
26	Leelakrishna Gadiboina	II B.Sc(MSCS)-EM	18590
27	Manoj Kumar Pallipattu	II B.Sc(MSCS)-EM	18591
28	Manoj Naramala	II B.Sc(MSCS)-EM	18592

29	Nanda Kumar Kala	II B.Sc(MSCS)-EM	18593
30	Pavan Kumar Reddy Vengareddy	II B.Sc(MSCS)-EM	18594
31	Penchala Narasimhulu Amooru	II B.Sc(MSCS)-EM	18595
32	Ramesh Gunda	II B.Sc(MSCS)-EM	18596
33	Ramu Goud Ramagouni	II B.Sc(MSCS)-EM	18597
34	Rithik.K.M	II B.Sc(MSCS)-EM	18598
35	Shireesha Kuruva	II B.Sc(MSCS)-EM	18599
36	Shyam Kumar Pokuru	II B.Sc(MSCS)-EM	18600
37	Siva Sankar Bijigam	II B.Sc(MSCS)-EM	18601
38	Sreedevi Chittoor	II B.Sc(MSCS)-EM	18602
39	Sreekanth.K	II B.Sc(MSCS)-EM	18603
40	Suresh Madigundu	II B.Sc(MSCS)-EM	18604
41	Thimmaiah Siruvella	II B.Sc(MSCS)-EM	18605
42	Vamsi Krishna Naidu Kommineni	II B.Sc(MSCS)-EM	18606
43	Venkata Amaranath Malasani	II B.Sc(MSCS)-EM	18607
44	Venkata Anirudh Reddy Venna	II B.Sc(MSCS)-EM	18608
45	Venkata Chanukya Kethineni	II B.Sc(MSCS)-EM	18609
46	Venkat Nisanth Kanupuru	II B.Sc(MSCS)-EM	18610
47	Venkateswarlu Bingi	II B.Sc(MSCS)-EM	18611
48	Venu Bharath.R.T	II B.Sc(MSCS)-EM	18612
49	Vijya Goud Ediga	II B.Sc(MSCS)-EM	18613
50	Vinay Sada	II B.Sc(MSCS)-EM	18614
51	Aravind Chakravarthy	II B.Sc(MPS)-TM	18474
52	Dastagiri Boya	II B.Sc(MPS)-TM	18476
53	Erri Swamy Getu	II B.Sc(MPS)-TM	18477
54	Marimeni Niranjan	II B.Sc(MPS)-TM	18478
55	Ramesh Boya	II B.Sc(MPS)-TM	18480
56	Yuva raj Kante	II B.Sc(MPS)-TM	18481
57	Nanda Kumar.S	II B.Sc(MPC)-EM	18451
58	Narendra Mattem	II B.Sc(MPC)-EM	18452
59	Obul Reddy Boggala	II B.Sc(MPC)-EM	18453
60	Pavan Kumar.N	II B.Sc(MPC)-EM	18454

Reddy

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"

STUDENTS ATTENDANCE LIST BATCH 1 (THEORY) 8:00 AM to 9:00 AM

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
1	18511	Anil Kumar Mukku	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P
2	18512	Anil Kumar Talari	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P
3	18513	Balaji.E	P	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P	P	P	P	P
4	18514	Balanjineyulu Mangala	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P
5	18515	Balarama Krishna Yadav A	P	P	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P	P	P	P
6	18516	Charan Kuamar Mandati	P	A	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P
7	18517	Chenchaiah Gunji	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
8	18518	Dinesh Kumar Mangali	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P
9	18519	Guru Sai Kumar.P.V	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
10	18520	Harshavardhan Reddy B	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
11	18521	Hemanth Kumar Paipuru	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P
12	18522	Hemasri Kuppa	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
13	18523	Jagan Goud Nandimalla	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
14	18524	Jayalakshmi Kailasam	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
15	18525	Mabbu Sharif Gajuguduru	P	P	P	A	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P
16	18526	Mahendra Dasari	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	A
17	18527	Manendra Kumar R	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	A
18	18528	Mohan Krishna Kotekallu	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
19	18529	Mounika Gali	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P
20	18530	Mounika Sanambatla	P	A	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
21	18531	Nagadinesh Nagisetty	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	A

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
22	18532	Narasimhudu Kukka	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P
23	18533	Naveen Kumar Kanchi	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
24	18534	Purushotham Reddy B	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P
25	18535	Raj Kumar Giddaiahgari	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P
26	18536	Rajesh Golla	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	A	P
27	18537	Rajesh Pannamareddi	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P
28	18538	Rakesh Gajjala	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	A
29	18539	Rasool Yendhoti	P	P	A	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P
30	18540	Ravi Teja Kamineni	P	P	P	P	P	A	P	P	P	A	P	P	P	P	P	P	P	P	P	P
31	18541	Saiteja Padma	P	P	P	P	P	P	A	A	P	P	P	P	P	P	P	P	P	P	P	P
32	18542	Sandhya Devarakonda	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A
33	18543	Sandhya Putturu	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
34	18544	Sasidhar Bottla	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	A	P	P	P	P
35	18546	Shaik Yasmin	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P
36	18547	Shireesha Gounori	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
37	18548	Siva Sankar Thati	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
38	18549	Sravan Kumar Chagalamarri	P	P	P	P	A	P	P	A	P	P	P	P	P	P	P	P	P	P	A	P
39	18550	Sreenivasulu Boya	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P	P	P
40	18551	Subbarayudu Bayanaboina	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
41	18552	Suresh Kota	P	P	P	P	P	A	A	P	P	P	P	P	P	P	P	P	P	P	P	P
42	18553	Suresh Mangali	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P
43	18404	Aravind Obalavaram	P	P	P	A	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
44	18405	Chandrasekhar.B	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	A	P	P
45	18406	Dathathreyudu Bandarupalli	P	P	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	A	P	P

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
46	18407	Hareesh Derangi	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P
47	18408	Munisekhar Pamanji	P	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P	P	P	P
48	18409	Sunil Chemuru	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P
49	18411	Suresh Jangamreddy	P	P	P	P	A	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P
50	18412	Vishnuvardhan Chalampalem	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P	P
51	18441	Bharadhwaz Yannam	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P
52	18442	Giri Kondari	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P
53	18443	Govardhan Dattam	P	P	P	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P	P	P
54	18444	Hemadri.P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
55	18445	Hemantha Kumar Aare	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P
56	18446	Kalathur Vetti Nithish	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P
57	18447	Karimulla Pinjari	P	P	P	P	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P	P
58	18448	Mallikarjuna Avula	P	P	P	P	P	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P
59	18449	Mahesh Vanka	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	A
60	18450	Manjunath Bhagathagalla	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P

Reddy

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"

STUDENTS ATTENDANCE LIST BATCH 1 (PRACTICAL) 4:00 PM to 5:00 PM

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
1	18511	Anil Kumar Mukku	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
2	18512	Anil Kumar Talari	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P	P
3	18513	Balaji.E	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P	P	P	P	P
4	18514	Balanjineyulu Mangala	P	P	A	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P
5	18515	Balarama Krishna Yadav A	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P
6	18516	Charan Kuamar Mandati	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P
7	18517	Chenchaiah Gunji	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P
8	18518	Dinesh Kumar Mangali	P	P	P	P	A	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P
9	18519	Guru Sai Kumar.P.V	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
10	18520	Harshavardhan Reddy B	P	P	P	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P	P	P
11	18521	Hemanth Kumar Paipuru	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	A
12	18522	Hemasri Kuppa	P	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P
13	18523	Jagan Goud Nandimalla	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P	P	P	P	P	P
14	18524	Jayalakshmi Kailasam	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P
15	18525	Mabbu Sharif Gajuguduru	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
16	18526	Mahendra Dasari	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P
17	18527	Manendra Kumar R	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A
18	18528	Mohan Krishna Kotekallu	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A
19	18529	Mounika Gali	P	P	P	P	A	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
20	18530	Mounika Sanambatta	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P	P	P
21	18531	Nagadinesh Nagisetty	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P
22	18532	Narasimhudu Kukka	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P
23	18533	Naveen Kumar Kanchi	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P
24	18534	Purushotham Reddy B	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P
25	18535	Raj Kumar Giddaiahgari	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P
26	18536	Rajesh Golla	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P
27	18537	Rajesh Pannamareddi	P	P	P	P	P	P	A	P	P	P	P	P	P	P	A	P	P	P	P	P
28	18538	Rakesh Gajjala	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P
29	18539	Rasool Yendhoti	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P
30	18540	Ravi Teja Kamineni	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P
31	18541	Saiteja Padma	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P
32	18542	Sandhya Devarakonda	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	A	P	P
33	18543	Sandhya Putturu	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P
34	18544	Sasidhar Bottla	P	P	P	A	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P
35	18546	Shaik Yasmin	P	A	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P
36	18547	Shireesha Gounori	P	P	P	A	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
37	18548	Siva Sankar Thati	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P	P	P	P	P	P
38	18549	Sravan Kumar Chagalamarri	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P	P	A	P	P
39	18550	Sreenivasulu Boya	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P
40	18551	Subbarayudu Bayanaboina	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
41	18552	Suresh Kota	P	P	P	P	P	P	P	P	P	A	P	A	P	P	P	P	P	P	P	P

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
42	18553	Suresh Mangali	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
43	18404	Aravind Obalavaram	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
44	18405	Chandrasekhar.B	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
45	18406	Dathathreyudu Bandarupalli	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
46	18407	Hareesh Derangi	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
47	18408	Munisekhar Pamanji	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
48	18409	Sunil Chemuru	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
49	18411	Suresh Jangamreddy	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
50	18412	Vishnuvardhan Chalampalem	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
51	18441	Bharadhwaz Yannam	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
52	18442	Giri Kondari	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
53	18443	Govardhan Dattam	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
54	18444	Hemadri.P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
55	18445	Hemantha Kumar Aare	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
56	18446	Kalathur Vetti Nithish	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
57	18447	Karimulla Pinjari	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
58	18448	Mallikarjuna Avula	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
59	18449	Mahesh Vanka	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
60	18450	Manjunath Bhagathagalla	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P

Reddy

NOTICE

SGSAC/Dept.of Computer Science/HOD/Certificate Course/2019-20

All the Second Year B.Sc MPCS, MSCS, MPC(EM & TM), MPS Batch 1 and Batch 2 enrolled students of Certificate Course in **MS-Office** are hereby informed that the **Final** Examination will be held on 15-02-2020.

Date: 13-02-2020


HOD

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
STUDENTS TEST ATTENDANCE BATCH 1

Date: 15-02-2020

Timinig: 4:00 to 5:00 PM

S.No	Roll No	Name of the Student	Signature
1	18511	Anil Kumar Mukku	M. Anil Kumar
2	18512	Anil Kumar Talari	T. Anil Kumar
3	18513	Balaji.E	Balaji.E
4	18514	Balanjineyulu Mangala	M. Balanjineyulu
5	18515	Balarama Krishna Yadav A	A. Balakrishna Yadav
6	18516	Charan Kuamar Mandati	Charan Kumar Mandati
7	18517	Chenchalah Gunji	Ch. Chenchalah
8	18518	Dinesh Kumar Mangali	M. Dinesh Kumar
9	18519	Guru Sai Kumar.P.V	P.V. Guru Sai Kumar
10	18520	Harshavardhan Reddy B	B. Harshavardhan Reddy
11	18521	Hemanth Kumar Paipuru	P. Hemanth Kumar
12	18522	Hemasri Kuppa	K. Hemasri
13	18523	Jagan Goud Nandimalla	N. Jagan Goud
14	18524	Jayalakshmi Kailasam	K. Jayalakshmi
15	18525	Mabbu Sharif Gajuguduru	G. Mabbu Sharif
16	18526	Mahendra Dasari	Mahendra Dasari
17	18527	Manendra Kumar R	R. Manendra Kumar
18	18528	Mohan Krishna Kotekallu	Mohan Krishna Kotekallu
19	18529	Mounika Gali	G. Mounika
20	18530	Mounika Sanambatla	S. Mounika
21	18531	Nagadinesh Nagisetty	N. Nagadinesh
22	18532	Narasimhudu Kukka	K. Narasimhudu
23	18533	Naveen Kumar Kanchi	K. Naveen Kumar
24	18534	Purushotham Reddy B	Purushotham Reddy B
25	18535	Raj Kumar Giddaiahgari	Raj Kumar Giddaiahgari
26	18536	Rajesh Golla	G. Rajesh
27	18537	Rajesh Pannamareddi	P. Rajesh
28	18538	Rakesh Gajjala	G. Rakesh
29	18539	Rasool Yendhoti	Y. Rasool
30	18540	Ravi Teja Kamineni	K. Ravi Teja
31	18541	Saiteja Padma	P. Saiteja
32	18542	Sandhya Devarakonda	D. Sandhya
33	18543	Sandhya Putturu	Sandhya putturu
34	18544	Sasidhar Bottla	S. Sasidhar Bottla
35	18546	Shaik Yasmin	Yasmin. S
36	18547	Shireesha Gounori	G. Shireesha

S.No	Roll No	Name of the Student	Signature
37	18548	Siva Sankar Thati	T. Siva Sankar
38	18549	Sravan Kumar Chagalamarri	C. Sravan Kumar
39	18550	Sreenivasulu Boya	B. Sreenivasulu
40	18551	Subbarayudu Bayanaboina	B. Subbarayudu
41	18552	Suresh Kota	Suresh Kota
42	18553	Suresh Mangali	M. Suresh
43	18404	Aravind Obalavaram	O. Aravind
44	18405	Chandrasekhar B	B. Chandrasekhar
45	18406	Dathathreyudu Bandrupalli	B. Dathathreyudu
46	18407	Hareesh Derangi	D. Hareesh
47	18408	Munisekhar Pamanji	Munisekhar Pamanji
48	18409	Sunil Chemuru	Sunil Chemuru
49	18411	Suresh Jangamreddy	Suresh J
50	18412	Vishnuvardhan Chalampalem	C. Vishnuvardhan
51	18441	Bharadhwaz Yannam	Y. Bharadhwaz
52	18442	Giri Kondari	Giri Kondari
53	18443	Govardhan Dattam	D. Govardhan
54	18444	Hemadri P	P. Hemadri
55	18445	Hemantha Kumar Aare	A. Hemanth Kumar
56	18446	Kalathur Vetti Nithish	Kalathur Vetti Nithish
57	18447	Karimulla Pinjari	P. Karimulla
58	18448	Mallikarjuna Avula	A. Mallikarjuna
59	18449	Mahesh Vanka	V. Mahesh
60	18450	Manjunath Bhagathagalla	B. Manjunath

Reddy

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
STUDENTS TEST ATTENDANCE BATCH 2

Date: 15-01-2020

Timing: 4:00 to 5:00 PM

S.No	Roll No	Name of the Student	Signature
1	18565	Ajay Kumar Bandar	B. Ajay Kumar
2	18566	Ajay Kumar Medikonda	M. Ajay Kumar
3	18567	Adishesu Mabbu	Adishesu mabbu
4	18568	Anitha Thalari	T. Anitha
5	18569	Aravind Naidu Buraga	B. Aravind Naidu
6	18570	Balaji Ballikura	B. Balaji
7	18571	Bharath Kumar Paidikalva	Bharath kumar paidikalva
8	18572	Bhaskar Kamatham	K. Bhaskar
9	18573	Chakri Paipuri	P. Chakri
10	18574	Charan Ulasala	U. Charan
11	18575	Chandini Velagambeti	V. Chandini
12	18576	Charan Tej Markapuram	M. charan Tej
13	18577	Deepa Kadirimangalam	K. Deepa
14	18578	Dhana Kumar Kota	K. Dhana kumar
15	18579	Eswar Kanakappagari	Eswar. K.
16	18580	Eswara Narasimha Vara Prasad	E. Narasimha vara prasad
17	18581	Govardhan Krishna Vallepu	V. Govardhan Krishna
18	18582	Guru Swamy Velampati	V. Guru Swamy
19	18583	Hari Prasad Mala	M. Hari Prasad
20	18584	Haritha Pujari	P. Haritha
21	18585	Hemanth Rajolu	Hemanth Rajolu
22	18586	Indhumathi Lalapeta	L. Indhumathi
23	18587	Jyothi Kuruva	K. Jyothi
24	18588	Kishore Kovuru	K. Kishore
25	18589	Lakshmi Manne	M. Lakshmi
26	18590	Leelakrishna Gadiboina	G. Leelakrishna
27	18591	Manoj Kumar Pallipattu	P. Manoj Kumar
28	18592	Manoj Naramala	N. Manoj
29	18593	Nanda Kumar Kala	K. Nanda kumar

S.No	Roll No	Name of the Student	Signature
30	18594	Pavan Kumar Reddy Vengareddy	Pavan Kumar Reddy . V
31	18595	Penchala Narasimhulu Amooru	penchala Narasimhulu Amooru
32	18596	Ramesh Gunda	G. Ramesh
33	18597	Ramu Goud Ramagouni	R. Ramu Goud
34	18598	Rithik.K.M	K.M. Rithik
35	18599	Shireesha Kuruva	Shireesha Kuruva
36	18600	Shyam Kumar Pokuru	P. Shyam Kumar
37	18601	Siva Sankar Bijigam	Siva Shankar Bijigam
38	18602	Sreedevi Chittoor	C. Sreedevi
39	18603	Sreekanth.K	K. Sreekanth
40	18604	Suresh Madigundu	Suresh madigundu
41	18605	Thimmaiah Siruvella	S. Thimmaiah
42	18606	Vamsi Krishna Naidu Kommineni	K. Vamsi Krishna Naidu
43	18607	Venkata Amaranath Malasani	M. Venkata Amaranath
44	18608	Venkata Anirudh Reddy Venna	V. Venkata Anirudh Reddy
45	18609	Venkata Chanukya Kethineni	K. Venkata Chanukya
46	18610	Venkat Nisanth Kanupuru	K. Venkat Nisanth
47	18611	Venkateswarlu Bingi	Bingi Venkateswarlu
48	18612	Venu Bharath.R.T	Venu Bharath.R.T
49	18613	Vijya Goud Ediga	E. Vijya Goud
50	18614	Vinay Sada	S. Vinay
51	18474	Aravind Chakravarthy	Aravind Chakravarthy
52	18476	Dastagiri Boya	R. Dastagiri
53	18477	Erri Swamy Getu	G. Erri Swamy
54	18478	Marimeni Niranjana	Niranjana Marimeni
55	18480	Ramesh Boya	R. Ramesh
56	18481	Yuva raj Kante	K. Yuva Raj
57	18451	Nanda Kumar.S	S. Nanda Kumar
58	18452	Narendra Mattem	M. Narendra
59	18453	Obul Reddy Boggala	B. Obul Reddy
60	18454	Pavan Kumar.N	N. Pavan Kumar

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SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"

Batch 1 & 2

ACADEMIC YEAR: 2019-20

Final Examination

48
50

15/02/2020

Max Marks: 50
Hours

Duration: 2

Name of the Student : N. Nagadinesh II, B.Sc. MPCs-EM

MULTIPLE CHOICE QUESTIONS EACH QUESTION CARRY 2 MARKS

1. Right clicking on something in word
 - A. Delete the object
 - B. Nothing the right mouse button is there for left handed people
 - ✓ C. Opens the short-cut list
 - D. Insert the picture
2. We can create the new shortcut key from
 - A. Options from Tools menu
 - ✓ B. Customize from Tools menu
 - C. Properties from File menu
 - D. Auto Correct from Toolbars menu
3. The autocorrect tool
 - A. Provides statistics about your document
 - B. Displays words with the same or similar meaning
 - C. Checks the grammar in the document
 - ✓ D. Checks for misspelled words as you type and correct them as defined
4. Which key or key combination will move the insertion point to the bottom of your document?
 - ✓ A. Ctrl+End
 - B. PageDown
 - C. Ctrl+Home
 - D. End
5. To display hyperlink fields in a Word document, you can press the
 - A. Ctrl+Shift+F9 key
 - B. Shift + F9
 - C. Ctrl+Alt key
 - ✓ D. None of above
6. To center worksheet titles across a range of cells, you must
 - ✓ a. Select the cells containing the title text plus the range over which the title text is to be centered
 - b. Widen the columns
 - c. Select the cells containing the title text plus the range over which the title text is to be unfettered
 - d. Format the cells with the comma style
7. How do you delete a column?

- a. Select the column heading you want to delete and select the Delete Row button on the standard toolbar
- b. Select the column heading you want to delete and select Insert Delete from the menu
- c. Select the row heading you want to delete and select Edit>Delete from the menu
- ✓ d. Right click the column heading you want to delete and select delete from the shortcut menu

8. How can you find specific information in a list?

- a. Select Tools > Finder from the menu
- b. Click the Find button on the standard toolbar
- ✓ c. Select Insert > Find from the menu
- ✓ d. Select Data > Form from the menu to open the Data Form dialog box and click the Criteria button

9. When integrating word and excel, word is usually the

- a. Server
- b. Destination
- c. Client
- ✓ d. Both b and c

10. When a label is too long to fit within a worksheet cell, you typically must

- a. Shorten the label
- ✓ b. Increase the column width
- c. Decrease the column width
- d. Adjust the row height

11. When you have completely finished working with a document you should

- ✓ A. Close it
- B. Save it
- c. Type it
- D. Edit it

12. To cut or copy text you must first

- A. Click the remove/duplicate button
- ✓ B. First highlight the text and click the Copy or Cut command from edit menu
- c. Click on File menu
- D. None of above

13. How many items can you copy to the Office Clipboard

- ✓ A. 24
- B. 20
- C. 12
- D. 10

14. The command cross reference is available in

- A. View menu
- B. Format menu
- ✓ C. Insert menu
- D. Tools menu

15. We can send the word documents to

- ✓ A. Microsoft PowerPoint
- B. Microsoft Excel

- c. Microsoft Access
- d. All of above

16. Which of the following are types of sound files?

- a. LOG files
- b. DAT files
- c. WAV files
- d. DRV files

17. Which of the following can you use to add times to the slides in a presentation? a.

- a. Microsoft graph
- b. Microsoft Table
- c. Microsoft Excel
- d. Microsoft Word

18. Which view in PowerPoint can be used to enter speaker comments? a. Normal

- b. Slide show
- c. Slide sorter
- d. Notes page view

19. The best way to insert a new slide in a presentation is to use the

- a. Normal view
- b. Special view
- c. Slide show view
- d. Slide sorter view

20. Which is the best view for getting your thoughts for a presentation out on the computer? ☒ a. Outline view

b. Notes page view

c. Slide sorter view

d. Slide view

21. The default page orientation of word document is

A. Landscape

☒ B. Portrait

C. Long sides

D. Double long side

22. You can print the documents

A. Print icon from Standard toolbar

B. Print command from File menu

C. Ctrl+P

☒ D. All of the above

23. The thesaurus tool

A. Finds repeated words in document

☒ B. Check for synonyms and antonyms

C. Identifies words with capitalization problems

D. All of the above

24. To view smaller text on the screen you can

A. Decrease the font points

☒ B. Decrease the zoom percentage

C. Increase the view percentage

D. Increase the editing percentage

25. Which of the following is not option for changing the case of the text?

A. Indent case

☒ B. Sentence case

C. Toggle case

D. Lower case

*****All the Best*****

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
MARKS SHEET BATCH 1

S.No	Roll No	Name of the Student	Marks in Figure	Marks in Words
1	18511	Anil Kumar Mukku	46	four Six
2	18512	Anil Kumar Talari	48	four Eight
3	18513	Balaji.E	42	four Two
4	18514	Balanjineyulu Mangala	44	four four
5	18515	Balarama Krishna Yadav A	44	four four
6	18516	Charan Kuamar Mandati	48	four Eight
7	18517	Chenchaiiah Gunji	46	four Six
8	18518	Dinesh Kumar Mangali	44	four four
9	18519	Guru Sai Kumar.P.V	42	four Two
10	18520	Harshavardhan Reddy B	42	four Two
11	18521	Hemanth Kumar Paipuru	44	four four
12	18522	Hemasri Kuppa	46	four Six
13	18523	Jagan Goud Nandimalla	48	four Eight
14	18524	Jayalakshmi Kailasam	50	fifty Only
15	18525	Mabbu Sharif Gajuguduru	42	four Two
16	18526	Mahendra Dasari	44	four four
17	18527	Manendra Kumar R	46	four Six
18	18528	Mohan Krishna Kotekallu	48	four Eight
19	18529	Mounika Gali	42	four Two
20	18530	Mounika Sanambatta	44	four four
21	18531	Nagadinesh Nagisetty	48	four Eight
22	18532	Narasimhudu Kukka	46	four Six
23	18533	Naveen Kumar Kanchi	44	four four
24	18534	Purushotham Reddy B	42	four Two
25	18535	Raj Kumar Giddaiahgari	44	four four
26	18536	Rajesh Golla	46	four Six
27	18537	Rajesh Pannamareddi	48	four Eight
28	18538	Rakesh Gajjala	44	four four
29	18539	Rasool Yendhoti	48	four Eight
30	18540	Ravi Teja Kamineni	42	four Two
31	18541	Saiteja Padma	42	four Two
32	18542	Sandhya Devarakonda	44	four four
33	18543	Sandhya Putturu	46	four Six
34	18544	Sasidhar Bottla	48	four Eight
35	18546	Shaik Yasmin	42	four Two
36	18547	Shireesha Gounori	44	four four
37	18548	Siva Sankar Thati	46	four Six
38	18549	Sravan Kumar Chagalamarri	48	four Eight
39	18550	Sreenivasulu Boya	42	four Two
40	18551	Subbarayudu Bayanaboina	44	four four

S.No	Roll No	Name of the Student	Marks in Figure	Marks in Words
41	18552	Suresh Kota	44	four four
42	18553	Suresh Mangali	46	four six
43	18404	Aravind Obalavaram	42	four Two
44	18405	Chandrasekhar.B	44	four four
45	18406	Dathathreyudu Bandarupalli	48	four Eight
46	18407	Hareesh Derangi	46	four six
47	18408	Munisekhar Pamanji	46	four six
48	18409	Sunil Chemuru	42	four Two
49	18411	Suresh Jangamreddy	44	four four
50	18412	Vishnuvardhan Chalampalem	48	four Eight
51	18441	Bharadhwaz Yannam	44	four four
52	18442	Giri Kondari	42	four Two
53	18443	Govardhan Dattam	44	four four
54	18444	Hemadri.P	46	four six
55	18445	Hemantha Kumar Aare	48	four Eight
56	18446	Kalathur Vetti Nithish	46	four six
57	18447	Karimulla Pinjari	48	four Eight
58	18448	Mallikarjuna Avula	44	four four
59	18449	Mahesh Vanka	46	four six
60	18450	Manjunath Bhagathagalla	46	four six

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SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
MARKS SHEET BATCH 2

S.No	Roll No	Name of the Student	Marks in Figure	Marks in Words
1	18565	Ajay Kumar Bandar	46	four six
2	18566	Ajay Kumar Medikonda	48	four eight
3	18567	Adishesu Mabbu	44	four four
4	18568	Anitha Thalari	42	four two
5	18569	Aravind Naidu Buraga	50	fifty only
6	18570	Balaji Ballikura	46	four six
7	18571	Bharath Kumar Paidikalva	42	four two
8	18572	Bhaskar Kamatham	48	four eight
9	18573	Chakri Paipuri	44	four four
10	18574	Charan Ulasala	46	four six
11	18575	Chandini Velagambeti	44	four four
12	18576	Charan Tej Markapuram	48	four eight
13	18577	Deepa Kadirimangalam	46	four six
14	18578	Dhana Kumar Kota	44	four four
15	18579	Eswar Kanakappagari	46	four six
16	18580	Eswara Narasimha Vara Prasad	48	four eight
17	18581	Govardhan Krishna Vallepu	40	forty only
18	18582	Guru Swamy Velampati	42	four two
19	18583	Hari Prasad Mala	44	four four
20	18584	Haritha Pujari	46	four six
21	18585	Hemanth Rajolu	48	four eight
22	18586	Indhumathi Lalapeta	42	four two
23	18587	Jyothi Kuruva	42	four two
24	18588	Kishore Kovuru	44	four four
25	18589	Lakshmi Manne	46	four six
26	18590	Leelakrishna Gadiboina	46	four six
27	18591	Manoj Kumar Pallipattu	44	four four
28	18592	Manoj Naramala	48	four eight
29	18593	Nanda Kumar Kala	44	four four
30	18594	Pavan Kumar Reddy Vengareddy	48	four eight
31	18595	Penchala Narasimhulu Amooru	46	four six

32	18596	Ramesh Gunda	42	four Two
33	18597	Ramu Goud Ramagouni	44	four four
34	18598	Rithik.K.M	46	four six
35	18599	Shireesha Kuruva	44	four four
36	18600	Shyam Kumar Pokuru	46	four six
37	18601	Siva Sankar Bijigam	48	four Eight
38	18602	Sreedevi Chittoor	46	four six
39	18603	Sreekanth.K	44	four four
40	18604	Suresh Madigundu	42	four Two
41	18605	Thimmaiah Siruvella	42	four Two
42	18606	Vamsi Krishna Naidu Kommineni	44	four four
43	18607	Venkata Amaranath Malasani	46	four six
44	18608	Venkata Anirudh Reddy Venna	48	four Eight
45	18609	Venkata Chanukya Kethineni	42	four Two
46	18610	Venkat Nisanth Kanupuru	44	four four
47	18611	Venkateswarlu Bingi	46	four six
48	18612	Venu Bharath.R.T	48	four Eight
49	18613	Vijya Goud Ediga	44	four four
50	18614	Vinay Sada	48	four Eight
51	18474	Aravind Chakravarthy	42	four Two
52	18476	Dastagiri Boya	44	four four
53	18477	Erri Swamy Getu	46	four six
54	18478	Marimeni Niranjana	48	four Eight
55	18480	Ramesh Boya	44	four four
56	18481	Yuva raj Kante	48	four Eight
57	18451	Nanda Kumar.S	46	four six
58	18452	Narendra Mattem	44	four four
59	18453	Obul Reddy Boggala	42	four Two
60	18454	Pavan Kumar.N	44	four four

Reddy

S G S ARTS COLLEGE, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN MS-OFFICE

Academic year : 2019 - 20

REPORT OF THE COURSE:

1. Name of the Course : MS-OFFICE
2. Name of the Course Coordinator : P Udaya
3. No of Students Admitted : 100 (60 per batch)
4. Faculty Members : 03

S.NO	Name of the Teacher	Paper	No of Hours (Classes) Engaged	
			Theory	Problem Solving / Practises
1.	Smt. B.Triveni	I	6	6
2.	Smt. N.Jayalakshmi	I	6	6
3.	Sri D.Chaithanya Kumar	I	8	8

5. Number of Students Appeared for Exam : Batch I : 60
Batch II : 60
6. Number of Students Passed : Batch I : 60
Batch II : 60
7. Percentage of Results : 100%
8. Date of Theory Examination : 15-02-2020
9. Name of the Examiner : B.Triveni, N.Jayalakshmi
10. Date of Declaration of Results : 16-02-2020

HoD


Signature of Course Coordinator

PHOTO GALLERY

Distribution of Certificates to B.S.c MPCs and MSCs Second year students for completing the Certificate Course on the Topic “MS Office “ by Guest of Honour Dr. N. Venugopal Reddy, H.O.D, Physics, S.V. Arts College, Tirupati.

Guest of Honor Dr. N. Venugopal Reddy & Staff of Computer Science Department



STUDENT NAME : P. SANDHYA II BSc MPCs EM



STUDENT NAME : T.ANIL KUMAR II BSc MPCs EM



Student Name: P.Manoj Kumar, II B.S.c-MSCs



Student Name: C.Sreedevi, II B.S.c-MSCs

Certificates of B.S.c MPCs and MSCs Second year students for completing the Certificate Course on the Topic "MS Office" .



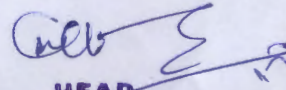
"OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

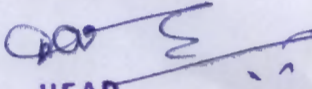
Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

Department of Chemistry

BATCH :: 2016-19		GROUP::MPC		
SEMESTER WISE PROGRESSION OF STUDENT RESULTS				
SEMESTER	TOTAL NUMBER OF STUDENTS IN THE CLASS	TOTAL NUMBER OF SLOW LEARNERS IN THE CLASS	TOTAL NUMBER OF ADVANCED LEARNERS IN THE CLASS	TOTAL NUMBER OF AVERAGE LEARNERS IN THE CLASS
II SEMESTER (Based on the results of I semester (2016-17))	44	23	2	19
III SEMESTER (Based on the results of II semester (2016-17))	44	15	4	25
IV SEMESTER (Based on the results of III semester (2017-18))	39	21	3	15
V SEMESTER (Based on the results of IV semester (2017-18))	35	14	4	17
VI SEMESTER (Based on the results of V semester (2018-19))	31	8	5	18


HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

BATCH :: 2016-19		GROUP: CBZ		
SEMESTER WISE PROGRESSION OF STUDENT RESULTS				
SEMESTER	TOTAL NUMBER OF STUDENTS IN THE CLASS	TOTAL NUMBER OF SLOW LEARNERS IN THE CLASS	TOTAL NUMBER OF ADVANCED LEARNERS IN THE CLASS	TOTAL NUMBER OF AVERAGE LEARNERS IN THE CLASS
II SEMESTER (Based on the results of I semester (2016-17))	24	22	NIL	2
III SEMESTER (Based on the results of II semester (2016-17))	27	4	7	16
IV SEMESTER (Based on the results of III semester (2017-18))	25	11	NIL	14
V SEMESTER (Based on the results of IV semester (2017-18))	26	12	1	13
VI SEMESTER (Based on the results of V semester (2018-19))	24	2	3	19


 HEAD
 DEPARTMENT OF CHEMISTRY
 S. G. S. ARTS COLLEGE
 TIROPATHI

STUDENT UNION OFFICE BEARERS:

Phone: (0877) 2264599

E-mail: sgsartscollegettd@gmail.com

Web: sgsac.edu.in



TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

2018-2019

Sno	Class	Name of the Representative	Roll No
1	III B.A-HPT-TM	Mr. S.Yerramala(Union Chairman)	16284
2	III B.Sc-BBC-EM	Mr. S.Snehalatha(Union Secretary)	16736
3	III B.Sc-MSCs-EM	Mr. P.Gurusai(Maths & Science Association Secretary)	16576
4	II B.Sc-BBC-EM	Kum. T.Jahanavi(Maths & Science Association Joint Secretary)	17672
5	III B.A-HEP-TM	Mr. B.Pavan Kumar(Arts Association Secretary)	16358
6	II B.A - HEP-EM	Mr. K.Rajendra Prasad(Arts Association Joint Secretary)	17307
7	III B.Com-CA-EM	Mr. N.Purushottam(Commerce Association Secretary)	0217008579
8	II B.Com-CA-EM	Kum. K.Deepa(Commerce Association Joint Secretary)	17167
9	III B.Sc-MZC-EM	Kum. P.Shilpa(Language Association Secretary)	16675
10	II B.Sc-MPCs-EM	Mr. M.Joseph(Language Association Joint Secretary)	17500
11	III B.Sc-MPS-TM	Mr. B.Sumanth Naik(Sports&Games Association Secretary)	317008004
BSc Class Representatives			
Sno	Class	Name of the Representative	Roll No
1	III B.Sc-MSCs-EM	Mr. D.Venkata Rami Reddy	16612

2	III B.Sc-MPCs-EM	Mr. T.Lokesh	16534
3	III B.Sc-MPC-EM	Mr.M.Sandeep Reddy	0317008027
4	III B.Sc-MPC-TM	Mr. D.L. Viswanath	0317008020
5	III B.Sc-MPS-TM	Mr. M.Baby	0317008037
6	III B.Sc-BBC-EM	Mr.B.Lokesh Reddy	0317008715
7	III B.Sc-MZC-EM	V.Naveen Sahit	0317008589
8	III B.Sc-CT&HM-EM	Mr. J.Hemanth	16760
9	III B.Sc-CBZ-TM	Mr,Chandra sekhar reddy	0317008664
10	II B.Sc-MSCs-EM	Mr. P.Lokesh	17556
11	II B.Sc-MPCs-EM	Mr. M.Prasanth	17515
12	II B.Sc-MPC-EM	Mr. A.K.Jagadeesh	17465
13	II B.Sc-MPC-TM	K.Chandu	17457
14	II B.Sc-MPS-TM	Mr. A.Jayakumar	17598
15	II B.Sc-BBC-EM	Mr. B.Mahesh	17681
16	II B.Sc-MZC-EM	Mr. P.Devananda Goud Mr.G.Jaya Kumar	17719 0318008263
17	II B.Sc-CT&HM-EM	Mr Aravindhan.V	17768
18	II B.Sc-CBZ-TM	Mr. M.Venkatesh	17643
19	I B.Sc-MSCs-EM	Mr. G.Ramesh	18596
20	I B.Sc-MPCs-EM	Mr. Natteti Venkateswarulu	18560
21	I B.Sc-MPC-EM	SurendraGuvvala M.Raghu Ram Kumr	18462
22	I B.Sc-MPC-TM	Chandrasekhar.B	18405
23	I B.Sc-MPS-TM	Mr. B.Dastagiri	18476

24	IB.Sc-BBC-EM	M.Marappa Mr. T.Bharat Kumar	18652
25	I B.Sc-MZC-EM	Chakrapani Tharun M.Gurramma	18704
26	I B.Sc-CT&HM-EM	Mr. M.Naveen	18772
27	I B.Sc-CBZ-TM	K.Durga ManendraAvula	18628
BCom Class Representatives			
28	III B.Com-CA-EM	Mr. P.Seshari Naidu	0217008503
29	III B.Com-GEN-EM	Mr. D.Rajasekhar	217008582
30	III B.Com-GEN-TM	Mr A.Koteswar Rao	217008501
31	II B.Com-CA-EM	K.Deepa Mr. V.Harish	17176
32	II B.Com-GEN-EM	K.Ramesh KishoreKullanivagaira	17095
33	II B.Com-GEN-TM	B.Babu Naik Giri Babu.K	17010
34	I B.Com-CA-EM	Mr D.Mahanandi Mr. M.Venkatsuresh	18187
35	IB.Com-GEN-EM	Mr. Y.Sasi Kumar Mr. C.Jaswanth	18069
36	I B.Com-GEN-TM	Mr. Y.Sai Kumar G.Kranthi Kiran	18042

BA Class Representatives			
37	III B.A-HEP-EM	Mr. B.Sunil	16361
38	III B.A-HEP-TM	Mr. A.Jayaram	16307
39	III B.A-HPT-TM	Mr. M.Venkataramana	16289
40	III B.A-ASCA-EM	Bhanu Prakash.B	16223
41	III B.A-MES-EM	B.Ashok	116008524
42	II B.A-HEP-EM	Dhana SekharAllavala	17343
43	II B.A-HEP-TM	NaveenPoola	17302
44	II B.A-HPT-TM	Mr.S.Yerramala	117008019
45	II B.A-ASCA-EM	Mr.P.Giddaiah	117008506
46	II B.A-MES-EM	Mr.G.Venkata Ratnam	17408
47	I B.A-HEP-EM	Mr. Surya Mahesh Bala	0119008075
48	I B.A-HEP-TM	Mr. K.Venkatesh	0119008072
49	I B.A-HPT-TM	Mr.M.Eswar Naik	0119008093
50	I B.A-ASCA-EM	Mr.P.Nagendra	0119008015
51	I B.A-MES-EM	S.Lalith Madhav	0119008082

HOSTEL COMMITTEE			
Sno	Committee Members	Designation	Role
1	Dr. C.Ramesh	Principal	Chief Warden
2	Sri P.Muni Bhaskar Rao	HOD,Mathematics	Warden
3	Sri K. Rajesh	Lecturer,Economics	Dy.Warden
4	E.Anil Kumar	Student,III BSc-MPCs-EM	Store Secretary
5	B.Sumanth Naik	Student,III BSc-MPS-TM	Sports Secretary
MESS COMMITTEE			
Sno	Name	Group	Role
1	G.Hariprasad	III BSc-MPCs-EM	Associate Members
2	P.Bharath Naveen	III BBC-EM	
3	C.Sreenu	III BA-MES-EM	
4	N.Raghu	III Bcom-CA-EM	
5	Subramanyam	II Bcom-Gen-EM	
6	A.Venkatesu	II BSc-MSCs-EM	
7	N.Udaya Kumar Reddy	II BA-ASCA-EM	

STUDENT UNION OFFICE BEARERS:

Department of Chemistry

To foster leadership qualities among the students, the institution has constituted a Student Council and the toppers in the academics are selected as representatives of student Council. The following are the students of chemistry who represent the student council.

2020-2021

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	Kum. M.Guramma	III BSc-MZC-EM	18712	UNION Chairmen
2	R.Gayathri	III BSc-MZC-EM		Life Sciences Association Secretary
3	A.Chandini	II BSc-BBC-EM		Life Sciences Association Joint Secretary

2019-2020

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	Kum. Jahnvi Takkolu	III BSC-BBC-E.M	17672	Life Sciences Association Secretary
2	Kum. M.Guramma)	II BSc-MZC-EM	18712	Life Sciences Association Joint Secretary
3	Mr. D.Sandeep	III B.Sc-MZC-EM	17741	Sports&Games Association Secretary

2018-2019

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	Mr. S.Snehalatha	III B.Sc-BBC-EM	16736	Union Secretary)
2	Kum. T.Jahanavi	II B.Sc-BBC-EM	17672	Maths & Science Association Joint Secretary
3	Kum. P.Shilpa	III B.Sc-MZC-EM	16675	Language Association Secretary

2017-2018

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	G.Venkateswarulu	III B.Sc-MPC-EM		Maths & Science Association Joint Secretary
2	S. Snehalatha	II B.Sc-BBC-EM		Language Association Joint Secretary

2016-2017

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	G.Venkateswarulu	II B.Sc-MPC-EM		Maths & Science Association Joint Secretary